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Tank 241-C-112 Characterization Report

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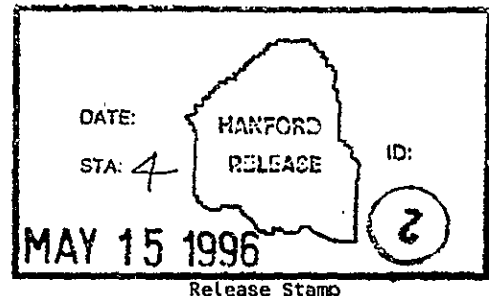
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Tank Characterization Data Report: Tank 241-C-112

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March 1996

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Office of Environmental Restoration and
Waste Management

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EXECUTIVE SUMMARY

Tank 241-C-112 is a Hanford Site Ferrocyanide Watch List tank that was most recently sampled in March 1992. Analyses of materials obtained from tank 241-C-112 were conducted to support the resolution of the Ferrocyanide Unreviewed Safety Question (USQ) and to support *Hanford Federal Facility Agreement and Consent Order*[1] (Tri-Party Agreement) Milestone M-10-00.

Analysis of core samples obtained from tank 241-C-112 strongly indicates that the fuel concentration in the tank waste will not support a propagating exothermic reaction. Analysis of the process history of the tank as well as studies of simulants provided valuable information about the physical and chemical condition of the waste. This information, in combination with the analysis of the tank waste, supports the conclusion that an exothermic reaction in tank 241-C-112 is not plausible. Therefore, the contents of tank 241-C-112 present no imminent threat to the workers at the Hanford Site, the public, or the environment from its ferrocyanide inventory. Because an exothermic reaction is not credible, the consequences of this accident scenario, as promulgated by the General Accounting Office, are not applicable.[2]

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[1]Ecology, EPA, and DOE, 1992, *Hanford Federal-Facility Agreement and Consent Order*. 2 vols. , Washington State Department of Ecology, U.S. Environmental Protection Agency, and U.S. Department of Energy, Olympia, Washington.

[2]Peach, J. D., 1990, "Consequences of Explosion of Hanford's Single-Shell Tank are Understated," (Letter B-241479 to C. M. Synar, Chairman of Environment, Energy, and Natural Resources Subcommittee, Committee on Government Operations, House of Representatives), GAO/RCED-91-34, General Accounting Office, Washington, D.C.

It is probable that tank 241-C-112 exceeds the 1,000 g-mol inventory criteria established for the Ferrocyanide USQ; however, extensive energetic analysis of the waste has determined a maximum exothermic value of -9 cal/g dry waste. This value is substantially below any levels of concern (-75 cal/g).[3] In addition, an investigation of potential mechanisms to generate concentration levels of radionuclides high enough to be of concern was performed. No credible mechanism was postulated that could initiate the formation of such concentration levels[4] in the tank.

Tank 241-C-112 waste is a complex material made up primarily of water and inert salts. The insoluble solids are a mixture of phosphates, sulfates, and hydroxides in combination with aluminum, calcium, iron, nickel, and uranium. Disodium nickel ferrocyanide and sodium cesium nickel ferrocyanide probably exist in the tank; however, there appears to have been significant degradation of this material since the waste was initially settled in the tank. Most of the ^{137}Cs precipitated during the scavenging campaign (1955 to 1957)[5] appears to still remain in the tank in an insoluble form, probably bound with the remaining ferrocyanide. Total cyanide analyses substantiate the energetics results. The soluble analytes are primarily sodium, nitrate, and nitrite (Table ES-1).

[3]Jewett, J. R., 1992, "Energy Measurements for Disqualifying Waste Tanks from Watch Lists," Meeting Minutes, October 22, 1992, Westinghouse Hanford Company, Richland, Washington.

[4]Dickinson, D. R., J. M. McLaren, G. L. Borsheim, M. D. Crippen, 1993, *Credibility of Drying Out Ferrocyanide Tank Waste Sludge by Hot Spots*, WHC-EP-0648, Westinghouse Hanford Company, Richland, Washington.

[5]Borsheim, G. L. and B. C. Simpson, 1991, *An Assessment of the Inventories of the Ferrocyanide Watchlist Tanks*, WHC-SD-WM-ER-133, Rev. 0, Westinghouse Hanford Company, Richland, Washington.

Comparisons of the calculated bulk inventories for various analytes of concern show that tank 241-C-112 is within established operating safety requirements for heat-load, organic content, and plutonium inventory. A substantial amount of free liquid remains in the tank. However, no effort to remove the remaining liquid is planned because there is less than 189,000 L (50,000 gal) of drainable liquid in the tank (thus meeting stabilization criteria). Tank 241-C-112 is considered a sound, non-leaking tank. Figure ES-1 presents a summary of tank status data for 241-C-112.

Table ES-1.

Estimated Major Analyte Inventory	Water	Na	NO$_3$	U	NO$_2$	P	Fe	Ca	Ni	Al	TOC
Avg. Wt% (wet solids)	53.0	9.3	.57	5.0	4.2	2.4	1.9	1.5	1.3	1.0	0.4

Estimated Total Mass of Waste in 241-C-112 (range): 608-611 Mg

Supernate: 48-175 Mg

Wet Solids: 434-563 Mg

Estimated Fission Product Inventory Range	^{137}Cs	^{90}Sr
Bulk Inventory. Ci (wet solids)	217, 200-286. 000	162, 000-183, 500
Heat Generation. w	1.025-1.350	1,090-1,230

Estimated Plutonium/Americium Inventory Range	^{238}Pu	$^{239/240}$Pu	^{241}Am
Bulk Inventory. Ci (wet solids)	5.9-7.7	67.3-87.1	330-430
Bulk Inventory. g (wet solids)	0.34-0.44	1.100-1.400	100-125

Analyte	Safety Issue Criteria	Calculated/Measured Value
Na$_2$NiFe(CN)$_4$ (wet solids)	1.000 g-mol	8,700-11,500 g-mol
ΔH (dry basis)	-75 cal	-9 cal
$^{239/240}$Pu	50 kg	1.1-1.4 kg
Temperature	300 °F (149 °C)	85 °F (29 °C)
Heat Load	.11.72 kw	2.26-2.44 kw
Organic Content (TOC. dry basis)	3.0 wt% TOC (10% sodium acetate equivalent)	0.75 wt% TOC

Figure ES-1. Tank: 241-C-112.

Tank: 241-C-112
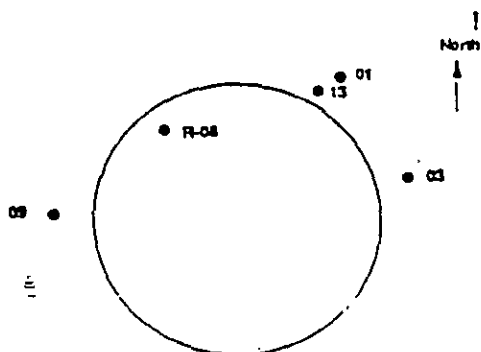
Riser and Drywell Locations
Tank 241-C-112

Tank Description
Type: Single Shell
Constructed: 1944
In-service: 11/46
Out of Service: 1976
Diameter: 75'
Usable Depth: 16'
Capacity: 530K gallons
Bottom shape: Dish
Hanford Coordinates:
    43.074' North
    48.398' West
Ventilation: Passive

Leak Detection System
Surface Level:
    FIC Riser- None
    Manual Tape Riser- R-08
    LOW Riser(s)- None

Number of External Drywells: 4
Number of Lateral Wells: None

Tank Status
Watch List: *Ferrocyanide*
Contents
    Type: Non-Complexed Waste
    Total Waste: 104K gallons
    Supernate Volume: OK gallons
    Drainable Interstitial Liquid: 32K gallons
Isolation Status
    Date Partially Interim Isolated: 12/15/82
    Date Interim Stabilized: 09/90
Surface Level/Leak Status
    Integrity Category: Sound
    Manual Tape Surface Level: 33.50 Inches (12/28/92)
    Last Photographed: 09/18/90
Temperature Status
    Highest temperature during 1992:
        90.50 deg F (08/09/92)
    Comments:
        Temperatures are stable.
Drywell Status
    Comments:
        Current drywell profiles were stable and consistent with established baseline
        profiles.

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LIST OF ACRONYMS

1C	first cycle decontamination waste
AA	atomic absorption
AEA	alpha energy analysis
AES	atomic emission spectroscopy
ANOVA	analysis of variance
CI	confidence interval
CLP	Contract Laboratory Procedure
CST	core sampling truck
CW	cladding waste
DLC	drainable liquid composites
DSC	differential scanning calorimetry
GAO	General Accounting Office
GC	gas chromatograph
GEA	gamma energy analysis
g-mol	gram-moles
HASM	Hanford Analytical Services Management
HDW-EIS	Hanford Defense Waste Environmental Impact Statement
HWVP	Hanford Waste Vitrification Plant
IC	ion chromatography
ICP	inductively coupled plasma - atomic emission spectroscopy
LCS	laboratory control samples
MS	mass spectrometry
NCAW	neutralized current acid waste
PNNI	Pacific Northwest National Laboratory
QA	Quality Assurance
QC	Quality Control
RCRA	*Resource Conservation and Recovery Act of 1976*
RPD	relative percent difference
SOW	statement of work
SST	single-shell tank
TC	total carbon
TGA	thermogravimetric analysis
TIC	total inorganic carbon
TOC	total organic carbon
TOD	total oxygen demand
Tri-Party Agreement	*Hanford Federal Facility Agreement and Consent Order*
TTP	Technical Task Plan
UR	uranium recovery
USQ	unreviewed safety question

1.0 INTRODUCTION

Analysis was conducted on materials obtained from tank 241-C-112 to support the closure of the ferrocyanide unreviewed safety question (USQ). Obtaining measurements that determine overall waste energetics is a key step in closing the ferrocyanide USQ and safety issue. In addition, several of the analytes contributing to the energetic properties of the waste need to be measured as a function of position (e.g., total cyanide and nitrate/nitrite present, water content, and the distribution and inventory of ^{137}Cs and ^{90}Sr in the tank). Other objectives that these measurements and inventory estimates support are as follows:

- Complete *Hanford Federal Facility Agreement and Consent Order* (Tri-Party Agreement) Milestone M-10-00 (Ecology et al. 1992) to sample and analyze two cores from each tank.

- Obtain estimates of both the concentration and total quantity of key analytes relating to other safety issues, such as organics and radionuclides.

- Provide input to risk assessment-based disposal decisions for the waste.

- Implement physical property measurements, such as rheology, bulk density, and particle size. These measurements are necessary for the design and fabrication of retrieval, pretreatment, and final waste disposal systems.

1.1 PURPOSE

This report summarizes the available information regarding the waste in tank 241-C-112, and arranges this information in a useful format for data users in various internal and external organizations.

1.2 SCOPE

This report presents a broad background of preliminary information that was available prior to core sampling, which initially guided the development of the sampling and analysis program. This material includes historical information about the ferrocyanide-scavenging program, transfer records, observations from in-tank photographs, and inferences from waste simulant studies. The results of tank 241-C-112 core sample analyses are summarized and presented, along with a statistical interpretation of the data. The information obtained from historical sources and synthetic waste studies are compared with the actual waste measurements in this report.

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2.0 PRESAMPLING INFORMATION AND EVALUATION

2.1 BACKGROUND

Radioactive wastes from defense operations have accumulated at the Hanford Site in underground waste tanks since the 1940's. During the 1950's, additional tank storage space was required to support the United States defense mission. To obtain additional tank storage volume within a short period of time and to minimize construction of additional storage tanks, Hanford Site scientists developed a process to scavenge radiocesium from tank waste liquids (Sloat 1954, Abrams 1956). Ferrocyanide compounds were used in a carrier-precipitation process to scavenge ^{137}Cs and other soluble radionuclides from the Hanford Site waste tanks. This treatment was used on U Plant waste effluent, bismuth phosphate first cycle decontamination waste (1C), and selected wastes that had been previously discharged to the tanks. The radionuclides settled in the waste tanks and the supernate was discharged to the cribs and trenches. As a result of this process, occupied waste volume in the waste tanks was greatly reduced, while minimizing the amount of long-lived radionuclides discharged to the ground.

In implementing this process, approximately 140 metric tons of ferrocyanide [as $Fe(CN)_6^{-4}$] were added to the tanks. The bulk of the ferrocyanide material is believed to remain in 18 to 24 single-shell tanks (SST). Ferrocyanide is a stable complex of iron(II) ion and cyanide, whose compounds are considered nontoxic because they do not appreciably dissociate in aqueous solutions (Burger 1984). In the presence of oxidizing materials such as nitrates and/or nitrites, ferrocyanide compounds can undergo uncontrolled exothermic reactions in the laboratory by heating them to high temperatures (above 280 °C [540 °F]). The reactive nature of ferrocyanide in the presence of an oxidizer has been known for decades, but the conditions under which the compound can undergo exothermic reactions have not been thoroughly studied. Because the scavenging process involved precipitating ferrocyanides from solutions containing nitrate and nitrite, the potential for a reactive mixture of ferrocyanides and nitrates/nitrites in the SSTs must be evaluated.

2.1.1 Tank 241-C-112 History

Groups of waste tanks that were physically located together and built at the same time are called tank farms at the Hanford Site. The original tank farms (B, C, T, U) were built in 1943 to 1944. Tank 241-C-112 was placed into service in 1946. Each tank has a diameter of 22.9 m (75 ft), an operating depth of 5.2 m (17 ft), and a nominal capacity of 2 million liters (530,000 gal). The basic design of a typical SST is shown in Figure 2-1. The tank was constructed of reinforced concrete with a mild steel liner covering its bottom and sides. The top of the tank is a concrete dome. Tanks such as 241-C-112 were all covered by at least 1.8 m (6 ft) of soil for shielding purposes (Anderson 1990). The tanks in the tank farms were connected in groups of three or four and overflowed from one to another (known as a cascade). Tank 241-C-112 is the last tank in a cascade that includes 241-C-110 and 241-C-111.

Figure 2-1. Typical Single-Shell Tank Diagram.
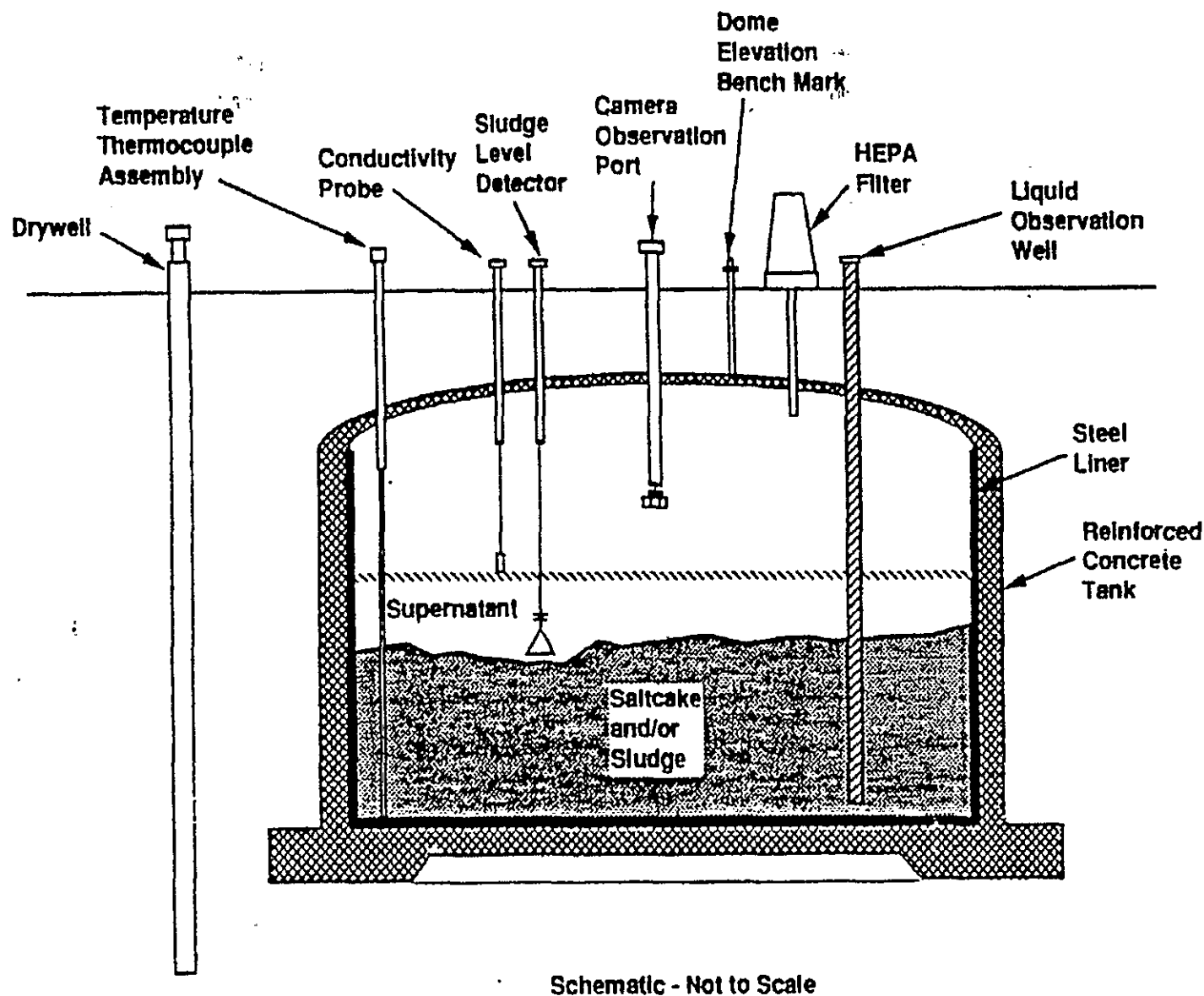
Note: Tank 241-C-112 has a dished bottom.

Cascades served several functions in Hanford Site waste management operations. By cascading tanks, fewer connections needed to be made during waste disposal; consequently, all three tanks were usable without having to connect the active waste transfer line directly to each individual tank. This handling method reduced the likelihood for personnel exposure to the waste and diminished the chances for a loss of tank integrity because of overfilling. Another benefit of the cascades was clarification of the wastes. When used in this manner, most of the solids in the waste slurries routed to the tanks settled in the first tank (241-C-110), and the clarified liquids cascaded on to the other tanks in the series (241-C-111 and 241-C-112). Supernate from the final tank in a cascade series was sometimes routed to a disposal trench. In this way, clarification reduced the potential amount of radiological contamination to the environment.

The first type of waste that tank 241-C-112 received and stored was 1C waste from the bismuth phosphate process (1946 to 1952). This waste would be comparatively high in bismuth, phosphate, and aluminum because aluminum decladding waste was combined with it. The waste was disposed to ground in 1952, leaving a 57,000-L (15,000-gal) heel. The tank was refilled with unscavenged uranium recovery (UR) waste in 1953 and 1954 (Anderson 1990). The UR waste solids were comparatively high in uranium and iron, and low in bismuth and aluminum. The available records do not show whether these wastes were added directly to the tank or through the cascade overflow line from tank 241-C-111. Neither of these waste types had any significant fuel content or heat-generating radionuclides (^{137}Cs or ^{90}Sr) that could contribute to the exothermic potential posed by the ferrocyanide wastes. In late 1955, tank 241-C-112 was emptied. The tank was then used for settling scavenged ferrocyanide waste until 1958. During ferrocyanide-scavenging operations, waste was not cascaded through the 241-C-110, -111, and -112 series. Tank 241-C-112 received the waste slurry in direct transfers from the process vessel (General Electric 1958).

Beginning in May 1955, unscavenged UR waste already stored in 200 East Area underground tanks at the Hanford Site was routed to the 244-CR vault for scavenging (refer to Figure 2-2). The 244-CR vault facility contained stainless steel tanks with chemical addition, agitation, and sampling capabilities. The pH was adjusted with HNO_3 and/or NaOH to pH 9.3 $\pm$ 0.7, and $Fe(CN)_6^{-4}$ and Ni^{+2} ion were added (generally to 0.005 $\underline{M}$ each) to precipitate ^{137}Cs. If laboratory analysis of the feed tank indicated additional ^{90}Sr decontamination was necessary, calcium nitrate was also added (Sloat 1955). There was also an effort to scavenge ^{60}Co with Na_2S. The scavenged waste was then routed to another tank for settling, sampling, and decantation to a crib. The settling tanks for this "In-Farm scavenged" waste were 241-C-108, 241-C-109, 241-C-111, and 241-C-112. These tanks are expected to exhibit significant similarity in chemical content and physical properties. However, no firm conclusions can yet be made, since it is the only tank of this group with analytical results available.

Figure 2-2. In-Farm Flowsheet.

Ferrocyanide Production by "In-Farm" Flowsheet

Settling Tanks C-108, -109, -111, -112
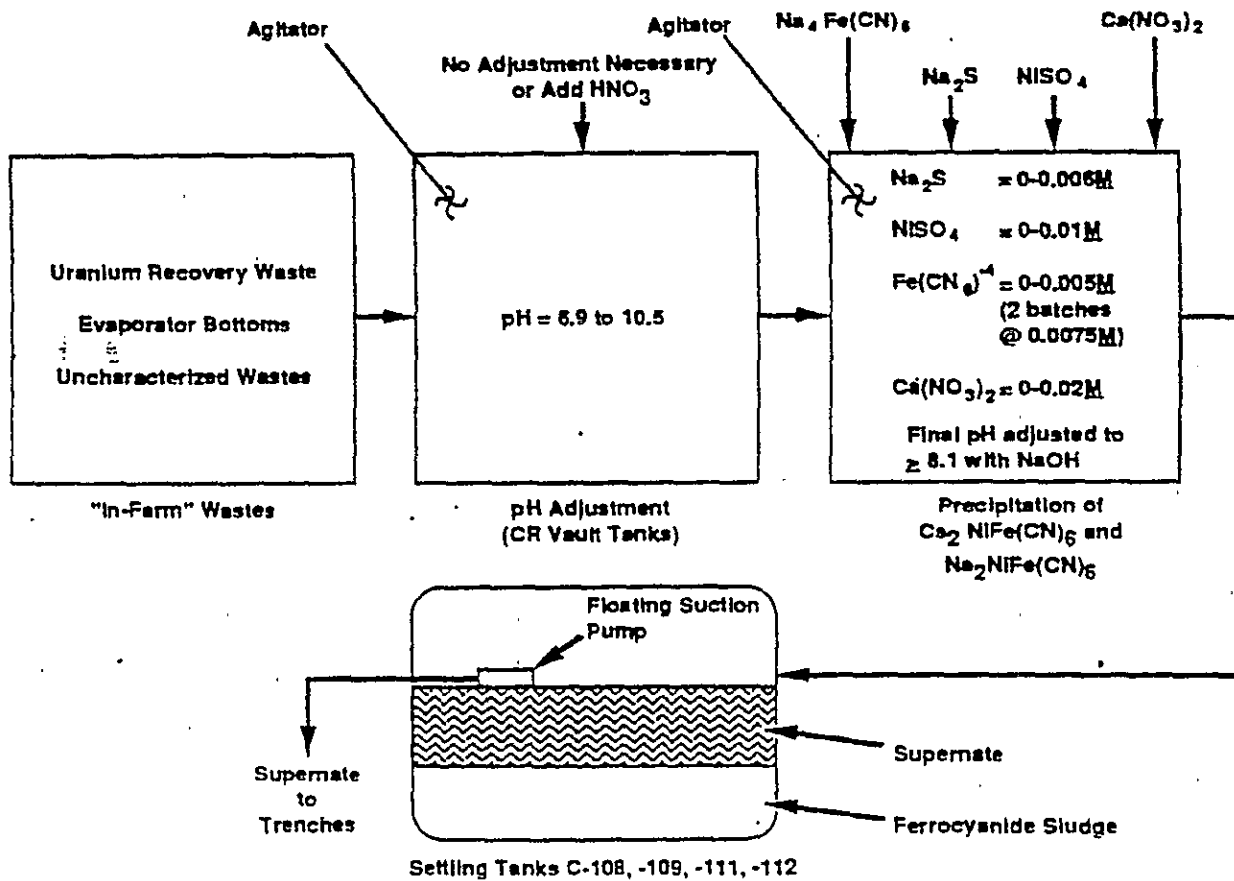
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The In-Farm precipitate comprises approximately 20 to 25% of the total ferrocyanide material in the Hanford Site tank farms. This material is expected to possess a much higher ferrocyanide concentration content than the more prevalent (70% of the total ferrocyanide material) U Plant material. Analytes that differentiate ferrocyanide waste from other wastes are nickel, calcium, and ^{137}Cs.

Over time, additional gravity settling may have compressed the waste layers, increasing the concentration of some of these analytes. However, the effect of radiation and high pH conditions from later waste additions on the waste matrix is largely unknown. Exposure of the waste to these conditions is believed to have degraded the ferrocyanide. However, laboratory results confirming that hypothesis are still pending (Lilga et al. 1992; Babad et al. 1993).

The first transfer of scavenged waste for settling was in the fourth quarter of 1955. In-Farm scavenging was completed in December 1957 (General Electric 1958). The inventory of solids in tank 241-C-112 at the end of the ferrocyanide-scavenging program, as calculated by the Borsheim-Simpson (1991) model, was 318,000 L (84,000 gal) with essentially no free supernate. The scavenging record (General Electric 1958) gives the tank level as 0.67 m (2 ft 2.5 in.) (323,000 L [85,400 gal]). *A History of the 200 Area Tank Farms* (Anderson 1990) reports a total volume of 318,000 L (84,000 gal), but lists only 174,000 L (46,000 gal) of that inventory as solids.

After the end of scavenging in late 1957, tank 241-C-112 remained in active service. However, the tank had relatively limited activity from 1958 to the end of its service life in 1980. The volume is shown as increasing by 189,000 L (50,000 gal) of liquid to approximately 507,000 L (134,000 gal) in the fourth quarter of 1958, but no incoming or outgoing transfers were noted. In the third and fourth quarters of 1960, a total of 996,000 L (263,000 gal) of highly alkaline cladding waste (CW) (a waste known to contain substantial amounts of solids) was added to the tank, but the reported solids inventory (174,000 L [46,000 gal]) did not change (Anderson 1990). Cladding waste solids would have settled on top of the ferrocyanide sludge already present.

Several small transfers with relatively high concentrations of ^{90}Sr occurred after 1958. Waste from the strontium semiworks/hot semiworks was added to the tank with the total volume listed as 2.07 M L (547,000 gal) at the end of 1964 (the reported solids inventory was still only 174,000 L [46,000 gal]). The listed volumes for the first quarter report in 1965 are a total volume of 2.04 M L (538,000 gal), with a solids volume of 485,000 L (128,000 gal) (Anderson 1990). This solids level measurement was apparently the first since additional waste was added to the tank following the last scavenging pumpout in 1958.

The reported waste volume remained essentially unchanged (between 2.01 and 2.04 M L [532,000 and 538,000 gal]) until a transfer of 1.29 M L (340,000 gal) to tank 241-C-104 in the first quarter of 1970. This transfer left a heel of at least 727,000 L (192,000 gal). A floating suction pump transfer would not have transferred any solids because the maximum reported solids level was 485,000 L (128,000 gal). In early 1970,

some B Plant ion-exchange waste (1.24 M L [327,000 gal]) from tank 241-C-110 and drainage to the C-301 catch tank (79,000 L [21,000 gal]) was added to tank 241-C-112. Between 1970 and 1975, the reported solids volume ranged between 454,000 and 522,000 L (120,000 and 138,000 gal), and the total volume reported decreased from 2.06 M L to 2.01 M L (543,000 to 532,000 gal) (Anderson 1990).

Tank 241-C-112 was suspected of leaking and was emptied of pumpable liquid to tank 241-C-103 in 1975-1976 (Anderson 1990); later surveillance never confirmed the suspected leak. Some solids may have been transferred, as the reported tank solids volume decreased from 485,000 L (128,000 gal) to 413,000 L (109,000 gal). However, the solids transferred would have been those that settled on top of the ferrocyanide solids (i.e., CW solids). The previously calculated volume of ferrocyanide sludge was 318,000 L (84,000 gal), with reported volumes ranging between 174,000 and 323,000 L (46,000 and 85,400 gal). Sludge volume in the tank may have decreased between 1958 and 1975 with further settling and compaction from the weight of overlying solids. However, the volume of the ferrocyanide sludge would not have increased during that time frame. There was no mixing equipment in tank 241-C-112 to move the settled ferrocyanide solids into the overlying solids layer. Therefore, it was concluded that no appreciable volume of ferrocyanide solids were transferred to tank 241-C-103 (Borsheim and Simpson 1991). WHC-EP-0640 Rev. 1 (Simpson et al. 1993), has the results of a model that represents the inventory changes in the tanks with various initial conditions and solid formation values after scavenging operations were completed.

The last major waste type was aluminum CW. These materials would be high in aluminum and silica, with a very high pH. However, the solids volume contribution to the tank is unknown because the majority of the solids would be deposited in the first tank to receive the wastes, which was not tank 241-C-112. The high pH of this waste is considered a significant factor affecting the state of the waste matrix. Other wastes had discernable impacts on the bulk characteristics of the tank contents as well. The strontium semiworks waste had a small volume of waste added, but would have a very high ^{90}Sr content because it included strontium recovery and purification waste losses. The B Plant ion-exchange waste was primarily liquid and was not expected to contribute significantly to the solids in the tank. Figure 2-3 presents a timeline histogram showing the waste deposition and stratification in tank 241-C-112, as indicated from the fill history.

2.1.2 Unreviewed Safety Question Declaration

Efforts have been underway since the mid-1980's to evaluate the potential of a ferrocyanide combustion reaction in Hanford Site SSTs (Burger 1989; Burger and Scheele 1990; Burger 1984). In 1987, the *Final Environmental Impact Statement, Disposal of Hanford Defense High-Level Transuranic and Tank Wastes*, hereinafter referred to as the HDW-EIS (DOE 1987), was issued. In the HDW-EIS, it was projected that the bounding "worst-case" accident in a ferrocyanide tank would be an explosion resulting in a subsequent short-term radiation dose to the public of 200 mrem.

Figure 2-3. Solids and Liquid Wastes - Tank 241-C-112.
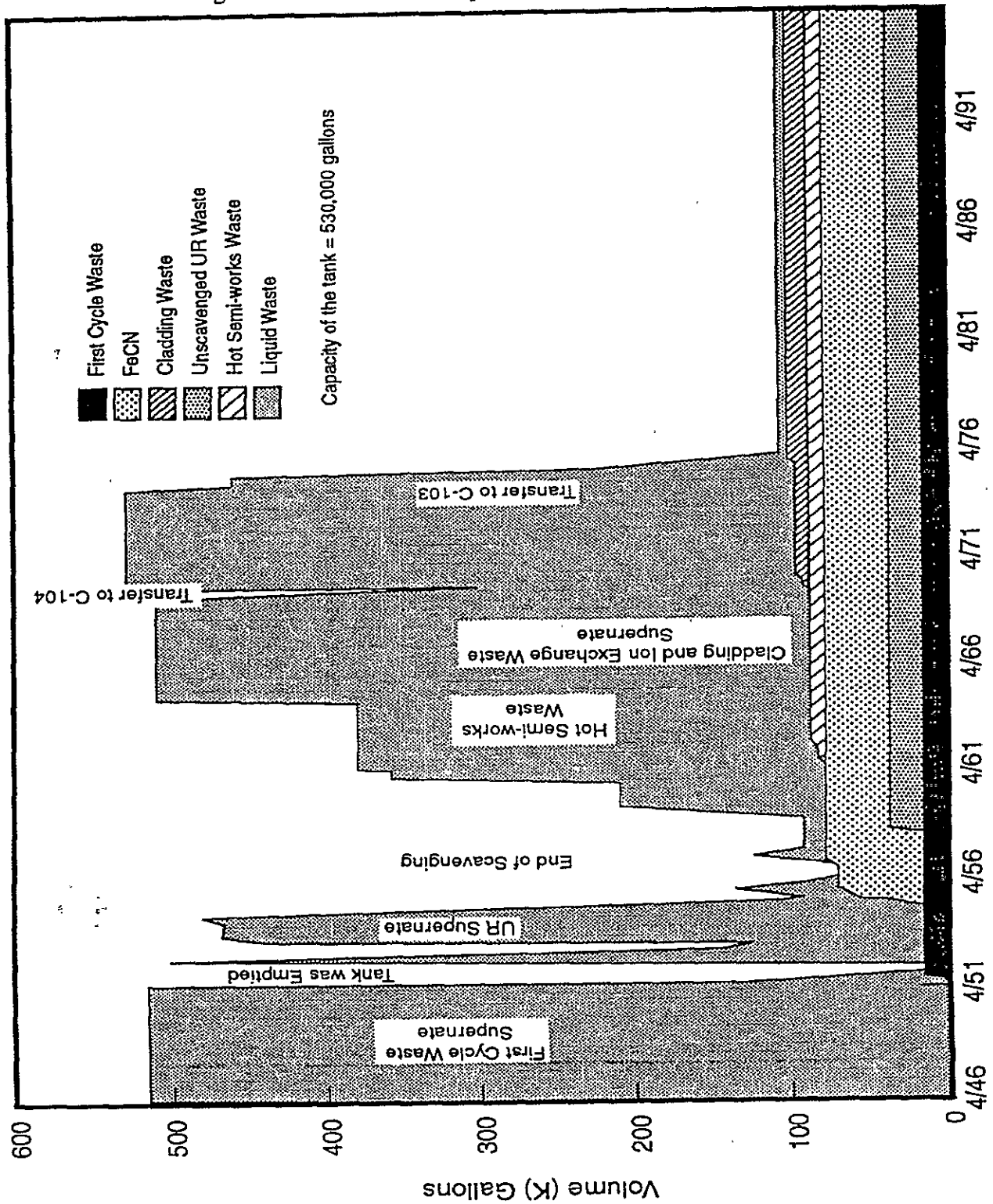
A later General Accounting Office (GAO) study (Peach 1990) postulated greater "worst-case" accident consequences, with independently calculated doses one to two orders of magnitude greater than the HDW-EIS. A special Hanford Site Ferrocyanide Task Team was commissioned in September 1990 to address all issues involving the ferrocyanide tanks, including the consequences of a potential accident. On October 9, 1990, the Secretary of Energy announced that a supplemental environmental impact statement would be prepared containing an updated analysis of safety issues for the Hanford Site SSTs, including a hypothetical ferrocyanide explosion. In October 1990, the ferrocyanide issue was also declared an USQ because the consequences of the accident scenario (as calculated by the GAO) were outside the bounds of the current safety analyses for SSTs. Furthermore, additional monitoring of tanks with designated USQs was mandated by Public Law 101-510 (1990).

Using a computer model output (Jungfleisch 1984), process knowledge, and transfer records, 24 waste tanks have been identified at the Hanford Site as potentially containing 1,000 g-mol (465 lb) or more of ferrocyanide as the $Fe(CN)_6^{-4}$ ion. On further investigation, six of these tanks are believed to have received less than 1,000 g-mol of ferrocyanide sludge and are therefore candidates for removal from the Watch List (Cash 1993). Tank 241-C-112 is on the Ferrocyanide Watch List because it was a known process tank during the ferrocyanide-scavenging campaigns.

2.2 EXPECTED TANK CONTENTS/CONDITIONS

Process knowledge obtained from historical records and waste simulants produced from the scavenging process flowsheets can be used to predict the major constituents and some general physical properties of the waste matrix in the waste tanks. Initially, the differences between the U Plant and In-Farm ferrocyanide sludges were not fully appreciated. However, further investigation of the simulants showed that the In-Farm process would be expected to precipitate approximately 1.0 to 1.3 vol% solids, and thus the sludge would have been deposited in the receiver tanks in layers approximately 3.6 to 6.1 cm (1.4 to 2.4 in.) thick. This is much less than the 4.25 vol% and 15- to 20-cm (6- to 8-in.) layers expected from the U Plant material. The In-Farm scavenged ferrocyanide tanks (such as tank 241-C-112) are expected to contain relatively soft sludge, which can be push-mode sampled. This expectation was supported by inspection of in-tank photographs. The other waste solids that were added to the tank after the scavenging campaign are also expected to be soft. During its operating history, tank 241-C-112 was never subject to any of the various in-tank solidification processes; consequently, there was no formation of hard salt cake on top of the sludge (as there was in the BY Tank Farm). No records of any previous core sampling have been found.

The most recent waste inventory measurement for tank 241-C-112 reports 394,000 L (104,000 gal) of waste with an estimated 121,000 L (~ 32,000 gal) of drainable liquids (Hanlon 1992). These figures translate to a waste depth of 115.1 cm (45.3 in.) at the tank centerline. Because the tank had less than 189,000 L (50,000 gal) of drainable liquid, it was

administratively interim-stabilized in September 1990, and is considered sound. Tank Farm Operations has installed a second thermocouple tree in tank 241-C-112, and the readings between the two thermocouple trees on opposite sides of the tank are consistent. The present maximum waste temperature in tank 241-C-112 is ~29 °C (85 °F), and the estimated heat load in the tank is less than 2.93 kW (10,000 Btu/hr). Tank 241-C-112 is considered to have one of the highest ferrocyanide concentrations of all the ferrocyanide Watch List SSTs (Borsheim and Simpson 1991).

In summary, various nickel ferrocyanide complexes (primarily disodium) are expected to be mixed with an interstitial solution containing sodium nitrate and nitrite. Cesium-137 is expected to be present as a mixed salt (possibly as $NaCsNiFe(CN)_6$); strontium-90 may be in several potential compounds: phosphate, sulfate, or carbonate. Both of these radionuclides have decayed through slightly more than one half-life, and therefore are not as abundant as when the scavenging waste was originally deposited. Other fission products with relatively short half-lives (such as ^{60}Co and [106]Ru) are not expected to be in abundance, especially with the limited number of waste additions for this tank. Hydrated transition metal oxides/hydroxides (including small amounts of transuranics) are also expected because of alkaline conditions. Other ions expected to be present are potassium, calcium, aluminum, and uranium. The supernate and interstitial liquid is expected to contain large amounts of sodium, nitrate, and nitrite ions. Table 2-1 presents estimates of some typical characteristic analyte concentrations in the waste streams disposed of in 241-C-112.

Table 2-1. Expected Concentrations for Characteristic Analytes (Schneider 1951; Jeppson and Wong 1993).

Waste Type	Identifier Analytes	Typical Process Stream Concentration (μg/g wet solids)
First Cycle Decontamination (1C)	Bi	7,100
	F	9,600
	U	620
	Si	2,400
Unscavenged Uranium Recovery (UR)	U.	1,700 - 32,500
Ferrocyanide (FeCN)	Ni	18,700
	Ca	30,000
	CN	91,000 - 113,000
Decladding Waste (CW)	Al	18,700
	U	340
	Si	2,600
Hot Semiworks (HS)	[90]Sr	No data, expected elevated ^{90}Sr levels

2.3 ANALYTICAL RESULTS FROM SIMULANT STUDIES

Physical and chemical measurements performed on simulants of ferrocyanide tank waste provide additional information and perspective regarding the condition and properties of the waste in tank 241-C-112.

2.3.1 Simulant Formulation: In-Farm 2 Flowsheet Material

The In-Farm 2 flowsheet material is considered to be an energetically conservative but reasonably close physical and chemical analogue of the ferrocyanide precipitate in tank 241-C-112 as it was deposited in the tank during the scavenging campaign. The In-Farm 2 flowsheet materials were prepared according to the following instructions (Jeppson and Wong 1993). The feed solution composition is listed in Table 2-2. Deionized water was used for feed solution and chemical addition makeup.

Table 2-2. Feed Solution Composition for
In-Farm 2 Flowsheet.

Component	Concentration (mol/L)
Sodium Nitrate (NaNO$_3$)	3.75 $\underline{M}$
Cesium Nitrate (CsNO$_3$)	0.00025 $\underline{M}$
Sodium Nitrite (NaNO$_2$)	1.25 $\underline{M}$
Sodium Sulfate (Na$_2$SO$_4$)	0.17 $\underline{M}$
Sodium Phosphate (Na$_3$PO$_4$)	0.16 $\underline{M}$

The product sludge was the precipitate produced when performing the following steps. This procedure mimicked the actual In-Farm 2 process that is illustrated in Figure 2-2. The feed solution was heated to 40 °C and the pH adjusted to 9.1 $\pm$ 0.5. The sodium ferrocyanide was then added to the solution, followed by nickel sulfate. The simulant solution was agitated for 1 hour, then struck with calcium nitrate. After the addition of calcium nitrate, the solution was agitated for another hour and allowed to settle. The settling was done for eight days and the supernate was decanted. The remaining sludge was centrifuged at 2,100 g for 14 hours and 1,820 g for 7 days in an attempt to simulate 3.6 and 30 gravity-years of settling respectively (Jeppson and Wong 1993). Selected physical properties for the two settled centrifuged sludges are presented in Table 2-3. Table 2-4 presents an estimate of the chemical composition of the In-Farm 2 simulant.

2.3.2 Simulant Physical Characteristics (see Jeppson and Wong 1993)

Table 2-3. Summary of In-Farm 2 Simulant Characterization Data.

	Property	In-Farm 2 sludge
30 gravity year	Water content, sludge	51 wt%
	pH, supernatant	9.42
	Bulk density, sludge	1.39 g/mL
3.6 gravity year	Bulk density, supernate	1.27 g/mL
	Particle density (dried sludge)	2.38 g/mL
	Particle size distribution, (by number)	97% < 2 μm Median diameter**: 0.76, 0.76 μm Acquisition Range: 0.5-150 μm
	Particle size distribution, (by volume)	100% < 110 μm Median diameter**: 14.3, 16.8 μm Acquisition range: 0.5-150 μm
	Hydraulic conductivity (permeability)	4.0 x 10^{-7} cm/s
	Total porosity	67.9%
	Thermal conductivity	1.82 W/m•K @ 39 °C 2.16 W/m•K @ 56 °C 2.82 W/m•K @ 68 °C 2.04 W/m•K @ 72 °C*

*Jeppson and Wong (1993) noted an anomalous data point, but were unable to explain the inconsistency of the observation.

**Two separate measurements.

2.3.3 Chemical Composition (see Jeppson and Wong 1993)

Table 2-4. Estimated Composition of Homogenized, Centrifuged,
Ferrocyanide Sludge Simulant.

Sludge component	Average In-Farm 2 weight fraction
Disodium mononickel ferrocyanide: $Na_2NiFe(CN)_6$	0.101
Sodium nitrate: $NaNO_3$	0.174
Sodium nitrite: $NaNO_2$	0.051
Sodium hydrogen phosphate: Na_2HPO_4	0.016
Sodium sulfate: Na_2SO_4	0.015
Calcium phosphate: $Ca_3(PO_4)_2$	0.073
Water	0.51
Percent mass balance subtotal	94.0
Percent unknown--likely includes $Fe_4(Fe(CN)_6)_3$, $Fe(OH)_3$, $Ni(OH)_2$, and other materials from trace impurities	6.0

2.3.4 Energetics Behavior of Ferrocyanide Sludge Simulant

Available chemical process information indicates that there were three significantly different types of ferrocyanide waste (Sloat 1954; Schmidt and Stedwell 1954). Nonradioactive waste simulants have been developed and tested using this information. In-Farm ferrocyanide waste, accounting for 20 to 25% of the total ferrocyanide waste, was formed from treating waste that was already stored in the tanks. The waste in tank 241-C-112 was produced using the In-Farm process. Most of this waste had less inert solids in the waste stream; therefore, it is believed to have been more concentrated in ferrocyanide than other ferrocyanide wastes. In-Farm simulants exhibit propagating exothermic activity when examined by differential and adiabatic scanning calorimetry (DSC and ASC).

Estimates of tank waste reactivity, which were developed after the ferrocyanide USQ was declared, were based on thermodynamic estimates (Colby and Crippen 1991). Several chemical reaction pathways were evaluated and heats of reaction were determined for each possible reaction from the published heats of formation of the reactants and the products. For the purpose of developing these estimates, the condition of the reactants are dry solid reagents at standard temperature and pressure in a stoichiometric ratio. The theoretical heats of reaction ranged in value from $\Delta H = -9.6$ kJ/g to $\Delta H = +19.7$ kJ/g of $Na_2NiFe(CN)_6$, and are listed below with their corresponding chemical reactions.

(1) $Na_2NiFe(CN)_6 + 54NaNO_3 + 22H_2O \longrightarrow 6Na_2CO_3 + FeO + NiO + 60NO_2 + 44NaOH$
$\Delta H = +19.7$ kJ/g of $Na_2NiFe(CN)_6$

(2) $Na_2NiFe(CN)_6 + 14NaNO_3 + 2H_2O$ ----> $6Na_2CO_3 + FeO + NiO + 20NO + 4NaOH$
$\Delta H = -0.7$ kJ/g of $Na_2NiFe(CN)_6$

(3) $Na_2NiFe(CN)_6 + 9NaNO_3$ ----> $5.5Na_2CO_3 + FeO + NiO + 7.5N_2O + 0.5CO_2$
$\Delta H = -6.8$ kJ/g of $Na_2NiFe(CN)_6$

(4) $Na_2NiFe(CN)_6 + 10NaNO_3$ ----> $6Na_2CO_3 + FeO + NiO + 6N_2O + 4NO$
$\Delta H = -5.7$ kJ/g of $Na_2NiFe(CN)_6$

(5) $Na_2NiFe(CN)_6 + 6NaNO_3$ ----> $4Na_2CO_3 + FeO + NiO + 6N_2 + 2CO_2$
$\Delta H = -9.6$ kJ/g of $Na_2NiFe(CN)_6$
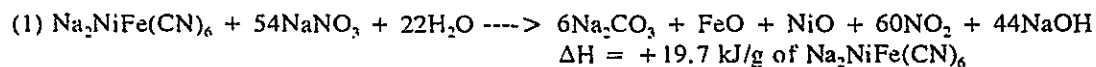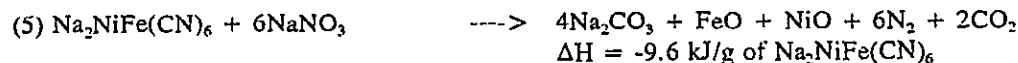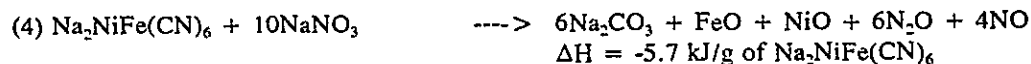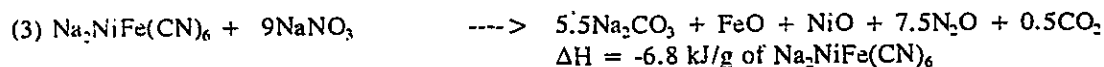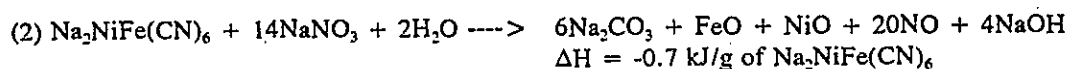
At temperatures below 1700 °C (3100 °F), the carbonate product is thermodynamically favorable and should predominate (Scheele et al. 1991). Note that considerably lower energy releases are obtained if the reaction is incomplete or if NO or NO_2 is formed rather than N_2 or N_2O. A three-component diagram illustrating the exothermic potential of various mixtures of ferrocyanide, nitrate, and inerts is presented in Figure 2-4. Further detail regarding the thermodynamic estimates of these mixtures is presented in Colby and Crippen (1991).

The waste simulants prepared using the In-Farm and U Plant process flowsheets were tested for chemical activity (Fauste 1992). Chemical and physical analyses of the In-Farm and U Plant waste simulants show that they contain an average of 51 and 66 wt% water, respectively, after centrifugation. The centrifugation was done to represent 30 gravity-years of compaction that may have occurred during storage. This amount of water in the waste matrix presents a tremendous heat sink that must be overcome before any reactions can become self-sustaining. During the DSC examinations, the samples exhibited large endotherms between room temperature and 150 °C (Jeppson and Wong 1993). Results from thermogravimetric analyses being run at the same time showed a large loss of mass (i.e., evaporation of water) in this same temperature range; thus, reactions were only able to occur in dry or nearly dry sample material. Average ferrocyanide content of the In-Farm 2 waste simulants is approximately 10.1 wet wt% (20.6 wt% dry). Table 2-5 presents the ΔH found for some simulant materials.

Table 2-5. Heats of Reaction of Various Simulants.

Material	ΔH (From Adiabatic Calorimetry)	Wt% Ferrocyanide (dry) [$Na_2NiFe(CN)_6$]	Calculated ΔH per gram $Na_2NiFe(CN)_6$
U Plant 1 simulant	-0.17 kJ/g of dry material	4.3	-3.95 kJ/g
U Plant 2 simulant (Bottom fraction)	-0.34 kJ/g of dry material	8.6	-3.95 kJ/g
In-Farm 1 simulant (Bottom Fraction)	-1.20 kJ/g of dry material	25.5	-4.71 kJ/g

The onset temperatures for propagating reactions to take place in the simulants range from 244 °C to 278 °C (471 to 532 °F). However, Arrhenius-type reactions may occur at lower temperatures (Fauske 1992).

NOTE: 4.18 J = 1 cal.

Figure 2-4. Ferrocyanide Tank 3-Component Diagram.

FERROCYANIDE TANK 3-COMPONENT DIAGRAM
REACTION PROPAGATION VS. CONCENTRATION

Note: Weight % Basis
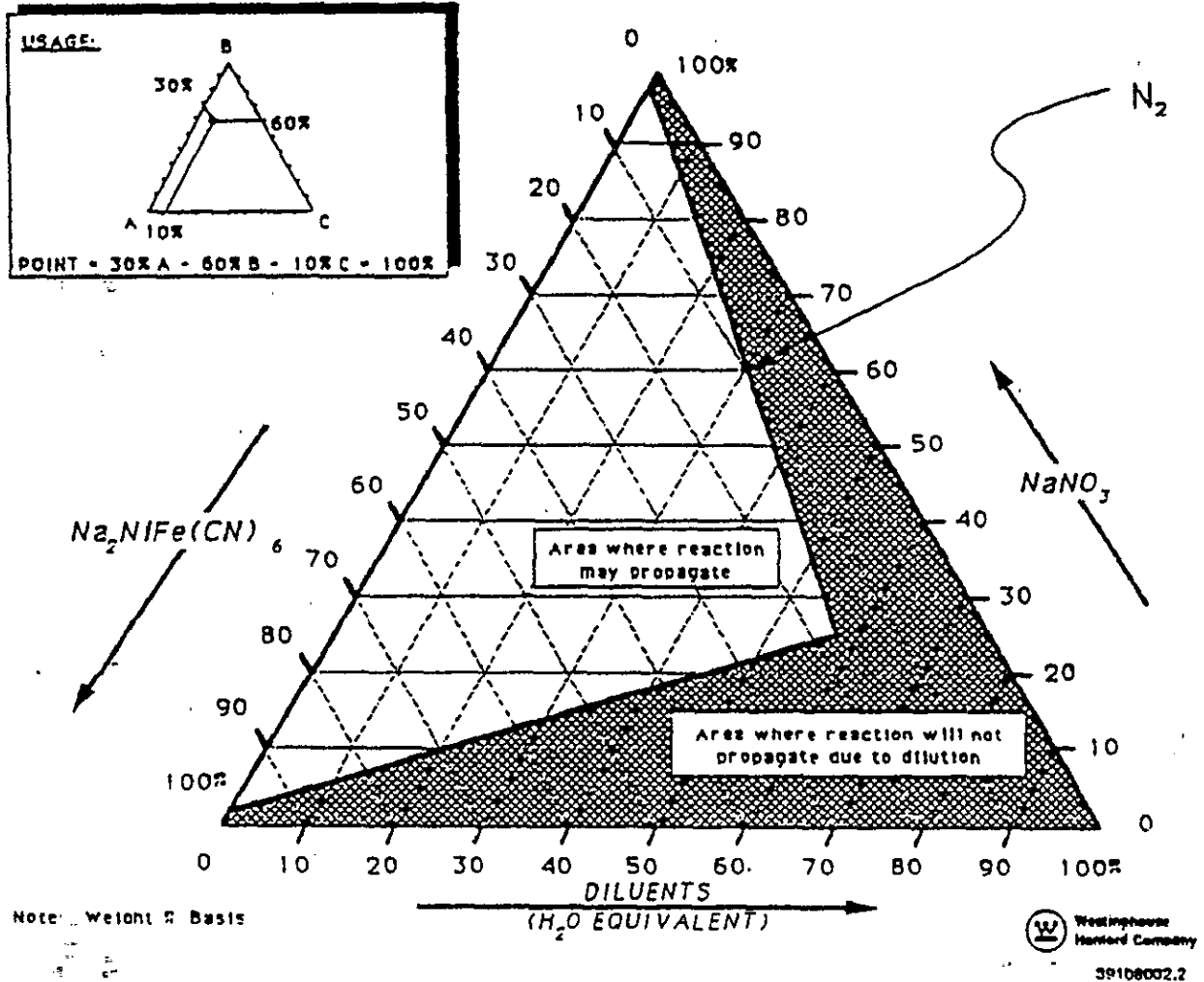
3.0 CORE SAMPLING

3.1 DESCRIPTION OF SAMPLING EVENT

Tank 241-C-112 was push-mode core sampled through three risers from March 19, 1992, to March 26, 1992. Two segments were expected from each core sample. Core 34 was obtained from riser #2; core 35 was obtained from riser #7; core 36 was obtained from riser #8. The core samples from tank 241-C-112 were obtained using a specially designed core sampling truck (CST). The sampling equipment is mounted on a rotating platform on the CST. Access to the interior of the tank is provided by various tank risers. These risers are pipes of various diameters leading into the tank dome from the ground. The riser configuration for tank 241-C-112 is given in Figure 3-1. A review of the tank farm operating records and a field inspection of the tank risers determine which risers can be used in the sampling operation. A riser is opened and the CST is positioned over the riser. The sampler is lowered into the tank through the drill string and pushed into the waste.

The sampler is constructed of stainless steel and is 48 cm (19 in.) long, with a 2.2-cm (7/8-in.) inside diameter, and has a volume of 187 mL (0.05 gal). Tank Farm Operations has determined that sampling events of one or two segments do not require hydrostatic head balance fluid. Therefore, none was used in this operation, which eliminated any potential problems with sample contamination. When a segment is captured by the sampler, it is sealed within a stainless steel liner, and the liner is placed within a shipping cask. The shipping casks are approximately 122 cm tall, 13 cm in diameter, and have 2.5 cm of lead shielding. This degree of shielding and containment protects workers from excessive radiological exposure and prevents any liquids from the sample (or the sample itself) from being lost.

The casks were transported to the 324 Shielded Materials Facility for gamma scanning, and then to the 325 Analytical Chemistry Laboratory for characterization analysis. Both facilities are operated by Battelle-Pacific Northwest National Laboratory (PNNL) in the 300 Area of the Hanford Site. Cores 34 and 35 arrived at the 324 Facility on March 25, 1992, and core 36 arrived on March 26, 1992.

3.2 CHAIN-OF-CUSTODY

A chain-of-custody record was kept during the sampling event for each segment that was sampled. The chain-of-custody form is a one-page record that is used to ensure that (1) the sample is safely and properly transported from the field to the laboratory, and (2) the correct personnel are involved in the sampling operation and transportation of the sample to the laboratory. Tables 3-1, 3-2, and 3-3 summarize the most important data contained in the field chain-of-custody.

Figure 3-1. Tank 241-C-112 Riser Configuration.

Not to Scale
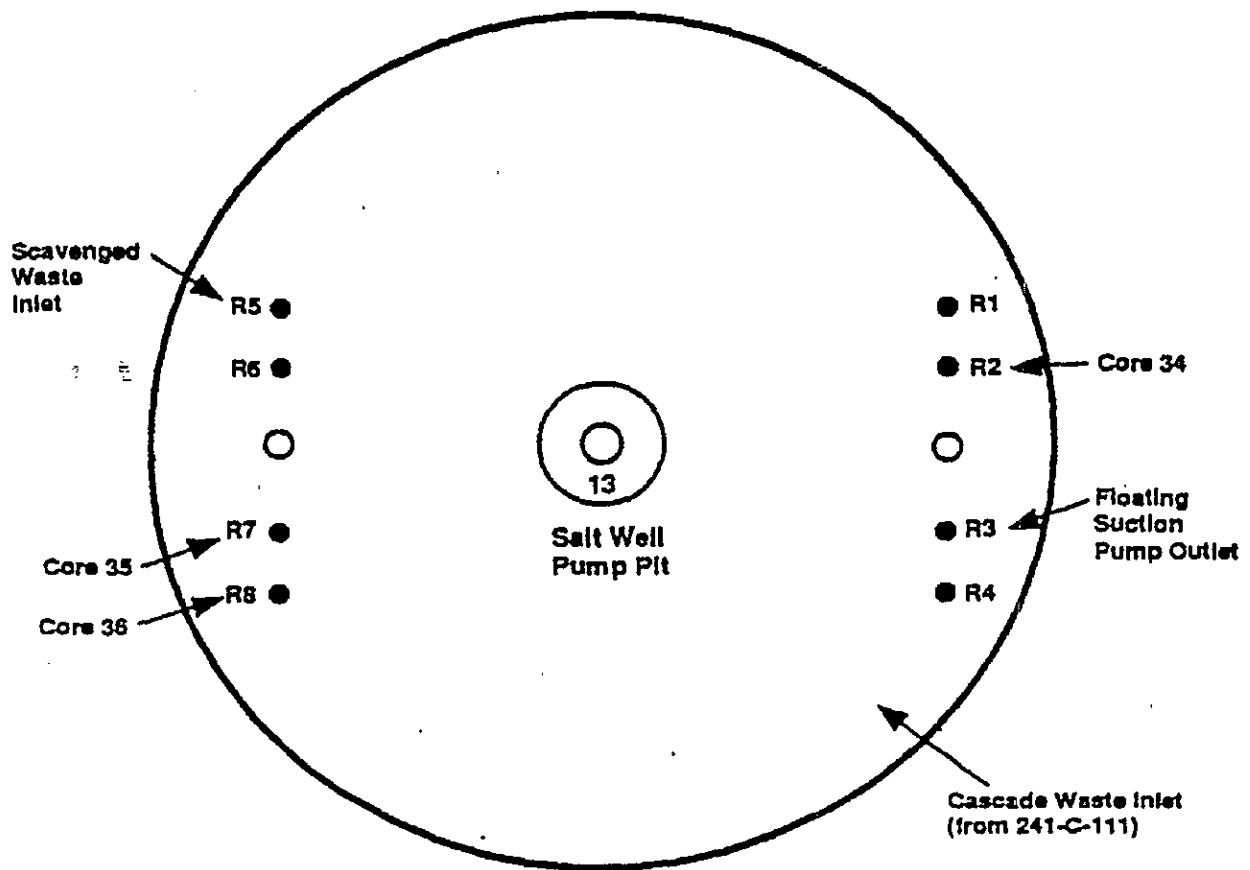
Tank 241-C-112 was constructed in 1943-1944;
it has a diameter of 22.9 m (75 feet) and nominal capacity of
2 million liters (530,000 gallons).

29304019.6

Table 3-1. Core 34--Chain-of-Custody Summary.

Sample	Core 34: 92-001	Core 34: 92-002
Place Taken	241-C-112 Riser 2	241-C-112 Riser 2
Date Taken	3/19/92	3/19/92
Date Released	3/25/92	3/25/92
Time Released	19:10	19:10
Sender	D. C. Hartley	D. C. Hartley
Receiver	M. R. Zumhoff	M. R. Zumhoff
Place Received	327 Building	327 Building
Time Received	20:55	20:55
Smearable Contamination	< DL alpha < DL beta-gamma	< DL alpha < DL beta-gamma
Dose Rate Through the Drill String	150 mR/hr	2 R/hr

Table 3-2. Core 35--Chain-of-Custody Summary.

Sample	Core 35: 92-003	Core 35: 92-004
Place Taken	241-C-112 Riser 7	241-C-112 Riser 7
Date Taken	3/22/92	3/22/92
Date Released	3/25/92	3/25/92
Time Released	19:10	19:05
Sender	D. C. Hartley	D. C. Hartley
Receiver	M. R. Zumhoff	M. R. Zumhoff
Place Received	327 Building	327 Building
Time Received	20:55	20:55
Smearable Contamination	< DL alpha < DL beta-gamma	< DL alpha < DL beta-gamma
Dose Rate Through the Drill String	1 mR/hr	500 mR/hr

Table 3-3. Core 36--Chain-of-Custody Summary.

Sample	Core 36: 92-005	Core 36: 92-006
Place Taken	241-C-112 Riser 8	241-C-112 Riser 8
Date Taken	3/24/92	3/24/92
Date Released	3/26/92	3/26/92
Time Released	19:05	19:05
Sender	D. C. Hartley	D. C. Hartley
Receiver	M. R. Zumhoff	M. R. Zumhoff
Place Received	327 Building	327 Building
Time Received	20:55	20:55
Smearable Contamination	< DL alpha < DL beta-gamma	< DL alpha < DL beta-gamma
Dose Rate Through the Drill String	2 R/hr	2 R/hr

A primary function of the chain-of-custody record is to provide radiation survey data. This is a record of the radiation dose that is emitted from the shipping cask. The dose rates in mrem/hour are measured from the top, sides, and bottom of the cask. These values are recorded on the chain-of-custody and represent the radiation being emitted directly from the sample. The last item recorded under the radiation survey data is the smearable contamination. Smearable contamination represents the radiation from waste material that is not sealed within the shipping cask; values greater than 100 mrem/hour are considered unsafe. Measurements are made both in the field and in the laboratory. No smearable contamination was found with these samples.

The chain-of-custody has several other important functions: (1) to provide a brief description of the cask, sampler, and the expected contents of the sampler (shipment, sample, and cask serial numbers for the specific sampling event); (2) to provide summary information about the analytical suite that the sample will undergo or reference the salient documentation; (3) to provide traceability for the sample during transport; and (4) to ensure sample integrity on arrival at the laboratory. This information is provided to ensure that each sample can be uniquely identified.

Copies of the chain-of-custody forms are on file at the Hanford Analytical Services Management (HASM) office. From inspection of the chain-of-custody records, there appear to be no irregularities in the sampling or transport of tank 241-C-112 samples from the field that would merit a safety or sample integrity concern (i.e., sample containment was not breached).

4.0 SAMPLE PREPARATION/SAMPLE EXTRUSION

4.1 DESCRIPTION OF GAMMA SCANNING EFFORT

The 324 Shielded Materials Facility has a device that was used for radionuclide measurements of irradiated fuel rods for the Fast Flux Test Facility program. The fuel pin reader was idle and there was the belief that this effort could be an efficient and innovative utilization of available resources in the Tank Waste Characterization/Waste Tank Safety Programs. Cores 34, 35, and 36 were transferred to the Shielded Materials Facility, and gamma scanning was performed on tank 241-C-112 sample segments using this device. Nine isotopes were scanned for: ^{137}Cs, ^{155}Eu, ^{154}Eu, ^{241}Am, ^{144}Ce, ^{134}Cs, ^{60}Co, ^{106}Ru, and ^{153}Gd. Of these, only the ^{137}Cs isotope had sufficient activity to be considered valid. The gamma scanning effort was undertaken as a means to obtain core sample information prior to extrusion. The data presented in Figures 4-1, 4-2, and 4-3 have been smoothed in a 5-point rolling average, to eliminate any anomalous peaks and to aid in interpretation of the scan.

Core 34

Core 34 was scanned in 2.5-mm (0.1-in.) increments at 500 seconds per increment on April 4, 1992. The scan was performed with the core in a vertical orientation, and the scan was begun at an axial location below the bottom of the segments (refer to Figure 4-1). Peak count rates for ^{137}Cs in core 34 were 2.72 counts/sec (1,361 counts in 500 seconds). Sample lengths were estimated from the activity signatures to be 8.6 cm (3.4 in.) for segment 92-001 and 36.3 cm (14.3 in.) for segment 92-002. These sample lengths were determined to be relatively reliable for the sample solids upon extrusion. However, there remained considerable uncertainty regarding sample recovery until the samples were extruded. In the case of segment 92-001, it appeared that suspended solids were concentrated at the bottom of the sampler while it was being scanned; thus, their activity signature masked the amount of cohesive solids in the sampler. After filtering, the liquids were found to have very little activity associated with them.

Core 35

The scanning effort for core 35 was subject to multiple mechanical difficulties. An anomalous peak was detected initially in the first scan on May 8, 1992. This result was suspicious and a rescan was recommended to confirm the result. The second scan performed on May 15 to verify this finding showed an abnormally low count rate (differing by an order of magnitude with core 34). The peak from the first scan was attributed to a power surge and shutdown of the device; the low count rate from the second scan was determined to be a calibration error. A third scan was performed on May 22 with no problems (refer to Figure 4-2). Core 35 was scanned the first two times in 2.5-mm (0.1-in.) increments at 500 seconds per increment; the third scan was done at a resolution of 1.27 mm (0.05 in.) and 500 seconds per increment. The scan was performed with the core in a vertical orientation, and the scan was begun at an axial location below the bottom of the segments.

Figure 4-1. 241-C-112 Gamma Scanning, Core 34 Results.

Gamma can Data -- SST C-112
Core 34 -- Segments 1 & 2 - Riser 2

^{137}Cs Profile
5-Point Rolling Average

Segment 1

Segment 2

Est. Hgt. Above Bottom at Sidewall (in)

Cs137 Counts/Second

Received 25-Mar-92
Scanned 4-Apr-92
500 Seconds/0.1 Inch

Revised June 16, 1992 HES

Gamma Scan Data — SST C-112
Core 35 — Segments 3 & 4 — Riser 7

137Cs Profile

10-Point Rolling Average (Seg 4B)
5-Point Rolling Average (Seg 3)

Segment 3

Segment 4B

NOTE: Segment 4B is a re-scan of Segment 4
to resolve discrepancies in original scan

Received 25-Mar-92
Scanned 8/12/22-May-92
500 sec/0.05 inch (Seg 4B)
500 sec/0.1 inch (Seg 3)

Cs137 Counts/Second

Est. Hgt. Above Bottom at Sidewall (in)

Revised June 2, 1992 HES

Figure 4-2. 241-C-112 Gamma Scanning, Core 35 Results.

Gamma Scan Data – SSi C-112

Core 36 – Segments 5 & 6 – Riser 8

137
- Cs Profile
5-Point Rolling Average

Segment 5

Segment 6

Est. Hgt. Above Bottom at Sidewall (in)

Cs137 Counts/Second

Received 25-Mar-92
Scanned 16-Apr-92
1275 seconds/0.1 In.

Revised June 16, 1992 HES
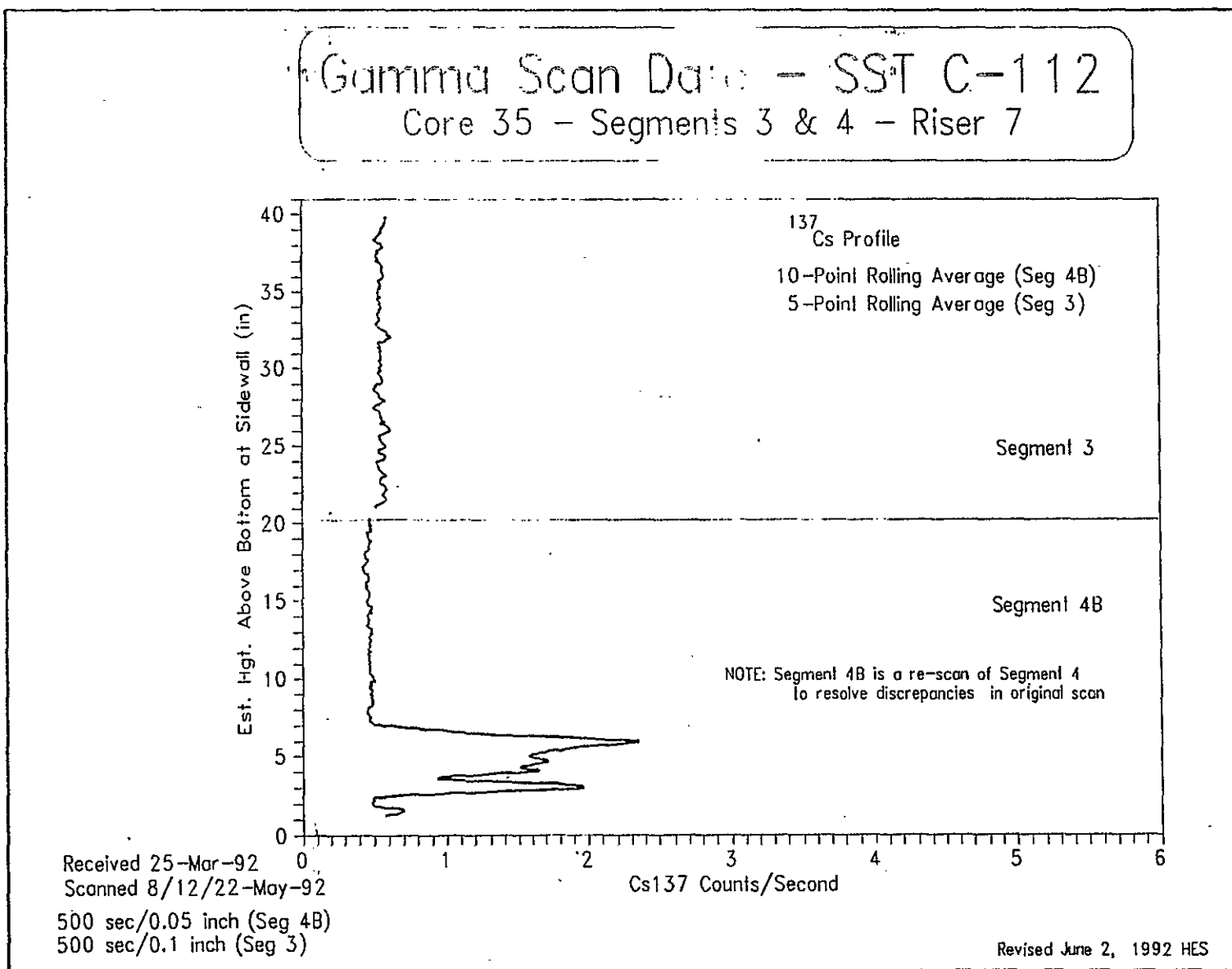
Figure 4-3. 241-C-112 Gamma Scanning, Core 36 Results.
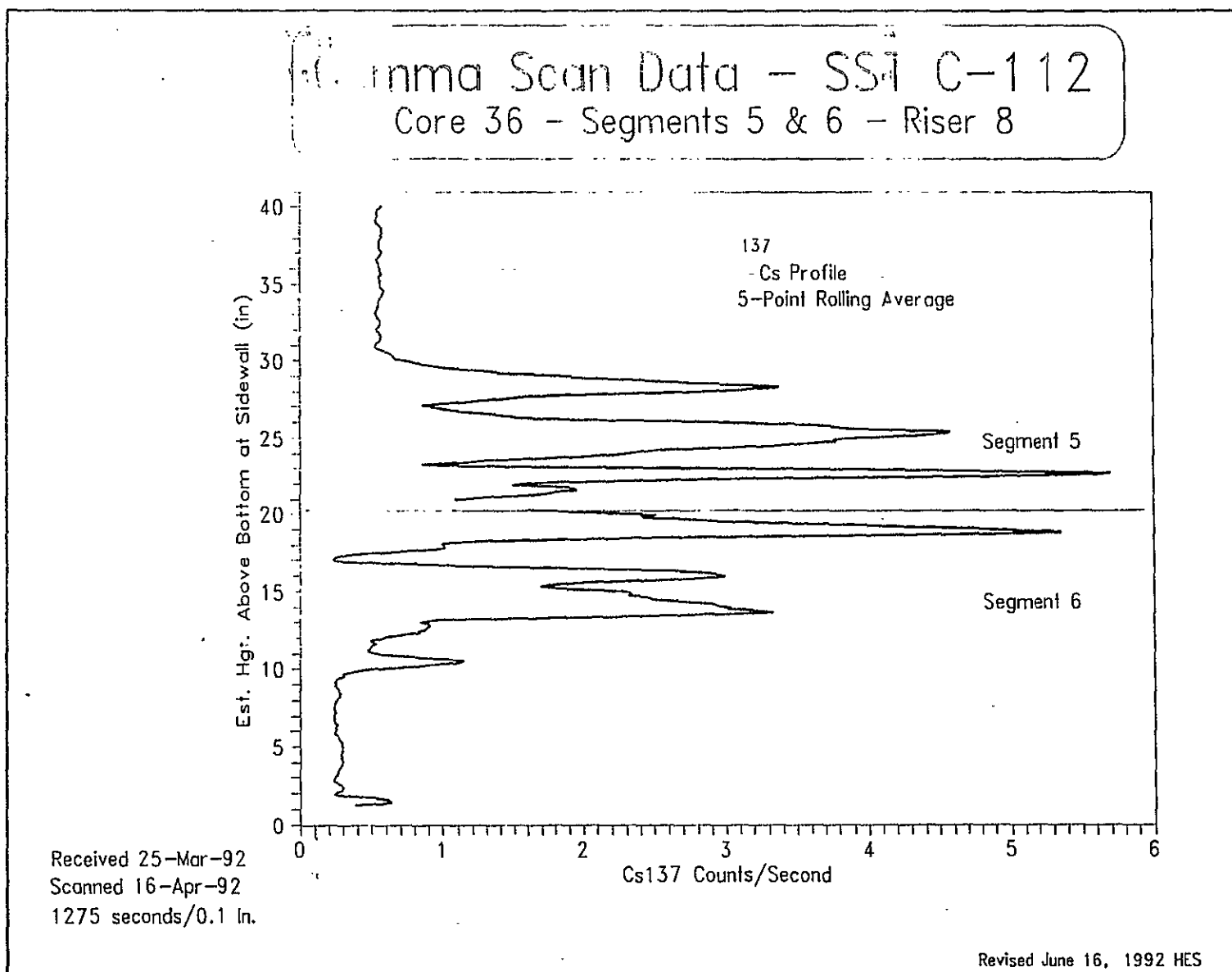
The sample valve was found open on segment 92-003 and the sampler was presumed empty; however, the sampler was scanned to determine if there was any contamination from contact with the tank contents. The gamma scanning measurements indicated no activity, and the sampler was empty upon extrusion. Peak count rates for ^{137}Cs in core 35, segment 92-004 was 2.83 counts/second (1,415 counts in 500 seconds; comparable with core 34). Sample length was estimated to be 10.2 cm (4 in.) for segment 92-004. Sample length on extrusion was found to be approximately 7.6 cm (3 in.) and, as with core 34, the liquids found had almost no activity.

Core 36

Core 36 was scanned in a slightly different manner than the other two. It was scanned in 2.5-mm increments and 1,275 seconds per increment on April 16, 1992 (refer to Figure 4-3). This was done to determine if the longer viewing time improved the detection and resolution of the radioisotopes. There was no noticeable improvement in the sensitivity of the device using the longer counting times. Peak count rates for ^{137}Cs in core 36 were 6.61 counts per second (8,430 counts in 1,275 seconds). This core sample had a much higher overall activity than the other two, better sample recovery, and a much more distinctive signature. Sample lengths were estimated to be 22.9 cm (9 in.) for segment 92-005 and 47 cm (18.5 in.) for segment 92-006. Sample length on extrusion was found to be approximately 21.8 cm (8.6 in.) and 43.9 cm (17.3 in.), respectively. There were no liquids found with these segments.

4.1.1 Tank 241-C-112 Gamma Scanning Summary

Gamma scans were performed as a scoping procedure on cores 34, 35, and 36 to obtain a qualitative measurement of the activity of the waste and to identify the major contributors. The gamma activity pattern obtained from the scans indicated some gross layering of the waste in the tank (i.e., differences in waste types) as well as differences in activity between individual batches. Of the nine isotopes scanned for, no significant gamma emitters were found in the tank waste except ^{137}Cs, although the sample had a relatively high gamma background. The activity of tank 241-C-112 waste material ranged between 0.15 and 2 R/hour, as measured through the drill string. No significant radiological activity was found in the drainable liquid in the tank. The ^{137}Cs appeared to be almost entirely associated with the solids, and thus was assumed not soluble.

Each of the three cores in tank 241-C-112 was expected to contain one full segment and a partial (3/4) second segment. This expectation was based on calculations based on the inventory values given in Hanlon (1992). Results from the gamma scans indicated there was less than anticipated recovery. At that time, alteration of the analysis plan was necessary because there was not enough sample to perform all of the requested analyses.

4.2 DESCRIPTION OF SAMPLE HANDLING AND ANALYTICAL SCHEME

4.2.1 Sample Breakdown Procedure

Because tank 241-C-112 has been identified as a Watch List tank, as detailed in Section 2.1.2, more extensive analytical measurements are required to resolve the safety concerns associated with this tank. To enhance the resolution of the assays for key analytes, the analysis horizon for characterization was determined to be one-quarter of a segment.

The sampler was removed from the shipping cask directly into the hot cell. At this time, the sampler must be placed into the horizontal position; hence, any free liquid at the top of the sampler has an opportunity to drain to the liner. The sample was then loaded into the mechanical extruder and removed by pushing it out from the back of the sampler with a piston. In this case, the sampler is pressed against a fixed piston, forcing the sample into the extrusion tray. If a full sample has been captured, the material nearest the valve was from a deeper part of the tank; the material near the piston was closer to the surface. The sample and any liquids were collected on a metal tray. Next, the mass of the segment and the approximate length were recorded. From this information, the bulk densities of the segments can be estimated. The sample volume is determined by measuring the length of the extruded sample using a linear unit volume of 9.85 mL/in. Each segment was divided into 12-cm (4-3/4-in.) subsegments. Figure 4-4 illustrates how the ferrocyanide SST segment sample was extruded and divided into subsegments. A video record of the extrusions of each of the segments from tank 241-C-112 was made, and color photographs documenting the extruded segments were taken.

Figure 4-4. Typical Single-Shell Tank Segment Extrusion.
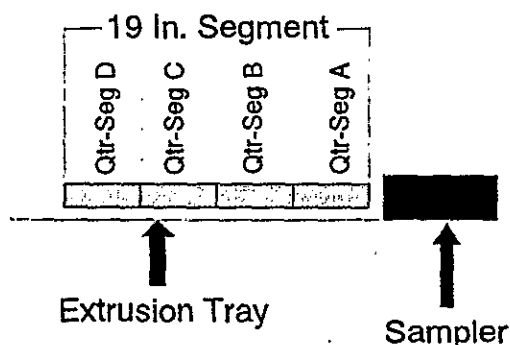
Several different styles of nomenclature are used for distinguishing core samples, sample segments, and subsegments in the existing literature. Two major conventions are used in the documentation relating to ferrocyanide (and core sampling in general). The first is designating the segment with the last two digits of the calendar year (92-) and then

numbering the segments sequentially (-001, -002, etc.). This system resets itself every calendar year. The second system distinguishes the tank, core, segment, and subsegment. The first (bottom) 12 cm (4-3/4 in.) of the extruded sample is assigned to the fourth subsegment and is uniquely identified (Tank ID - Core No. - Segment No. - D). The following three 12-cm (4-3/4-in.) sections of the extruded segment are labelled as C, B, and A, respectively. An example of this naming protocol for the third subsegment from the second segment of the first core is (241-C-112-Core 34-Segment 2-B). If the extruded segment is less than 48 cm (19 in.) long, then the same naming convention applies until no solid material is left to make a complete 12-cm (4-3/4-in.) subsegment. The first 12 cm (4-3/4 in.) is be assigned to the D subsegment (etc.). This second system of naming is the primary convention used in this report. Where no tank identification is given in this report, it should be understood as meaning tank 241-C-112.

4.2.2 Homogenization Tests

The subsegment and core composite samples are homogenized using a mechanical mixer prior to analysis. This is done so that aliquots removed for analysis will be representative of the entire subsegment or core composite. Aliquots of the homogenized tank waste from core 34-2C and 2D, core 35-2D, and core 36-1C and 2D were taken to determine the efficacy of the homogenization procedure. The samples were split into duplicates, acid digested, and assayed by inductively coupled plasma - atomic emission spectroscopy (ICP) and gamma energy analysis (GEA). This procedure is done to determine if the degree of mixing achieved by the as-planned homogenization procedure was sufficient for the remaining samples to be homogenized and prepared for analysis. If the analytes from the aliquots are within a relative percent difference (RPD) of 10%, the samples are considered homogenized. If there are several analytes that are not within the specified RPD, the samples are mixed further and re-assayed. Once homogenization was indicated, the remaining samples were homogenized via the required procedure and prepared for analysis. The investigators reported that the samples from tank 241-C-112 exhibited substantial resistance to homogenization. Generally, the samples had to be blended twice before the ICP results were considered satisfactory. The GEA never showed satisfactory homogenization; it indicated the distribution of radionuclides remained irregular even after the second homogenization. However, this behavior was not unexpected because the simulant materials were very resistant to dissolution. The acid digestion preparation was probably insufficient to completely dissolve the sample, and the ^{137}Cs was associated with the insoluble materials. In the future, homogenization tests of suspected ferrocyanide tanks should use a potassium hydroxide (KOH) fusion sample preparation procedure because it provides more complete dissolution of the sample.

4.2.3 Subsegment-Level Analyses

The objectives of subsegment-level analyses are to provide (1) information as a function of depth pertaining to the overall waste energetics, (2) the distribution of ^{137}Cs and ^{90}Sr, (3) the concentration and solubility of the CN⁻ present in the sample, and (4) a higher resolution for determining bulk tank composition for certain analytes. To accomplish these goals, the limited suite of analyses listed in Table 4-1 was performed on each homogenized subsegment. These analyses were conducted using the analytical procedures identified in Tables I5-1 and I5-2 of WHC-EP-0210, Rev 3 (Hill et al. 1991), and as amended in Hill (1991). Brief descriptions of the sample preparation and assay methods are presented.

Table 4-1. Subsegment-Level Analysis.

Direct	Fusion Dissolution	Water Leach
TOC/TIC TGA DSC Total CN⁻ Wt% H_2O	ICP (Metals) GEA (^{137}Cs) ^{90}Sr	IC (Anions) CN⁻ pH GEA

DSC = Differential scanning calorimetry
GEA = Gamma energy analysis
ICP = Inductively coupled plasma - atomic emission spectroscopy (AES)
TGA = Thermogravimetric analysis
TIC = Total inorganic carbon
TOC = Total organic carbon.

Direct analyses are assays performed on the sample matrix with little or no sample preparation. Several direct analyses were performed relating to the energetic properties of the waste: total organic carbon (TOC), scanning thermogravimetric analysis (TGA), DSC, total cyanide, and gravimetric weight percent water.

The TOC was determined using the hot persulfate method. That method dissolves a sample in a sulfuric acid solution (90 °C+) to liberate inorganic carbon (carbonate). Potassium persulfate ($K_2S_2O_8$) is then added, and organic carbon is converted to CO_2, which is measured coulometrically. The difficulty encountered in solubilizing the sample matrix in the homogenization tests makes the results of this assay potentially unreliable.

Scanning TGA and DSC are useful in determining the thermal stability or reactivity of a material. TGA measures the mass of a sample while the temperature of the sample is increased at a constant rate. In DSC analysis, the heat absorbed/evolved over and above the usual heat capacity of the substance is measured while the substance is exposed to a linear increase in temperature.

Total cyanide analysis was done using a developmental procedure developed at PNNL. The sample was dissolved in a solution of ethylenediaminetetraacetic acid and ethylenediamine and placed in a microdistillation apparatus. The total cyanide content was determined by argentometric titration.

The gravimetric weight percent water was determined by drying the sample for 12 to 24 hours in an oven at 103 to 105 °C and measuring the difference in the weight of the sample.

Analyses that were performed on fusion-prepared samples were ICP and GEA for radionuclides. Fusion dissolution analyses are assays performed on the sample matrix after it has been fused with potassium hydroxide in a nickel crucible and dissolved in acid. This preparation dissolves the entire sample, whereas other sample preparation procedures may not completely dissolve the sample matrix. However, one significant disadvantage of fusion preparation is that large amounts of potassium hydroxide are required to bring a sample into solution. Because of this high dilution factor, trace elements are less likely to be correctly quantified, if they are detected at all. Elements that occur in abundance (major metals) or are highly insoluble are likely to be detected better by the fusion results than by any other sample preparation. Generally, fusion dissolution is the preferred method of analyzing radionuclide content, with the exception of ^{14}C and ^3H (tritium). However, the sample preparation specified in the test instructions for ^{14}C (water digestion) is likely not the best for the ferrocyanide waste. Difficulty with dissolving the sample with a water leach, and volatility associated with a fusion preparation, will bias the ^{14}C results low for both sample preparations. An adequate sample preparation method for ^{14}C is not available for this sample matrix; however, ^{14}C is not expected to be a significant contributor to the radionuclide content of the waste.

Water leach (or water digestion) analyses are assays performed after the sample matrix has been digested in distilled/deionized water; the water is then analyzed for soluble analytes. The soluble anions are determined by ion chromatography (IC). The primary anions analyzed in this manner are fluoride, chloride, nitrate, nitrite, phosphate, and sulfate. In addition, free cyanide and pH were also analyzed from water digestion samples.

4.2.4 Rheological and Physical Measurements

Only one 25-mL aliquot (from the second segment of core 36) was used for rheological and physical measurements. Viscosity, settling properties, fluid behavior, and shear strength were some of the primary characteristics investigated. The sample tested for these properties was not homogenized prior to analysis.

4.2.5 Subsegment Level Archive

Several analyses (adiabatic calorimetry, ferrocyanide speciation, and total oxygen demand [TOD]) have been identified by the Waste Tank Safety Programs as requiring developmental work. A sufficient amount of sample from each subsegment has been archived to perform these analyses when the procedures for these analyses have been developed. The adiabatic calorimetry assay will be performed on each subsegment if an exotherm of predetermined parameters is detected by DSC analysis. The boundaries for performing adiabatic calorimetry have been determined to be when the DSC exotherm is greater than -75 cal/g and the sample has 15 wt% water or less; or when the exotherm is greater than -125 cal/g, even if the sample has greater than 15 wt% water. Because of sample consumption constraints, the TOD test cannot be run for the subsegment from the rheology segment.

4.2.6 Core Composite Level Analysis

One composite from each core was built and analyzed in accordance with the complete baseline case core composite scenario detailed in Section 6.1 of WHC-EP-0210 (Hill et al. 1991) and as amended by Hill (1991). The type and number of analytical tests performed are similar to the suite done on the subsegments, but are much more extensive. The free liquid from the segments in core 34 was combined and analyzed as a separate liquid core composite. The free liquid from the segment in core 35 was also analyzed as a liquid core composite.

Selected radionuclides were measured on some of the water digestion samples to determine the type and number of water soluble radionuclides. ICP and atomic absorption (AA) spectroscopy were also performed on some of the water digestion samples. These assays were performed to determine the amount of soluble metal cations (ICP) or arsenic, mercury, or selenium (AA). In most cases, these analytes were below the detection limits in the water digestion samples, suggesting that most of the analytes are not water soluble

Acid digestion is a preparation method where the sample is dissolved in a mixture of nitric and hydrochloric acids. This preparation brings most of the insoluble metals into a solution with a minimum amount of dilution, and is usually best for detecting trace and some major metals. These properties are the reason that acid digestion is generally used as the sample preparation for the homogenization tests. The analyses performed on this preparation were the ICP, GEA, and AA analysis (the AA analysis used nitric acid only). IC analysis was not performed with the acid digestion preparation solution.

Major metals that were well quantified with fusion ICP analysis for tank 241-C-112 were aluminum, calcium, iron, sodium, and uranium; phosphorous is a non-metallic analyte detected by the ICP. In the case of these elements, the fusion result is the preferred value to repeat. Although the assay was performed in a nickel crucible, nickel values from the fusion

preparation will be reported because they are important to interpreting the overall results. This is done with the understanding that they may be biased high. A zirconium crucible was initially recommended for use with these assays to eliminate any potential nickel bias, but the sample matrix reacted with the zirconium during the fusion procedure. However, potassium readings from the ICP fusion are not reported because potassium hydroxide was used to dissolve the sample and the potassium results are not important to characterizing the waste. Some of the primary radionuclides that are measured using this sample preparation are neptunium, plutonium, strontium, cesium, and technetium. A total alpha and total beta count were performed on the fusion dissolution samples as well.

A U.S. Environmental Protection Agency Contract Laboratory Procedure (CLP) type organics speciation analysis was performed on the core composites. No CLP target compounds or tentatively identified compounds were detected in levels above accepted quantitation limits (HASM 1993), and they were not expected to contribute to the sample matrix.

In previous characterization sampling, the core composites were built using quantities of segments based on a proportion of the total weight of sample for the core (Winters et al. 1990a,b). This method assumed that the sample obtained is representative of what is in the tank. However, when partially filled segments are obtained, this procedure assumes that the tank does not contain any waste in this area. Incomplete recovery for a segment is more likely the result of sampling problems rather than voids in the waste.

The approach used in this analysis effort was to composite equal quantities of the homogenized subsegment material and assume that whatever is obtained in a partial subsegment is representative of a whole subsegment. Some inaccuracies may be introduced from this method because of density differences between subsegments. However, the inaccuracies introduced from density differences would probably be small; those deviations are minimal compared to the other errors inherent in core sampling and analysis. If full segments are obtained for the entire core, and the homogenization procedure is satisfactory, there will be little difference between the two approaches.

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5.0 ANALYTICAL RESULTS: TANK 241-C-112

5.1 TANK 241-C-112 CORE SAMPLE RECOVERY

As shown in Figure 5-1, the last 3.8 cm (1.5 in.) of the 48.3-cm (19-in.) sampler does not secure a sample from the bottom segment. In addition, the location of the risers, the dished bottom of the tank, and safety margins in the sampling protocol preclude obtaining samples from the entire waste depth in the tank. Thus, the maximum recovery for the top segment from tank 241-C-112 is 3.8 cm (1.5 in.) above the bit bottom to the waste surface. The next segment will likewise not obtain the lowest 3.8 cm (1.5 in.), but should include the 3.8 cm (1.5 in.) from the *N-1* segment for a full 48.3-cm (19-in.) segment. Segment recoveries were based on the maximum recoverable volume for the segment, regardless of solid/liquid ratio. In the upper segments of tank 241-C-112 (92-001, -003, and -005), the maximum recoverable amount of waste is 33.8 cm (13.3 in.) (131 mL) and 48.3 cm (19 in.) (187 mL) for the lower segments (92-002, -004, and -006). Tables 5-1 and 5-2 present the initial measurements and observations regarding the core samples on extrusion, and an estimate of the core recovery on a volume basis.

Table 5-1. Tank 241-C-112 Core Sample Description Summary.

Core No.	Segment	Core Recovery (Vol. basis)	Total Mass g	Comments
Core 34 Upper	92-001	87.0%	136.9	Liquid contained suspended solids. Solids portion was 1.3 cm (0.5 in.) long.
Core 34 Lower	92-002	74.9%	211.8	Grey/white streak at edge of solids. Solid segment was 36.1 cm (14.2 in.) long.
Core 35 Upper	92-003	0%	N/A	No sample recovered; valve remained open.
Core 35 Lower	92-004	34.8%	109.1	Liquid contained suspended solids. Solid segment was 7.6 cm (3 in.) long.
Core 36 Upper	92-005	64.9%	105.8	Medium brown color; no drainable liquid. Solid segment was 21.8 cm (8.6 in.) long.
Core 36 Lower	92-006	90.9%	263.7	Thin brown sludge at bottom of segment with the material gaining consistency and gradually changing color to grey/white moving up the core. Solid segment was 43.9 cm (17.3 in.) long.

Figure 5-1. Current Condition of Tank 241-C-112.
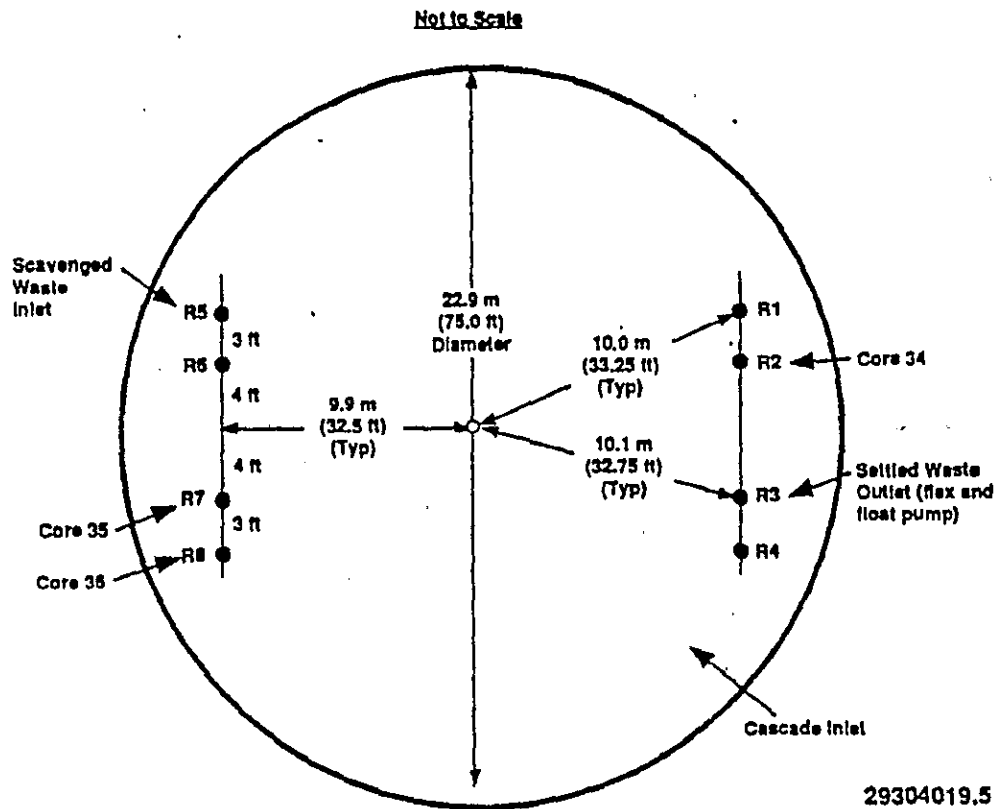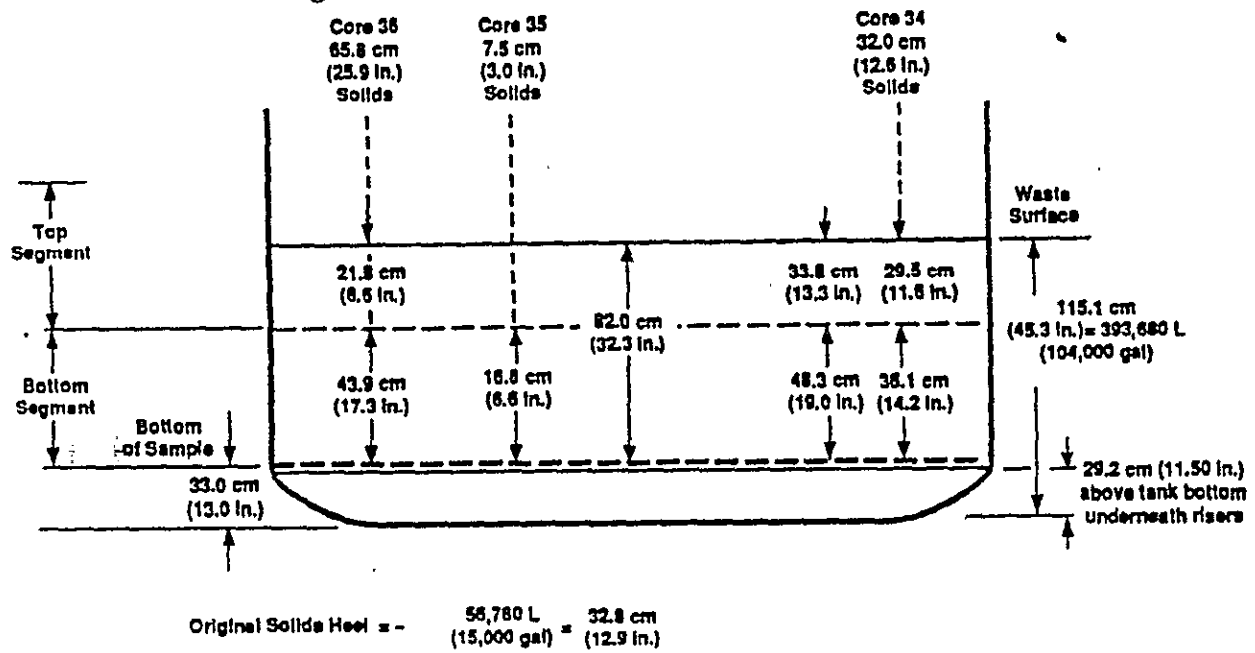
Table 5-2. Tank 241-C-112 Core Sample Physical Characteristics Summary.

Core No.	Segment	Solids Sample Mass (g)	Liquid Sample Mass (g)	Solids Sample Volume (mL)	Liquid Sample Volume (mL)	Solids Bulk Density (g/mL)	Liquids Bulk Density (g/mL)
Core 34	92-001	20.99	115.89	14	100	1.5	1.2
Core 34	92-002	175.75	36.07	110	30	1.6	1.2
Core 35	92-003	N/A	N/A	0	0	N/A	N/A
Core 35	92-004	58.7	50.35	30	35	2.0	1.4
Core 36	92-005	105.8	0.0	85	0	1.2	None
Core 36	92-006	263.7	0.0	170	0	1.6	None

Solids: wet solids
Liquid: drainable (free) liquid.

General characteristics of tank 241-C-112 waste materials are as follows:

- Drainable liquids were all rust to dark brown in color and contained significant amounts of suspended solids. After filtering, the liquids were dark yellow.

- Core samples ranged from grey/white to tan/dark brown in color. No sharp boundaries were observed in the samples. The changes in color occurred gradually over the sample length.

- The samples also ranged in consistency from a thin slurry to a very thick, chunky sludge. They appeared to be saturated with liquid.

- The samples slumped somewhat, but held their shape relatively well (high viscosity, non-Newtonian fluids).

Summary tables of the most significant components are also provided. Analysis of the samples was performed at the PNNL Analytical Chemistry Laboratory facility in the 300 Area of the Hanford Site. The full data package is available from the HASM office (HASM 1993).

5.2 TANK WASTE REMEDIATION SYSTEM PROGRAM ELEMENT CHARACTERIZATION SYNOPSIS

This section provides selected results obtained from core sampling for some of the most pertinent analytes for the various Tank Waste Remediation System (TWRS) program elements, including Hanford Waste Vitrification Plant (HWVP), Retrieval, Pretreatment, and

Waste Tank Safety. Analytes of interest will be reported on a level of resolution commensurate with the available data and program direction. Watch List tanks will have segment or subsegment level analyses reported, while non-Watch List tanks are analyzed only on a core composite basis. Generally, analytes of interest to multiple programs will only be reported in one section. Further detail can be found in the body of the report or in the data packages.

5.2.1 Retrieval Program Characterization Data Summary

A major objective of the Characterization program is to measure the physical properties of the waste to support waste retrieval technology development. The physical characteristics of tank waste are required to develop design criteria for waste retrieval equipment, to provide a basis for simulated waste development, and to provide a basis for validation of equipment testing using design criteria and simulated waste. The analytical methods to determine the physical properties of the waste as it actually exists in the tank require a substantial amount (50 to 100 g) of unhomogenized sample. In some cases, the limited amount of sample recovered constrains the number of analyses that can be performed.

Performing the rheological/physical measurements once for each stratum of waste in a tank is believed to be sufficient to characterize the entire tank contents. Selected rheological and physical properties are presented in Table 5-3; further information regarding these analytes can be found in Section 5.3.

Table 5-3. Retrieval Program Data Summary.

Analyte	Data Range
Specific Gravity (g/mL)	
--solids	1.2 - 2.0
--liquids	1.2 - 1.4
Shear Strength	16,000 dynes/cm^2
Viscosity (cP @ 29 °C)	
--undiluted	160,000 - 220,000
--1:1 dilution	50 (high shear) - 400 (low shear)
--3:1 dilution	~2
Particle Size (μm)	
--Number distribution	97% < 2 μm
--Volume distribution	100% < 85 μm

5.2.2 Pretreatment Program Characterization Summary

The majority of the programmatic decisions pertaining to the design of pretreatment and final disposal systems will be based upon the average characteristics of the tank waste. Therefore, the majority of the laboratory analyses will be conducted on representative core composites. Liquid composites and strata composites will be built under some circumstances, and will be analyzed with fewer assays. As noted previously, segment (and subsegment) analyses will be performed, when directed.

Trace chemical analytes of interest are presented in Table 5-4. This table indicates selected, minor analytes of known interest; further chemical (and radiological) characterization information on primary analytes can be found in Sections 5.4 to 5.6. At this time, there are analytes for which methods are in development, or being phased in as part of a technology transfer effort. In these cases, samples will be archived until the requisite method has been developed and implemented, or samples will be shipped between the onsite laboratories (222-S and 325) and possibly even to offsite laboratories for analysis.

Table 5-4. Pretreatment Program Data Summary.

Analyte	Core Composite Values		
	Core 34	Core 35	Core 36
Minor ICP Analytes (μg/g)			
--B	100	110	140
--Cr	240	210	220
--Mg	430	780	540
--Si	1500	2500	843
--Zr	20	10	30
pH (Drainable liquid)	10.30	10.47	N.M.
pH (Water leach)	10.33	9.77	9.2

N.M. = No measurement.

5.2.3 Waste Tank Safety Program Characterization Data Summary

5.2.3.1 Criticality Safety. The criticality safety program has indicated that plutonium and uranium isotopic analyses on each core composite and the bottom most 15.2 cm (6 in.) of each core is required to alleviate the concern for the potential of tank criticality. For the ferrocyanide tanks, the analyses are performed on the bottom subsegment. As requested, the analyses will indicate whether the fissile species have settled in a concentrated layer at the bottom of a tank. Therefore, upon extruding the last segment in a core, the waste to be tested shall be homogenized before a small aliquot is taken and analyzed for plutonium and uranium isotopic analyses by mass spectrometry (MS) prior to homogenization. Tables 5-5 and 5-6 present the results of these analyses for plutonium and some transuranics; uranium results can be found in Table 5-28.

Table 5-5. Core Composite Transuranics (fusion preparation).

Core No.	^{237}Np (μCi/g)	^{238}Pu (μCi/g)	$^{239/240}$Pu (μCi/g)	^{241}Am$_{GEA}$ (μCi/g)	^{241}Am$_{AEA}$ (μCi/g)	Total α (μCi/g)
Core 34	6.62E-04	0.0137	0.155	0.76	0.613	0.95
Core 35	1.20E-03	0.0137	0.151	1.05	0.763	1.18
Core 36	4.09E-04	0.0033	0.0593	ND	0.0612	0.17

AEA = Measurement by alpha energy analysis
GEA = Measurement by gamma energy analysis
ND = Not detected.

Table 5-6. Plutonium Isotopic Distribution.

Core No.	^{238}Pu mass fraction	^{239}Pu mass fraction	^{240}Pu mass fraction	^{241}Pu mass fraction	^{242}Pu mass fraction
Core 34	0.00037	0.93885	0.05846	0.00153	0.00077
Core 35	N.M.	N.M.	N.M.	N.M.	N.M.
Core 36	0.00023	0.95887	0.03943	0.00103	0.00046

N.M. = Plutonium isotopic distribution not measured, plutonium concentration too low for mass spectroscopy analysis.

5.2.3.2 Ferrocyanide Tanks. During the 1950's, ferrocyanide compounds were used to scavenge ^{137}Cs from the supernate of Hanford Site waste tanks. The potential for an exothermic reaction in the cesium-nickel ferrocyanide/sodium nitrate complex must be evaluated in waste tanks believed to contain 1,000 g-mol or more of ferrocyanide precipitates. The characterization objectives in support of resolution of this USQ and in support of retrieval, pretreatment, and final disposal systems design are as follows:

- Determine the waste energetics behavior in the tanks.

- Determine the spatial distribution of fuel, ^{137}Cs and ^{90}Sr.

- Determine the concentration of total CN$^-$ and the speciation of ferrocyanide present in the waste.

- Satisfy the general characterization requirements for physical, chemical, and radiological analytes.

To achieve the above objectives and to enhance the resolution of the vertical distribution study for key analytes, the analysis horizon for characterization of layering is one-quarter of a segment. The data from tank 241-C-112 indicate that the tank meets the present criterion for placement on the Ferrocyanide Watch List (i.e., greater that 1,000 g-mol ferrocyanide, estimated from total cyanide measurements). Figures 5-2, 5-3, and 5-4 depict the core samples and provide values for analytes of importance to the Ferrocyanide Safety Program as a function of position.

5.2.4 Hanford Waste Vitrification Plant Program Characterization Data Summary

The HWVP program has characterization needs in addition to those described for core sampling. Transforming waste into glass is primarily for the disposal of high-level/transuranic solids in a geologic repository. The vitrification process will be performed after the solids have been pretreated. Therefore, the core sample information will provide preliminary bounding design conditions for the glass plant. Further characterization for technology development and regulatory compliance will be necessary on the pretreated waste that will be fed to the vitrification plant. The analytical requirements for the HWVP program are identified in the *Hanford Waste Vitrification Plant Feed Characterization Requirements* Revision 4, (Wagner 1992).

Neutralized current acid waste (NCAW) is expected to be pretreated by a sludge washing process prior to becoming the first feed to HWVP. The high-heat waste in tanks C-105 and C-106 will also be an early feed to HWVP and is expected to be pretreated in the same manner (one of the two cores from tank C-106 will be analyzed as an early feed tank). Other tanks, such as 241-C-112, may be considered for early retrieval and processing, based on technical and programmatic criteria. Some of the characterization objectives in support of design of retrieval, pretreatment, and final disposal systems are as follows:

- Provide extensive characterization of the chemical and radiological contents of the waste (solids and supernate) as it currently exists in the tanks to support processibility assessments and to verify whether the composition variability study envelope coverage for key analytes is adequate.

- Provide sufficient data to make an estimate of the waste fraction that will remain after sludge wash pretreatment and become feed for HWVP.

- Determine the physical and rheological properties of the waste before and after simulated sludge washing to support the design of a waste retrieval system.

- Satisfy the general characterization requirements for physical, chemical, and radiological analytes.

Figure 5-2. Segments 1 and 2, Core 34 Measurements and Observations.

137Cs μCi/g	90Sr μCi/g	239/240Pu μCi/g	Segment 1	H₂O wt%	Total CN⁻ wt% (dry basis)	NO₂⁻ wt%	NO₃⁻ wt%	ΔH cal/g
			1A					
			1B					
			1C					
240	1300	N.M.	1D	45	0.52	6.0	7.9	-5.1
			Segment 2					
			2A					
610	4900		2B	53	0.43	5.4	7.0	-6.7
800	1100		2C	58	0.83	4.9	6.5	-6.1
510	2500		2D	52	0.75	4.6	6.0	-6.7
0.044	0.35	N.M.	Drainable Liquid	N.M.	0.46	5.5	7.2	N.M.
750	3500	0.155	Core Composite	38	0.97	6.2	8.0	-4.4

29304019.10

Drainable Liquid   N.M. = No measurement

Solids

Figure 5-3.   Segments 1 and 2, Core 35 Measurements and Observations.

Analytes				Analytes				
^{137}Cs $\frac{\mu Ci}{g}$	^{90}Sr $\frac{\mu Ci}{g}$	$^{239/240}$Pu $\frac{\mu Ci}{g}$	Segment 1	H_2O wt%	Total CN⁻ wt% (dry basis)	NO_2^- wt%	NO_3^- wt%	ΔH cal/g
No Sample			1A / 1B / 1C / 1D					
			Segment 2					
			2A / 2B / 2C / 2D					
.007	.23	N.M.	Drainable Liquid	N.M.	0.29	4.6	5.9	N.M.
700	3200	0.151	Core Composite	34	N.M.	3.5	4.4	-6.0
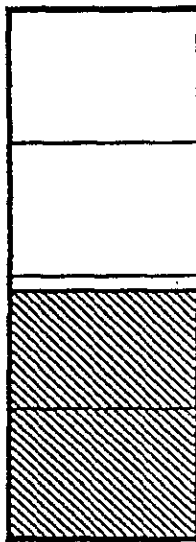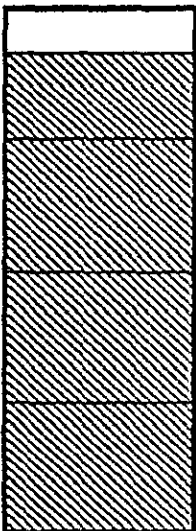
Drainable Liquid   N.M. = No measurement   29304019.8

Solids

Figure 5-4. Segments 1 and 2, Core 36 Measurements and Observations.

Analytes   **Analytes**

^{137}Ce $\frac{\mu Ci}{g}$	^{30}Sr $\frac{\mu Ci}{g}$	$^{239/240}$Pu $\frac{\mu Ci}{g}$		Segment 1	H_2O wt%	Total CN⁻ wt% (dry basis)	NO_2^- wt%	NO_3^- wt%	ΔH cal/g
			1A						
			1B						
560	1900		1C		49	N.M.	4.8	6.2	-5.8
1200	15		1D		58	0.72	5.1	6.7	-8.0
				Segment 2					
880	20		2A		57	0.92	4.9	6.6	-5.2
530	70		2B		41	0.75	3.0	4.3	-3.7
100	140		2C		64	0.40	3.2	4.6	No Exotherm
40	200		2D		56	0.56	3.5	5.1	No Exotherm
800	510	0.06		Drainable Liquid Core Composite	45	0.71	5.3	7.2	-8.6

Drainable Liquid - None   N.M. = No measurement   29304019.9

▨ Solids

Planned early retrieval of some tanks for HWVP necessitates an increased evaluation of rheological and physical properties of the tank waste. To this end, selected rheological and physical properties are measured on the first and last segments of both cores taken from these tanks as a minimum. Rheological measurements will also be performed on other segment material if a unique stratum is identified in the remaining segments.

The analytical program for HWVP not only entails determining whether a waste type is suitable for disposal as glass, but also includes determining the physical and chemical characteristics of the glass for process control purposes and to ensure regulatory compliance. Sampling and analysis plans will be developed on an individual basis for each tank or process batch. The characterization needs for these efforts include analyses for metals, water-soluble anions, radionuclides, semi-volatile organics, and rheological and physical testing for both the HWVP feed and vitrified product.

Presently, tank 241-C-112 is not scheduled as an early feed to the HWVP. However, in recognition that the tank may be considered in the future as a potential early retrieval/processing candidate, two selected groups of analytes are presented in this summary. Table 5-7 provides a set of analytes of interest to the vitrification process control. Table 5-8 presents analytes of interest to the regulatory permitting of such a facility.

Table 5-7. HWVP Process Stream Analytes of Concern.

Analyte	Core 34 Composite (μg/g)	Core 35 Composite (μg/g)	Core 36 Composite (μg/g)
PO_4^{3-}	62,000	34,500	52,500
F^-	1,000	300	450
Cl^-	1,300	750	1,050
TOC	3,100	1,200	3,100

Table 5-8. HWVP Regulatory Operation Analytes of Concern.

Analyte	Core 34	Core 35	Core 36
Hg (D. Liquid)	3.53	1.30	4.40
Pb (D. Liquid) (Composite)	< DL 3,300	< DL 4,600	N.M. 1,050
Cr(VI) (D. Liquid)	130	93	75*
^{14}C (D. Liquid) (Composite)	0.004 N.M.	0.002 N.M.	0.002** N.M.
^{99}Tc (D. Liquid) (Composite)	0.112 0.139	0.084 0.097	N.M. 0.107
^3H (D. Liquid)	0.232	0.002	0.016*

*Water leach.
**Based on a single measurement.

5.3 TANK 241-C-112 CORE SAMPLE RHEOLOGICAL/PHYSICAL MEASUREMENTS

Physical properties such as shear strength, viscosity, particle size, and settling properties were measured. These measurements are necessary for the design and fabrication of retrieval, pretreatment, and final waste disposal systems.

5.3.1 Shear Strength

The shear strength of tank 241-C-112 core 36 was measured on a combined, unhomogenized sample obtained from both segments of the core. The shear strength measurements were made at ambient temperature using a shear vane connected to a viscometer and rotated at 0.3 rpm. Shear strength (τ_s) is a semiquantitative measurement of the force required to move the sample. Because shear strength is dependent on sample handling, the measurement was taken without any sample homogenization. The rheology sample was generated by taking small aliquots from the bottom segment of core 36 at various positions. The aliquots were transferred to a sample jar and allowed to settle for 10 weeks to let the sample recover from the disturbance of sampling and extrusion. The extended delay between sample and analysis was specified because it is believed that the longer the sample sits undisturbed, the more likely it is to return to its (nearly) original condition; therefore, the shear measurement is likely to be more representative. The shear stress (S_r) of the sample was recorded as a function of time and the shear strength was calculated using Equation 1.

$$\tau_s = \frac{[\%\tau/100]*S_r*4.9E+05}{\dfrac{\pi*H_v*D_v^2}{2} + \dfrac{\pi*D_v^3}{6}} \tag{1}$$

where

$\%\tau/100$ = The ratio of the total torque which is recorded as full scale on the plot of the shear stress

S_r = Shear stress

$4.9E+05$ = Maximum torque of the viscometer head (dynes)

H_v = Shear vane height (0.635 cm)

D_v = Shear vane diameter (0.635 cm)

Shear strength for the sample was found to be 16,000 dynes/cm^2. The RPD between initial and duplicate measurements of the sample was less than 1%.

5.3.2 Viscosity as a Function of Shear Rate

Viscosity measurements (as a function of shear rate) were performed on the composite sample and the 1:1 sample:water dilution of the sample at ambient hot cell temperatures 29 to 32 °C (84 to 90 °F) and at 95 °C (203 °F). At 95 °C the undiluted core composite sample dried too quickly to obtain an accurate viscosity measurement; therefore, no data are presented for the undiluted sample at that temperature. The rheological properties for the undiluted sample were not characteristic of any accepted behavior models. In addition, the undiluted sample behavior at high shear rates is considered suspect. Viscosity of the undiluted sample at low shear ranged from 160,000 to 220,000 cP. The 1:1 dilution of the composite sample exhibited yield-pseudoplastic behavior. Plots of the measurements can be found in the summary data packages or in the full validated data packages (HASM 1993). Viscosities as a function of shear rate for the 1:1 dilutions ranged from 400 cP (at low shear rates) to 50 cP (at high shear rates).

Further measurements of the shear stress as a function of shear rate were made on the 1:3 sample:water dilution samples at ambient temperature and at 95 °C (203 °F). The ambient samples were run in duplicate; however, due to drying of the sample, only a single measurement could be performed at 95 °C (203 °F). All of the 1:3 diluted samples had viscosities near the detection limit of the apparatus (2 cP). The diluted samples also exhibit yield-pseudoplastic behavior. But, at viscosities near the detection limit, accurate modeling of the flow properties with this data becomes difficult. The viscosity of the sample was observed to decrease significantly as the temperature increased.

The data from the rheograms for the 1:1 dilution were fit to a nonlinear yield power-law model (Equation 2). Sample and duplicate measurements were run at ambient and 95 °C.

$$S_\tau = \alpha + \beta\gamma^n$$

(2)

where

S_τ = Shear stress
α = Yield stress (not a fit parameter)
β = Consistency factor
γ = Shear rate (0 to 468 s^{-1})
n = Flow behavior index.

Table 5-9 presents the power law model parameters for the 1:1 sample dilutions at 29 and 95 °C.

Table 5-9. Power-Law Model Parameters for Tank 241-C-112 Material.

Sample	Temperature (°C)	Trial	α, Yield Stress (Pa)	β, Consistency Factor (Pa•s)	n, Flow Behavior Index
1:1 Dilution	29	S	6.8	0.279	0.576
1:1 Dilution	29	D	5.8	0.302	0.534
1:1 Dilution	95	S	3.6	0.079	0.68
1:1 Dilution	95	D	4	0.097	0.648

S = Sample
D = Duplicate.

A rheogram for a material with a yield stress has two sections. The first section is a straight line beginning at the origin and climbing up the ordinate. This portion of the rheogram records the material as it acts like a solid or gel. When sufficient force is applied to the material to make the gel yield, the rheogram breaks sharply to the right; recording the material's behavior as a fluid. The point on the rheogram at which the sample's behavior transfers from a solid or gel to a fluid is the yield point or yield stress. The consistency factor in this model is analogous to viscosity. The flow behavior index indicates the degree of deviation from Newtonian behavior. For values less than 1, the behavior is considered pseudoplastic (Bird et al. 1960).

5.3.3 Slurry Flow Properties

Turbulent flow is necessary to keep particles in suspension and prevent the accumulation of the solids in retrieval and/or pretreatment process equipment. Characteristics necessary for turbulent flow were calculated for the 1:1 dilution slurry using the parameters determined from measurement and a curve-fitted rheological model (refer to Table 5-10).

Table 5-10. Turbulent Flow Model Calculations.

Sample	Temp. (°C)	Trial	Pipe Dia. (in.)	Velocity (m/s)	Critical Flow Rate (L/min)	Reynolds Number
1:1 Dilution	29	S	2	1.9	246	4,425
1:1 Dilution	29	D	2	1.7	220	4,470
1:1 Dilution	29	S	3	1.7	496	4,920
1:1 Dilution	29	D	3	1.6	447	4,908
1:1 Dilution	95	S	2	1.3	163	5,190
1:1 Dilution	95	D	2	1.3	170	5,214
1:1 Dilution	95	S	3	1.2	329	6,002
1:1 Dilution	95	D	3	1.2	344	5,997

S = Sample
D = Duplicate.

5.3.4 Particle Size Measurement

Particle size analysis is performed by placing a small amount of sample in a dispersant (which is the liquid used to disperse and suspend the particles from the solid sample). The prepared sample was placed in a particle size analyzer. The apparatus measures particle size by passing a thin beam of laser light through the dispersant. The diameter of a particle of matter in the dispersant can be determined by the amount of light that it blocks as the particle passes through the beam. The dimension measured by this method is the value across the short diameter of the particle. This means that if a particle is oblong, the machine estimates the shortest length across the particle (i.e., the width of the oblong shape, not the length). The term "diameter" throughout this text will be used to describe any linear profile of any shape.

An important consideration involving the analysis of particle size is the dispersant used. The primary concern involved with the dispersant is dissolving the particle. Any particles existing in the tank that are soluble in the dispersant will dissolve or decrease in size during the analysis. Depending on the dispersant, the particle size analysis may not represent the true particle size distribution in the tank. In the case of tank 241-C-112, a mixture of water and glycerin was used as the dispersing medium. If a "true" particle size distribution is required, the mother liquor (drainable liquid) of the tank should be used, if possible, because the tank particulates are already in equilibrium with the tank mother liquor. The high insolubility of the waste matrix suggests that the particle size data acquired should be acceptable. However, if the ferrocyanide waste has been hydrolyzed by high-pH waste, this assumption may not be completely accurate.

The mean particle size in the number distribution ranges from 0.83 microns to 0.95 micron in diameter for the tank core samples. Table 5-11 presents the summary results of the measurements. Plots of the distributions are presented as Figures 5-5 and 5-6. The first graph is the probability number density. The number density graph is plotted over the acquisition range of the device (from 0.5 to 150 microns). The numbers of particles in each size range (shown as a percentage of the whole) are graphed against their respective size ranges to form a distribution curve. It can be seen from Figure 5-5 that the most common occurrences (modes) for particle size range between 0.5 and 1.0 microns. The majority (over 90%) of the measured particles fit within the narrow band of 0.4 to 1.5 microns, and over 97% of the particles have a diameter of less than 2 microns.

Table 5-11. Particle Size Distribution by Number: 97% < 2μm (both cores).

Sample	Mean (μm)	Median (μm)
Core 34, subsegment 2D, Initial	0.83	0.76
Core 34, subsegment 2D, Duplicate	0.94	0.83
Core 36, 92-005 (random sample)	0.95	0.84

The particle size in the volume distribution ranges from 0.4 microns to 80 microns in diameter between the two cores. Table 5-12 presents the summary results of the measurements. Under the assumption that the density of the solid material within the tank is constant, the volume distribution is also the best estimation of the mass particle size distribution of the tank. The analyzer calculates particle volume as the cube of the diameter. These distributions are presented as Figures 5-7 and 5-8.

Table 5-12. Particle Size Distribution by Volume: 100% < 30μm (34-2D); 100% < 85μm (36).

Sample	Mean (μm)	Median (μm)
Core 34, subsegment 2D, Initial	8.68	6.05
Core 34, subsegment 2D, Duplicate	9.60	6.32
Core 36, 92-005 (random sample)	33.77	33.26

As with the number distribution, the volume distribution is represented by a probability volume density graph. The average particle size in the volume distribution is considerably larger than that of the number distribution. In core 34 most of the particles are within the 2.0 to 20 micron range. In core 36 most of the particles are much larger, with particle volumes widely dispersed in the 20 to 80 micron range.

Figure 5-5.   Core 34, Particle Size Number Density.
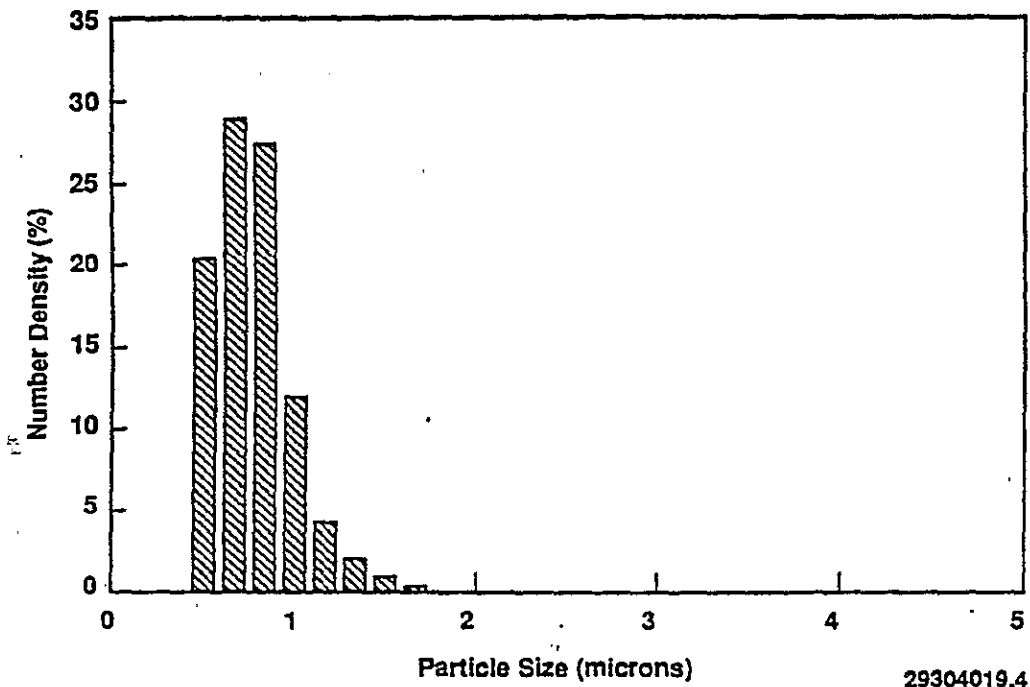
29304019.4

Figure 5-6.   Core 36, Particle Size Number Density.
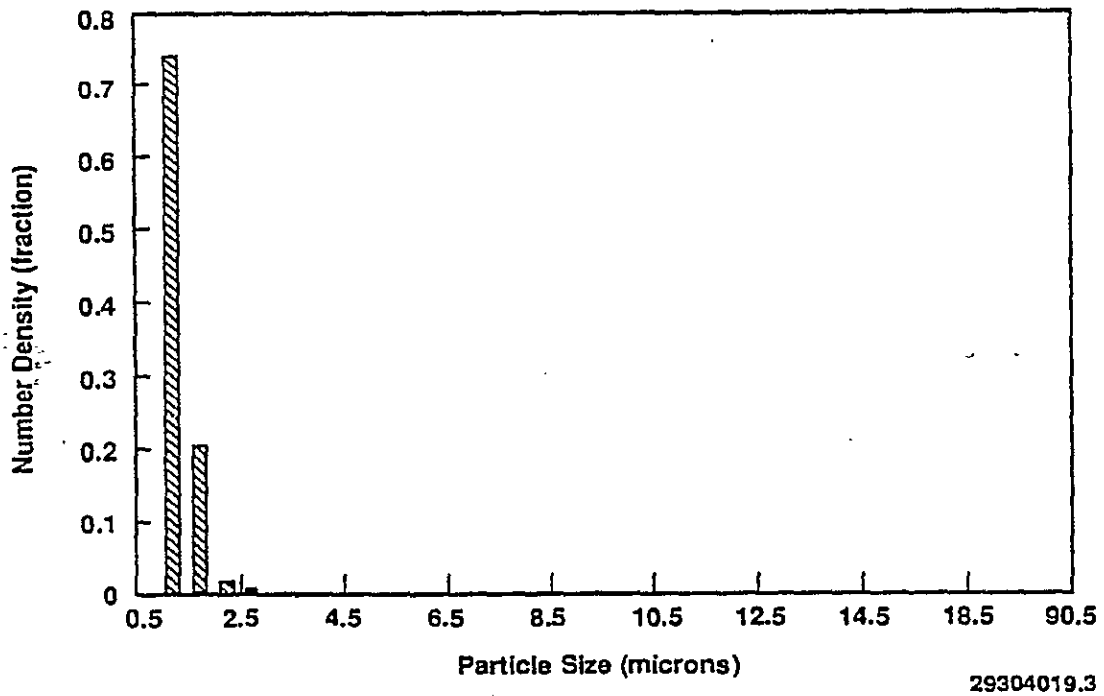
29304019.3

Figure 5-7. Single-Shell Tank Core 34, Particle Size Volume Density.
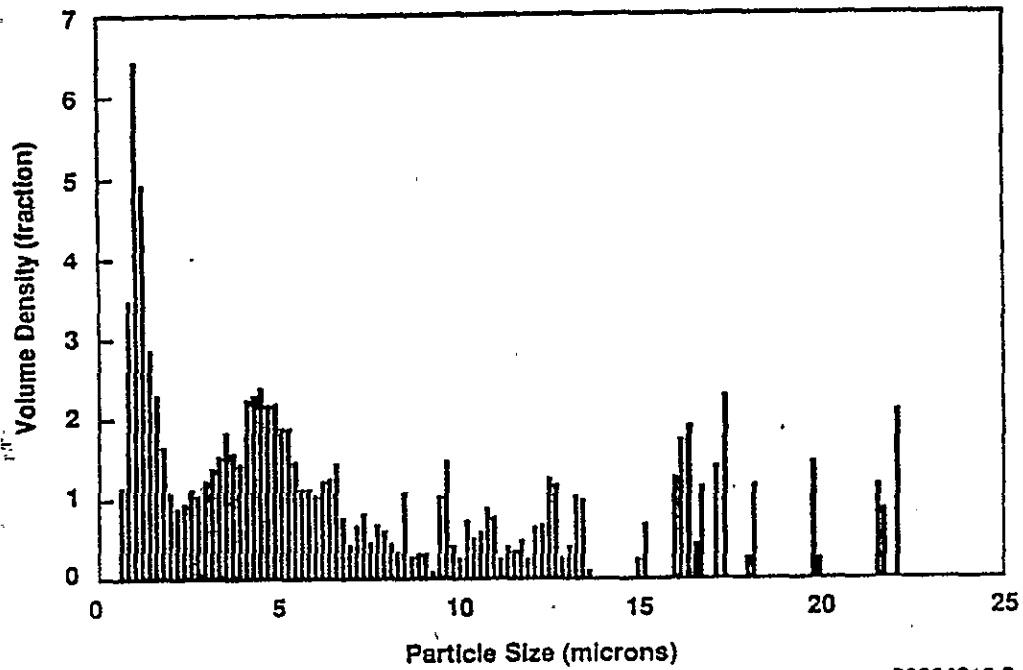
29304019.2

Figure 5-8. Single-Shell Tank Core 36, Particle Size Volume Density.
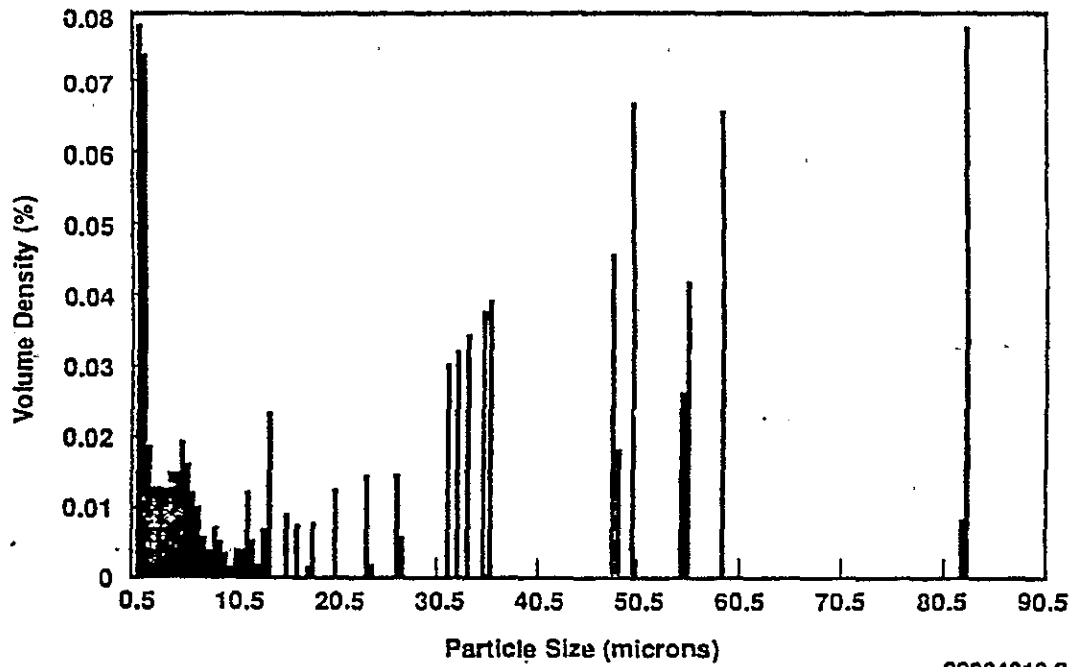
29304019.2

The disparity between the two core sample measurements possibly indicates a difference in waste type. In core 34-2D, over 50% of the particles in the sample have a diameter of less than 6 microns, with close agreement from the duplicate measurement. In the core 36 sample over 50% of the particles have a diameter of less than 33 microns. In the retrieval and subsequent treatment of the tank wastes, it may be desirable to design pumping or filtration systems for the tank particulate. Therefore, the volume distribution of the particles should not be neglected (i.e., particles with diameters of over 33 microns should be considered in these designs).

5.3.5 Settling Behavior of Diluted Samples

This section analyzes the settling behavior for the 1:1 and 3:1 water/sample dilutions and the viscosity as a function of shear rate on the 3:1 dilution. All results for the as-received material and the viscosity for the 1:1 dilution have been previously reported (HASM 1993). The physical properties reported here include settling rates and volume percent settled solids, and weight percent and volume percent centrifuged solids for the 3:1 dilution and settling rates and volume percent settled solids for the 1:1 dilution. The experimental procedures used to perform these measurements were reported previously (HASM 1993).

The physical properties of the 1:1 and 3:1 dilutions are summarized in Table 5-13.

Table 5-13. Physical Properties Summary.

Property	Segment	
	1:1 Dilution	3:1 Dilution
Settled Solids (vol%)	74.4	72.4
Centrifuged Solids		
Vol%	NM	21.1
Wt%	NM	27.0
Density (g/ml)		
Sample	NM	1.11
Centrifuged Supernate	NM	1.01
Centrifuged Solid	NM	1.39

NM = No measurement.

No settling was observed in the as-received segment samples over a period of 3 days. Two dilutions (1:1 and 3:1 water to sample) were prepared, and the volume percent settled solids for each of the dilutions are plotted as a function of settling time (as shown below in Figures 5-9 and 5-10).

Figure 5-9. Settling Rate Data for Tank 241-C-112 Core 36, 1:1 Dilution.
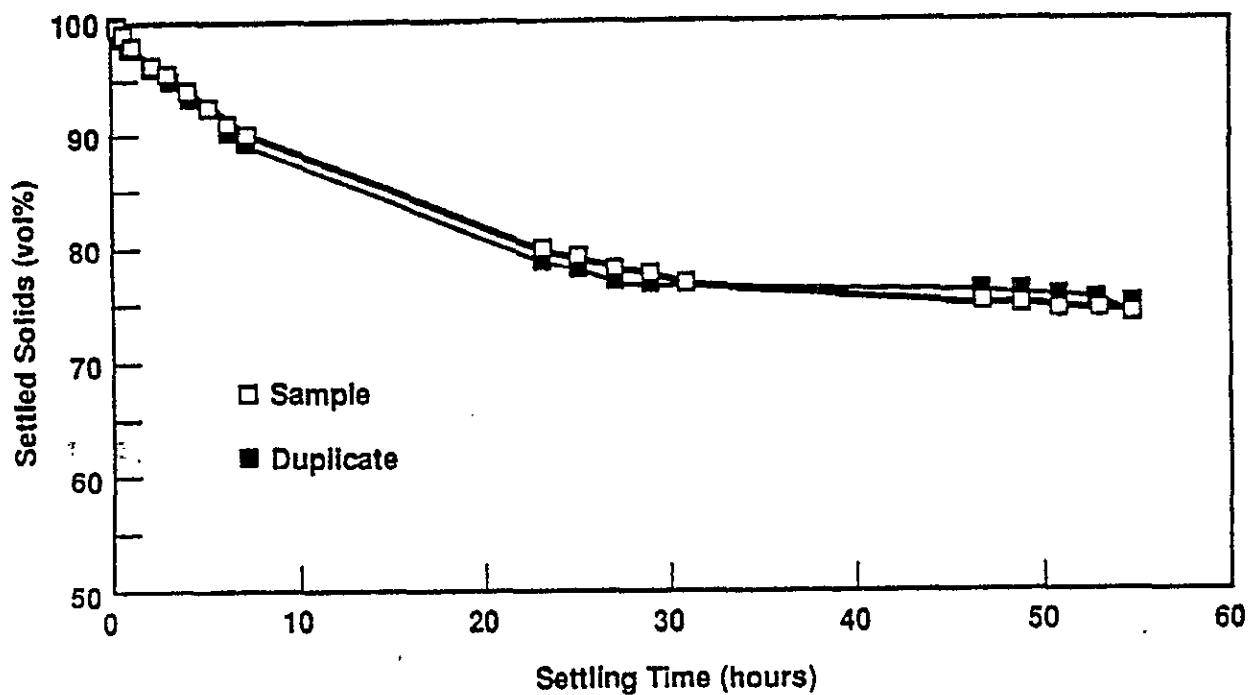
Figure 5-10. Settling Rate Data for Tank 241-C-112 Core 36, 3:1 Dilution.
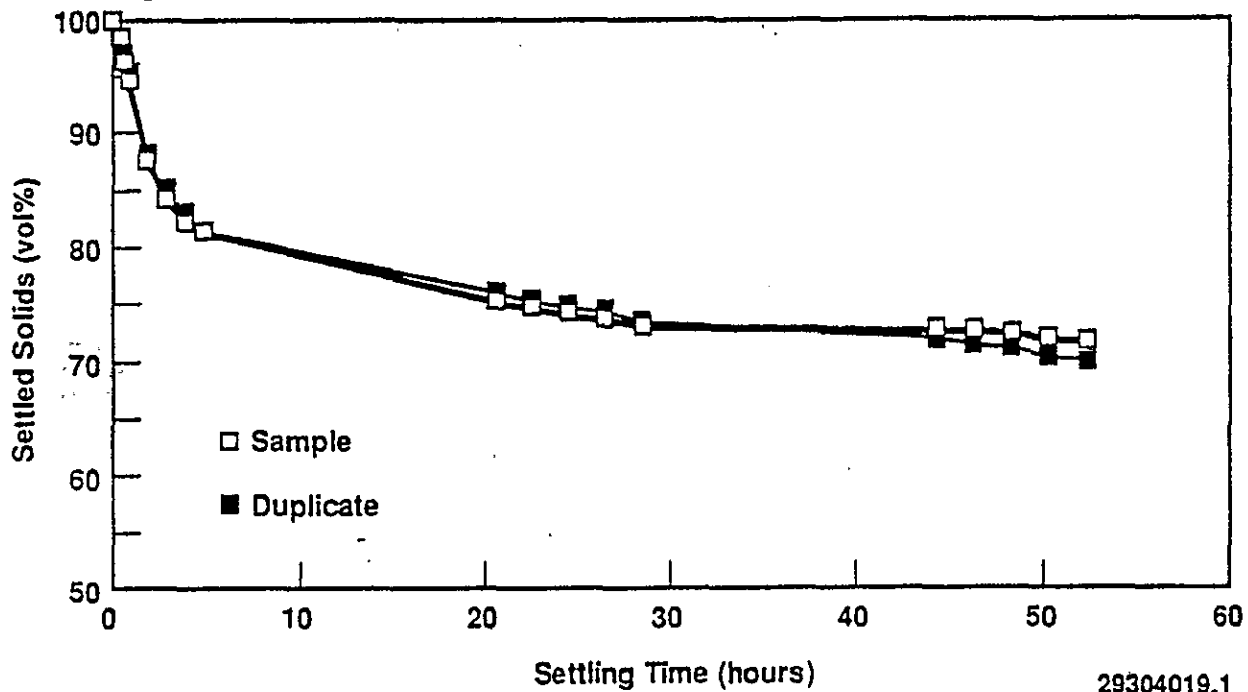
29304019.1

The 1:1 dilution reaches a final volume percent settled solids of 74.4% (avg.). Settling continues throughout the 3-day period, but the majority of the settling is seen in the first 30 hours. The 3:1 dilution reaches a final volume percent settled solids of 72.4%. The majority of the solids settling is complete within 24 hours. Figures 5-9 and 5-10 illustrate the setting behavior over time.

5.4 ANALYTICAL RESULTS--INDUCTIVELY COUPLED PLASMA-ATOMIC EMISSION SPECTROSCOPY

5.4.1 Inductively Coupled Plasma Assays--General Comments

No online multiple inter-element corrections were performed for matrix interferences. The ICP has built-in correction capability to adjust for moderate matrix interferences; however, this function does not perform well on samples containing weight percent quantities of iron, aluminum, or uranium. Single-pass off-line corrections were performed to correct for high aluminum, iron, and uranium content in the samples. As requested, process blank values have not been subtracted from the reported values (except for nickel-fusion and Quality Control [QC] results). In the water digestion and drainable liquid assays, the single most prevalent element is sodium, by at least an order of magnitude. In the fusion assays, some elements (lead, for example) appear to be at high concentrations because of the large dilution factors required for fusion samples. Those analytes may actually only be present in concentrations marginally above the detection limit. Selenium routinely demonstrated a low bias; however, it is not a major analyte in the waste matrix. An estimate of the detection limit for any analyte can be obtained by multiplying the analyte's detection limit factor, based on dilution by the appropriate sample "$\mu g/g$ factor" found in the data packages. Analytes reported in the data tables are those consistently contributing significant amounts to the composition of the waste matrix; average values for the analytes are reported to 3 significant figures based on two replicates. The full range of ICP analytes can be found in the full data packages (HASM 1993). All reported concentration values are based on grams of wet sample, unless otherwise specified.

Tables 5-14, 5-15, and 5-16 provide ICP analyte concentration information on the core composites as a function of the sample preparation.

Table 5-14. Water Digestion Chemical Composite Data--ICP Average Values.

Analyte	Core 34 (μg/g sample)	Core 35 (μg/g sample)	Core 36 (μg/g sample)
Al	315	330	814
Ca	240	339	385
Fe	1,630	712	1,390
Na	105,000	60,500	108,000
Ni	1,000	407	792
P	6,380	5,630	16,500
U	715	460	4,600

Table 5-15. Acid Digestion Chemical Composite Data--ICP Average Values.

Analyte	Core 34 (μg/g)	Core 35 (μg/g)	Core 36 (μg/g)
Al	23,300	44,600	5,530
Ca	21,700	11,500	17,000
Fe	20,100	26,000	19,500
Na	95,100	70,800	103,000
Ni	17,700	11,100	10,600
P	19,100	18,800	30,800
U	12,100	79,200	83,900

Table 5-16. Fusion Digestion Chemical Composite Data--ICP Average Values.

Analyte	Core 34 (μg/g)	Core 35 (μg/g)	Core 36 (μg/g)
Al	29,800	45,000	6,410
Ca	28,900	15,000	20,400
Fe	22,800	32,400	26,000
Na	115,000	81,600	121,000
Ni*	30,000	19,900	12,800
P	22,000	21,200	36,800
U	14,400	89,700	105,000

Tables 5-17, 5-18, and 5-19 provide ICP analyte concentrations as a function of depth, i.e., for the subsegments.

Table 5-17. Tank 241-C-112 Core 34 ICP Analyte Trending (fusion prep on subsegments).

Subsegment	Al (μg/g)	Ca (μg/g)	Fe (μg/g)	Na (μg/g)	Ni* (μg/g)	P (μg/g)
1D	17,600	28,200	8,500	91,000	22,700	12,800
2B	32,400	21,700	14,700	89,000	22,200	11,100
2C	17,300	29,900	10,200	90,000	28,200	19,700
2D	26,900	23,800	19,700	78,600	23,500	19,500
W. Leach (Core. Comp)	315	240	1,630	105,000	1,000	6,380
D. Liquid	<DL	<DL	1,200	85,700	790	3,800

Table 5-18. Tank 241-C-112 Core 35 ICP Analyte Trending (fusion prep on subsegments).

Subsegment	Al (μg/g)	Ca (μg/g)	Fe (μg/g)	Na (μg/g)	Ni* (μg/g)	P (μg/g)
2D	45,000	15,000	32,400	81,600	19,900	22,200
W. Leach (Core Comp.)	330	339	712	60,500	407	5,630
D. Liquid	<DL	410	750	69,900	440	3,960

Table 5-19. Tank 241-C-112 Core 36 ICP Analyte Trending (fusion prep on subsegments).

Subsegment	Al (μg/g)	Ca (μg/g)	Fe (μg/g)	Na (μg/g)	Ni* (μg/g)	P (μg/g)
1C	14,700	28,900	36,000	81,400	21,800	19,300
1D	4,000	28,600	7,500	81,700	23,000	19,300
2A	2,900	21,400	11,300	91,800	11,000	21,200
2B	3,100	8,900	8,900	90,600	4,800	25,200
2C	2,100	3,000	15,100	106,000	1,000	33,600
2D	3,000	2,100	31,200	105,100	900	29,900
W. Leach (Core Comp.)	814	385	1,390	108,000	792	16,500
D. Liquid	None	None	None	None	None	None

*Ni concentrations are potentially biased high. Values are derived from ICP fusion performed in a Ni crucible. However, in each case the blank value was an order of magnitude less (or greater) than the measurement. The fusion values are comparable with acid digestion values, where both are available.

5.4.1.1 Core 34. The ICP assays for the water leach of the core composite and drainable liquid samples were found to be quite similar. Because the drainable liquid is in equilibrium with the solids and the water leach is done at a 100:1 dilution, the similarity is unexpected. This similarity suggests that some soluble compounds have crystallized from the drainable liquid in the solids in addition to the insoluble materials. The water leach of drainable liquid samples were compared and used to determine RPDs for the analytes. RPDs were very close for all major analytes, within 2.5%.

The acid digestion assay for the drainable liquid gives results similar to those obtained from the core composite water leach. This treatment was done on a liquid sample to solubilize any particles suspended in the sample not removed by filtration. The major elements found were sodium, phosphorus, iron, uranium, nickel, potassium, and calcium. A full QC suite was performed on the drainable liquid; the percent recovery for spiked samples and controls were excellent in each case (all recoveries were greater than 92.6%). On the solid samples, spike concentrations for iron, sodium, and nickel were insufficient for quantitation (i.e., the concentration of these analytes in the sample was too high for the spike to be detected). Results from the subsegments and core composites correlate reasonably well with the fusion results. RPDs are good (most are within 20%), except for calcium, iron, manganese, and zinc. The calcium anomaly is believed to be the result of a high system blank. The iron, manganese, and zinc RPDs are outside the typical 20% precision criteria. This behavior is demonstrated throughout the acid digestion analyses and was not unexpected. Both the In-Farm simulant matrix and homogenization samples were highly resistant to acid dissolution, and the samples were expected to show similar behavior. The ICP homogenization test results showed no significant difference between the top and bottom samples of subsegments 2C and 2D, respectively. However, high RPDs for iron, manganese, zinc, and aluminum for each segment sample indicate the possibility of heterogeneity due to a crystalline or particle inclusion.

The RPDs for the fusion ICP assay were found to be generally good for the subsegment major analytes, routinely within 20%. However, there was not enough sample to perform a duplicate measurement on the core composite. The fusion results correlate reasonably well with the acid digestion assay, within 30% for major analytes. This difference between acid and fusion digestion ICP results further indicates the high insolubility of the waste. A high iron process blank was found in this sample, and its source is undetermined.

5.4.1.2 Core 35. The ICP assays for the direct drainable liquid and acid digestion of the drainable liquid samples were found to be almost identical and quite similar to the water leach of the core composite. RPDs were very good for all major analytes (within 20%).

The acid digestion assay for the drainable liquid gives results similar to those obtained from the core composite water leach. A full QC suite was performed on the drainable liquid percent recovery for spiked samples, and controls were excellent in each case (above 90% for major analytes). The predominant elements found were sodium, phosphorus, iron, uranium, nickel, potassium, and calcium. On the solid samples, spike concentrations for

iron, sodium, and nickel were again insufficient for quantitation. RPDs were generally higher than core 34, but within 20%, except for calcium, iron, and low concentration analytes near the detection limit in the 5x dilution run (the sample was diluted to five times its original volume). Concentrations of 8 to 9% uranium, 4 to 5% aluminum, and 3% iron make accurate inter-element correction for matrix interferences difficult; thus, the high concentration levels of many "unlikely" analytes (such as lanthanum, neodymium, and thallium) are probably the result of inadequate interelement correction. RPDs for most analytes significantly above the quantitation limit were within the duplicate precision criteria of 20%, except for aluminum, uranium, and phosphorus. Duplicate analyses for this assay were unusual--nearly all analytes found in the duplicate are 15 to 20% lower than the original. The exception in this case is aluminum, which was 30% higher. This behavior suggests that the poor RPDs are the result of an aluminum compound inclusion. The homogenization test showed no significant difference for subsegment 2D top and bottom. The RPDs were acceptable; consequently, the homogenization of the samples was considered adequate. The homogenization test values compared well with the core composite results. Some small silicon inhomogeneity was observed.

The fusion assay appears to improve quantitation for aluminum and silicon. Results for analytes that require significant uranium correction and have low concentration should be considered qualitative (chromium, calcium, titanium, barium, lanthanum). RPDs for major analytes are good (within 15%).

5.4.1.3 Core 36. There was no drainable liquid associated with these samples. Water leach RPDs of the core composites were found to be poor (generally above 25%). This behavior is attributed to differences in particle size and solubility in the aliquots used to make the composite.

The acid digestion values for the core composite correlate well with the fusion results. RPDs on 1x and 5x dilutions are very good except for calcium and analytes near the detection limit in the 5x dilution run (potassium, lanthanum, neodymium, thallium, and vanadium). RPDs for all analytes that are above the quantitation limit were within the 20% process control criteria. Percent recovery for spiked samples and controls was good except for those analytes for which the spike was less than 25% of the measured sample concentration. Digestion spike concentrations for aluminum, calcium, iron, manganese, sodium, nickel, and uranium were insufficient for percent recovery quantitation. As previously noted, selenium demonstrates a low bias (50%) for both spiked samples and controls. The homogenization tests on subsegments 1C, top and bottom, and 2D, top and bottom, showed no significant variation within the subsegment (<10% RPD). However, a large variation is observed between the two subsegments. This behavior was expected from the process history of the tank (e.g., the waste heel that remained prior to adding scavenged waste). Based on sample RPDs, the homogenization of the samples appears adequate.

The RPDs for the fusion ICP assay were found to be generally good for the subsegments and core composite major analytes, each within 20%. RPD for iron is good, contrasted with poor RPD for iron in the acid assay. This also suggests a minor sample inhomogeneity.

5.5 ANALYTICAL RESULTS--ION CHROMATOGRAPHY

5.5.1 Ion Chromatography Assays--General Comments

IC analyses were performed on water-leached samples of the subsegments, core composites, and diluted samples of the drainable liquid (where applicable). Matrix components in some of the samples were found to affect detector performance reversibly during the analysis for free cyanide; thus, a modification to the procedure using pulsed electrode cleaning between sample injections was incorporated to overcome that effect. In addition, the free cyanide assay produced much higher results than those anticipated from the simulant studies. No spike or control standard was used for the free cyanide analysis; therefore, these results may not be representative of the free cyanide in the samples. Quantitation for fluoride (and possibly chloride) was compromised by a co-eluting matrix interference, probably organics of some type. This supposition is supported by the TOC results from the water leach samples. The TOC values, although not high, are large enough to potentially interfere with fluoride and chloride detection. Further information regarding the IC analytes can be found in the full data packages (HASM 1993).

Table 5-20 shows the concentration of IC analytes for each core composite sample. Tables 5-21, 5-22, and 5-23 provide IC analyte concentrations as a function of depth. Tables 5-24, 5-25, and 5-26 provide additional information on other anions such as total cyanide as a function of depth. These anions were not determined by the IC method. Reported pH values for the subsegments and core composites are for 1:100 diluted samples; therefore, only the pH measurement of the drainable liquids (direct pH measurements) are meaningful. The TOC and total inorganic carbon (TIC) assays are not considered capable of measuring the total cyanide in the waste because they depend on acid dissolutions to perform the analyses.

5.5.1.1 Core 34. Fluoride spike recovery in the core composite was 50%; the poor recovery was attributed to matrix interference. All other spike recoveries (where applicable) ranged between 87 and 139%, indicating some minor matrix interferences (generally biased high). Control-standard-recoveries ranged from 87 to 112%, indicating that the analysis was in control at the time of the assays. The drainable liquid results are similar to the core composite and subsegment water leach results. This suggests a liquid in saturated equilibrium with the solid core material. RPDs were within 15% for all IC anions.

Table 5-20. Anion Analyses--Composite Data Results.

Analyte	Core 34 (μg/g)	Core 35 (μg/g)	Core 36 (μg/g)
NO_2^-	62,000	34,500	52,500
NO_3^-	80,000	43,500	71,500
PO_4^{3-}	19,100	17,850	49,500
SO_4^{2-}	15,600	8,600	13,700
Cl^-	1,300	750	1,050
F^-	1,000	300	450
Free CN^-	2,050	780	1,300
Total Carbon[1]	11,700	6,100	7,000
TOC[1]	3,100	1,200	3,100
TIC[1]	8,600	4,900	4,000
Total Cyanide[2]	9,700	NM	7,100

[1]Total Carbon, TOC, and TIC are not IC analyses, but are probably present as anions (TOC + TIC = Total carbon). Thus, it seems appropriate to include them with this table.

[2]Total Cyanide is not an IC anion. Presently it is a developmental assay; however, the total cyanide assay is important in interpreting the data.

IC = Ion chromatography (NO_2^-, NO_3^-, PO_4^{3-}, SO_4^{2-}, Cl^-, F^-, Free CN^-)
NM = No measurement
TIC = Total inorganic carbon
TOC = Total organic carbon
Note: All IC results are obtained from a water leach preparation.

Table 5-21. Tank 241-C-112 Core 34 IC Analyte Trending.

Subsegment	NO_2^- (μg/g)	NO_3^- (μg/g)	PO_4^{3-} (μg/g)	SO_4^{2-} (μg/g)	Cl^- (μg/g)	F^- (μg/g)
1D	60,000	79,000	11,650	14,400	1,100	1,000
2B	53,500	69,500	12,100	13,050	1,000	900
2C	48,500	64,500	11,500	11,750	900	900
2D	45,500	59,500	17,500	11,150	850	950
W. Leach (Core Comp.)	62,000	80,000	19,100	15,600	1,300	1,000
D. Liquid	55,000	72,000	11,250	11,650	1,000	600

Table 5-22. Tank 241-C-112 Core 35 IC Analyte Trending.

Subsegment	NO$_2^-$ (µg/g)	NO$_3^-$ (µg/g)	PO$_4^{3-}$ (µg/g)	SO$_4^{2-}$ (µg/g)	Cl$^-$ (µg/g)	F$^-$ (µg/g)
2D	34,500	43,500	17,800	8,600	850	300
D. Liquid	46,000	58,500	12,900	10,900	900	300

Table 5-23. Tank 241-C-112 Core 36 IC Analyte Trending.

Subsegment	NO$_2^-$ (µg/g)	NO$_3^-$ (µg/g)	PO$_4^{3-}$ (µg/g)	SO$_4^{2-}$ (µg/g)	Cl$^-$ (µg/g)	F$^-$ (µg/g)
1C	48,000	62,000	16,600	12,050	900	450
1D	51,000	66,500	17,800	13,000	1,000	500
2A	48,500	66,000	22,850	12,400	900	500
2B	30,000	42,500	20,400	8,050	600	400
2C	32,000	46,000	40,500	8,600	650	400
2D	35,000	51,000	55,000	9,500	700	1,150
W. Leach (Core Comp.)	52,500	71,500	49,500	13,700	1,050	450
D. Liquid	None	None	None	None	None	None

Note: All IC results are obtained from a water leach preparation.

Table 5-24. Tank 241-C-112 Core 34 Misc. Analyte Trending.

Subsegment	TIC (µg/g)	TOC (µg/g)	Free CN$^-$ (µg/g)	Wt% Total Carbon	pH
1D	7,700	4,900	1,900	1.3	9.77
2B	5,400	3,000	1,600	0.8	9.89
2C	5,200	3,100	1,400	0.8	9.31
2D	6,600	4,000	1,200	1.1	9.72
W. Leach (Core Comp.)	8,600	3,100	2,050	1.2	10.33
D. Liquid	5,600	2,000	1,600	0.7	10.30

Table 5-25. Tank 241-C-112 Core 35 Misc. Analyte Trending.

Subsegment	TIC (μg/g)	TOC (μg/g)	Free CN⁻ (μg/g)	Wt% Total Carbon	pH
2D	3,700	2,500	800	0.6	9.77
W. Leach (Core Comp.)	4,900	1,200	800	0.6	9.77
D. Liquid	4,700	1,400	1,000	0.6	10.47

Table 5-26. Tank 241-C-112 Core 36 Misc. Analyte Trending.

Subsegment	TIC (μg/g)	TOC (μg/g)	Free CN⁻ (μg/g)	Wt% Total Carbon	pH
1C	4,000	8,200	1,100	1.1	9.12
1D	5,400	4,900	1,200	1.0	9.57
2A	4,500	3,900	1,200	0.9	9.54
2B	2,500	2,700	700	0.5	8.92
2C	3,400	2,900	800	0.6	9.29
2D	2,900	2,300	900	0.5	9.36
W. Leach (Core Comp.)	5,300	1,400	1,300	0.7	9.2
D. Liquid	None	None	None	None	None

5.5.1.2 Core 35. IC analyses were performed on drainable liquid and core composite water leach samples. Samples exhibited characteristic poor spike recovery for fluoride (40%). Spike recoveries for the other anions indicate minimal matrix interferences. The water leach sample spike recoveries for NO_3^-, PO_4^{3-}, and SO_4^{2-} were significantly higher than normal (136%, 129%, and 122%, respectively). This behavior was attributed to sample inhomogeneity by the investigator. Control standard recoveries ranged from 83% to 112%; chloride showed a recovery of 132%. Other chloride spike controls run at the same time showed 96% and 92% recovery. Analyses were considered in control at the time of the assays. The drainable liquid samples exhibited good RPDs for all anions, within 5%. Analyte concentrations in the drainable liquid corresponds with the concentrations in the water leach samples.

5.5.1.3 Core 36. IC analyses were performed on subsegment and core composite water leach samples. No drainable liquid was recovered with the core 36 samples. Subsegment 36-1C water leach samples exhibited characteristic poor spike recovery and RPD for fluoride (37% and 22%, respectively). Spike recoveries for the other anions ranged from 80% to 111%, indicating minimal matrix interferences. Control standard recoveries ranged

from 91 to 107%. The analyses were considered in control. Consistent behavior between subsegments was found for NO_2^-, NO_3^-, and SO_4^{2-}, chloride, and free cyanide (i.e., when NO_2 is low, free cyanide is low, etc.), possibly indicating a matrix interference for free cyanide from the other anions. Phosphate and fluoride showed significant differences and do not, in general, track the behavior of the other anions. This is possibly associated with an intrinsic sample inhomogeneity relative to these anions or a heterogeneity associated with tank 241-C-112 itself (i.e., a waste heel). The core composite water leach results magnify the effect of the sample inhomogeneity/tank heterogeneity. This poor reproducibility of sample and duplicate is common for inorganic water leach assays on this sample matrix. RPDs are consistently high for all analytes, ranging from 22 to 43%.

5.6 ANALYTICAL RESULTS--RADIOCHEMISTRY

5.6.1 Radiochemistry Assays--General Comments

Analyses appear to be consistent. Total beta measurements calculated using ^{90}Sr detector efficiencies are largely in agreement with the sum of the major beta emitters, ^{90}Sr and ^{137}Cs. Similarly, the total alpha values show good agreement with the sum of the neptunium, plutonium, and americium/curium values. The gamma energy analyses results obtained from wet chemistry correlate with the results from the segment gamma scans. Detection of ^{137}Cs and most other radionuclides was observed to increase as a function of sample preparation. This was attributed to the ability of the sample preparation to dissolve the waste (KOH fusion dissolves the sample better than acid; acid dissolves the sample better than water). Uranium measurements were obtained from ICP fusion and laser fluorimetry, and show reasonably good agreement between core 34 and 36. Core 35 has a large discrepancy between the two measurements, which can be attributed to matrix interferences or differences in sample particle size. Alpha energy analysis (AEA) and GEA show good agreement for ^{241}Am. GEA analytical values are back-corrected to January 1, 1992, to account for decay.

Tables 5-27, 5-28, and 5-29 show the radionuclide concentrations found in the core composite samples. Tables 5-30, 5-31, and 5-32 show fission product concentration and uranium concentrations as a function of depth.

Table 5-27. Core Composite Fission Products (fusion prep).

Core No.	^{137}Cs (μCi/g)	^{90}Sr (μCi/g)	^{154}Eu (μCi/g)	^{155}Eu (μCi/g)	^{60}Co (μCi/g)	Total Beta (μCi/g)
Core 34	750	3,500	1.25	1.27	0.03	7,070
Core 35	700	3,200	2.01	2.22	2.1 E-04	7,000
Core 36	800	510	0.156	ND	0.006	1,700

ND = Not detected.

Table 5-28. Core Composite Uranium.

Core No.	$U_{ICP\ fusion}$ (μg/g)	U_{FL} (μg/g)	^{238}U mass fraction	^{235}U mass fraction
Core 34	14,400	17,700	0.993107	0.006715
Core 35	89,700	44,300	0.993112	0.006761
Core 36	105,000	94,050	0.993100	0.006780

FL = Uranium measurement by laser fluorimetry.

Table 5-29. Core Composite Transuranics (fusion preparation).

Core No.	^{237}Np (μCi/g)	^{238}Pu (μCi/g)	$^{239/240}Pu$ (μCi/g)	$^{241}Am_{GEA}$ (μCi/g)	$^{241}Am_{AEA}$ (μCi/g)	Total α (μCi/g)
Core 34	6.62E-04	0.0137	0.155	0.76	0.613	0.95
Core 35	1.20E-03	0.0137	0.151	1.05	0.763	1.18
Core 36	4.09E-04	0.0033	0.0593	ND	0.0612	0.17

AEA = Measurement by alpha energy analysis.
GEA = Measurement by gamma energy analysis.
ND = Not detected.

Table 5-30. Tank 241-C-112 Core 34 Radionuclide Trending (fusion preparation).

Subsegment	^{137}Cs (μCi/g)	^{90}Sr (μCi/g)	U_{ICP} (μg/g)
1D	240	1,300	None
2B	610	4,900	3,000
2C	800	1,100	5,700
2D	510	2,500	20,000
W. Leach (Core Comp.)	6.16	27.88*	715
Acid Digestion (D Liquid)	0.04	0.35	1,130

*Value from total beta analysis.

Table 5-31. Tank 241-C-112 Core 35 Radionuclide Trending
(fusion preparation).

Subsegment	^{137}Cs ($\mu Ci/g$)	^{90}Sr ($\mu Ci/g$)	U_{ICP} ($\mu g/g$)
2D	700	3210	89,700
W. Leach (Core Comp.)	5.2	N.M.	460
Acid Digestion (D Liquid)	0.007	0.23	928

Drainable liquids are measured directly.

Table 5-32. Tank 241-C-112 Core 36 Radionuclide Trending
(fusion preparation).

Subsegment	^{137}Cs ($\mu Ci/g$)	^{90}Sr ($\mu Ci/g$)	U_{ICP} ($\mu g/g$)
1C	560	1,900	4,400
1D	1,200	15	3,100
2A	880	20	40,000
2B	530	70	170,000
2C	100	140	110,000
2D	40	200	58,000
W. Leach (Core Comp.)	N.M.	11.8*	4,600
D. Liquid	None	None	None

*Value from total beta analysis.

5.6.1.1 Core 34 Radiochemistry. Mass fractions for isotopic plutonium are
^{239}Pu = 0.93936; ^{240}Pu = 0.0579. The concentration of ^{238}U was too high to determine ^{238}Pu
by MS; thus, AEA of the separated plutonium fraction was used to determine ^{238}Pu. Process
sample blanks show small levels of contamination for selected isotopes. However, the
contamination is negligible when compared to the activity in the solid samples.
Radiochemical recoveries for specific analytes of interest for all samples are as follows: ^{90}Sr
(92 to 105%), uranium (101%), ^{237}Np (94%), plutonium (92%), and americium/curium
(103%).

5.6.1.2 Core 35 Radiochemistry. Uranium measurements from laser fluorimetry had an average chemical recovery of 104%. The neptunium, plutonium, and americium fractions were separated and counted. Plutonium concentration was too low to perform isotopic analysis by MS. Alpha energy analyses were used to determine isotopic plutonium ratios and other alpha emitter concentrations. Problems were encountered in performing the plutonium analysis of the drainable liquid; erratic results and behavior of the sample during analysis was observed.

5.6.1.3 Core 36 Radiochemistry. Mass fractions for isotopic plutonium are ^{239}Pu = 0.95887; ^{240}Pu = 0.03943. The concentration of ^{238}U was too high to determine ^{238}Pu by MS; thus, AEA of the separated plutonium fraction was used to determine ^{238}Pu. Process sample blanks show small levels of contamination for selected isotopes, similar to core 34. The contamination is negligible when compared to the activity in the solid samples. These levels are significant for tritium and ^{60}Co measured by GEA; however, neither of those analytes contribute substantially to the radiological content of the waste. Radiochemical recoveries for specific analytes of interest for reagent or simulated matrix standards are as follows: ^{90}Sr (101%), uranium (107%), ^{237}Np (89%), plutonium (96 plutonium), and americium (100%).

5.7 ANALYTICAL RESULTS--ENERGETICS

Scanning TGA and DSC were performed on subsegment and core composite material obtained from tank 241-C-112. These two thermal analysis techniques are useful in determining the thermal stability or reactivity of a material. In DSC analysis, heat flow over and above the usual heat capacity of the substance is measured while the substance is exposed to a linear increase in temperature, i.e., dT/dt = Constant (where T=temperature, and t=time). While the substance is being heated, air is passed over the waste material to remove any gases being released. The onset temperature for an endothermic or exothermic event on a DSC is determined graphically. The endpoints of the event are determined and a line is drawn between them to establish a base. A line tangent to the initial side of the event is drawn until it intersects the base. From that point of intersection, a vertical line is constructed to the temperature scale at the bottom of the DSC curve; that temperature is the onset temperature of the event.

TGA measures the mass of a sample while the temperature of the sample is increased at a constant rate. Again, dT/dt is constant; the X-axis is representative of the running time of the analysis as well as the temperature increase of the sample during analysis. The Y-axis represents the weight percent of the sample and is effectively unitless. As with the DSC, air is passed over the sample during heating. Any decrease in the weight percent of the sample represents a loss of gaseous matter from the sample either through evaporation or through a reaction that forms gas phase products.

DSC is often used to measure thermal decomposition temperatures, heats of reaction, reaction temperatures, melting points, and solid-solid transition temperatures. TGA is used to measure thermal decomposition temperatures, water content, and reaction temperatures. The two methods often provide complementary information.

5.7.1 Remarks on the Interpretation of Differential Scanning Calorimetry/Thermogravimetric Analysis Data

Tables 5-33 and 5-34 summarize the results of the thermal analyses performed. There are three significant features seen on both the DSC and TGA plots. Several minor endotherms appear on the DSC, but only endotherms greater than 20 J/g were considered as significant features of the plot. The values presented in the tables do not exactly match the values recorded on the DSC and TGA plots. This is because interpreting these semi-quantitative analyses requires considerable experience and judgement on the part of the analyst. The values presented represent the best summary evaluation of the data (Tingey 1993). Although the temperature ranges observed for the various transitions in the DSC and TGA assays do not exactly match, the weight losses and thermal events in the observed transitions are considered related and usually in the same vicinity.

Table 5-33. Thermogravimetric Analysis Results
from Tank 241-C-112.

Core Sample	Total Wt% loss	Transition 1 Wt% loss	Transition 2 Wt% loss	Transition 3 Wt% loss
34-1D	20.2	6.6	13.6	N.A.
34-2B	56.8	52	5.4	-0.6
34-2C	48.7	45	4.0	-0.3
34-2D	39.3	33	6.3	0
34-Comp.	40.5	35	6.1	-0.6
35-2D	48.0	42	6.0	0
36-1C	54.6	46	8.9	-0.3
36-1D	51.8	52	2.0	-0.2
36-2A	53.7	52	1.9	-0.2
36-2B	41.0	38	3.1	-0.1
36-2C	44.8	41	3.8	0
36-2D	50.8	47	3.5	0.3
36-Comp.	54.2	44	2.9	0.3

Table 5-34. Differential Scanning Calorimetry Energetics Results
from Tank 241-C-112.

Core Sample	Transition 1			Transition 2			Transition 3		
	Range (°C)	Onset (°C)	ΔH (J/g)	Range (°C)	Onset (°C)	ΔH (J/g)	Range (°C)	Onset (°C)	ΔH (J/g)
34-1D	58-110	62	30	206-312	215	515	356-444	375	-17
34-2B	30-240	30	847	260-300	276	-12	300-400	349	
34-2C	30-240	30	795	260-300	267	-13	300-400	360	
34-2D	33-240	33	930	260-300	289	-17	300-400	347	
34-Comp.	34-240	34	734	260-300	276	-11	300-400	357	
35-2D	34-195	34	780	225-290	230	-12			
36-1C	34-240	34	1,070	260-300	267	-11	300-380	301	31
36-1D	32-230	32	1,310	260-310	277	-16	(a)	NA	
36-2A	30-230	30	1,110	277-300	280	-10	300-400	305	35
36-2B	33-235	30	870	260-325	298	-9	325-400	330	28
36-2C	33-240	32	830	(a)	NA	None	305-407	320	36
36-2D	34-172	34	1,060	(a)	NA	None	300-395	328	45
36-Comp.	34-172	34	880	270-325	288	-19	(a)	NA	

(a) This transition is not quantifiable.
NOTE: To convert from J to cal, divide by 4.18.
NOTE: Negative ΔH indicates an exotherm.
NA = Not applicable.

There is a concern regarding the choice of cover gas affecting the DSC and
TGA results. Air was used in the assays instead of an inert gas because that is what the test
instructions directed. However, oxygen in the air may contribute to the oxidation of the
sample and alter the reaction (Pederson et al. 1993). This condition is not considered
representative of the potential reaction conditions in the tank; therefore, future DSC/TGA
tests will be performed under an inert cover gas.

5.7.2 General Comments on the Differential Scanning Calorimetry/Thermogravimetric Analysis Behavior of the Samples

The first transition in each sample is endothermic, begins at the lower temperature limit of the analysis (30 °C), and is essentially complete between 140 °C and 200 °C. The most likely phenomenon occurring in this region are the release of the bulk and interstitial water in the core sample material. The endotherms exhibited in this region are quite substantial (typically, 700 to 1,200 J/g). These values are per gram of wet sample; if divided by the mass fraction lost during analysis, they range from 1,400 to 2,600 J/g and correspond generally with the heat of vaporization of water (2,260 J/g). In addition, the majority of the weight percent change observed in the TGA curve occurs over this same temperature range. The TGA water content corresponds closely (but not exactly) with the water loss observed in a gravimetric weight percent solids determination. Some of this discrepancy is attributable to the time elapsed between the two assays. In some cases, up to two months pass between the gravimetric weight percent solids assay and the TGA measurement. Extended exposure to the ambient hot cell conditions are believed to have dried the sample somewhat in that period of time; thus, some of the water content measurements may be biased low. The warm, dry conditions in the hot cell will remove moisture between the time of the sample assays. However, the results from the two methods are generally in agreement (core 35 was an anomaly).

Exotherms and additional weight loss are routinely detected between 260 to 300 °C in all the samples. Similar exotherms and weight changes have been observed in previous thermal analysis studies of $Cs_2NiFe(CN)_6$ (Scheele et al. 1991) and other simulant materials (Bechtold 1992; Jeppson 1993). As reported previously, the dried simulant materials demonstrate much larger exothermic responses than those observed in tank 241-C-112 waste. However, the magnitude of the exotherms observed correlates roughly with the predicted exotherms derived from the amount of cyanide present in the waste (refer to Table 5-35 and 5-36), based on the Fauske (1992) determined value of -3.95 kJ/g $Na_2NiFe(CN)_6$. The weight losses are attributed to the loss of gaseous reaction products and waters of hydration.

There is an observable third transition range, but here the energetic behavior is not readily quantifiable. Initially, Cores 34 and 35 appeared to have some significant exothermic behavior detected in the temperature range between 300 to 440 °C. In the core 36 materials, the investigator detected no exothermic activity in this temperature range. This difference in behavior was unexpected, especially in light of small, but detectable, cyanide and organic levels in the waste, suggesting the potential for a corresponding exotherm. Further analysis of the DSC results suggests that the observed activity was the consequence of subtracting a rapidly changing baseline measurement from the analytical results. Therefore, it is believed that there is no quantifiable exothermic activity in that temperature range. This interpretation is contrary to previously released information. A minor weight gain was sometimes observed in this temperature range, most likely an oxidation of some type.

Table 5-35. Tank 241-C-112 Core 34 Energetic Comparison.

Subsegment	Wt% Total Cyanide (dry)	Equivalent Wt% Total Ferrocyanide (dry)	Theoretical Heat of Reaction (cal/g dry waste)	Measured Heat of Reaction (cal/g dry waste)
1D	0.52	1.06	-10.0	-5.1
2B	0.43	0.87	-8.3	-6.7
2C	0.83	1.69	-15.9	-6.1
2D	0.75	1.52	-14.4	-6.7
Composite	0.97	1.97	-18.6	-4.4

NOTE: 1 cal = 4.18 J.

Table 5-36. Tank 241-C-112 Core 36 Energetic Comparison.

Subsegment	Wt% Total Cyanide (dry)	Equivalent Wt% Total Ferrocyanide (dry)	Theoretical Heat of Reaction (cal/g dry waste)	Measured Heat of Reaction (cal/g dry waste)
1C	NM	NM	NM	-5.8
1D	0.72	1.46	-13.8	-8.0
2A	0.92	1.87	-17.7	-5.2
2B	0.75	1.52	-14.4	-3.7
2C	0.40	0.81	-7.7	No exotherm
2D	0.56	1.14	-10.8	No exotherm
Composite	0.71	1.44	-13.6	-8.6

NM = No measurement.
NOTE: 1 cal = 4.18 J.

The properties related to energetics are illustrated for each core in Tables 5-37, 5-38, and 5-39. The results for the samples from 34-1D, 36-2C, and 36-2D indicate that these samples differ in thermal behavior from most of the other samples, further suggesting a difference in waste type.

The TOC and TIC assays are not considered capable of measuring the total cyanide in the waste because they depend on acid dissolutions to perform the analyses.

Table 5-37. Tank 241-C-112 Core 34 Energetics Trending.

Subsegment	Wt% Total Cyanide (dry)	Wt% Total Organic Carbon	Wt% Total Carbon	Wt% Water (Grav.)	Wt% Water (TGA)	Average Heat of Reaction (kJ/g dry waste)
1D	0.52	0.49	1.3	45	20	-0.02
2B	0.43	0.30	0.8	53	57	-0.03
2C	0.83	0.31	0.8	58	49	-0.03
2D	0.75	0.40	1.1	52	39	-0.03
Composite	0.97	0.31	NM	38	41	-0.02

Table 5-38. Tank 241-C-112 Core 35 Energetics Trending.

Subsegment	Wt% Total Cyanide (dry)	Wt% Total Organic Carbon	Wt% Total Carbon	Wt% Water (Grav.)	Wt% Water (TGA)	Average Heat of Reaction (kJ/g dry waste)
2D	NM	0.25	0.6	34	48	-0.02

Table 5-39. Tank 241-C-112 Core 36 Energetics Trending.

Subsegment	Wt% Total Cyanide (dry)	Wt% Total Organic Carbon	Wt% Total Carbon	Wt% Water (Grav.)	Wt% Water (TGA)	Average Heat of Reaction (kJ/g dry waste)
1C	N.M	0.82	1.1	49	55	-0.02
1D	0.72	0.49	1.0	58	52	-0.03
2A	0.92	0.39	0.9	57	54	-0.02
2B	0.75	0.27	0.5	41	41	-0.02
2C	0.40	0.29	0.6	64	45	No exotherm
2D	0.56	0.23	0.5	56	51	No exotherm
Composite	0.71	0.14	0.7	45	47	-0.04

Heats of Reaction are calculated using the TGA wt% water value.
NOTE: 1 kcal = 4.18 kJ.

5.8 ANALYTICAL RESULTS - POTENTIAL WASTE CONSTITUENTS

5.8.1 Mass Balances

A method to help ensure that the data are acceptable is to perform a mass balance on the core composite sample data. This activity functions as a rough QC check, and also provides insight to some of the properties of the matrix. To do this, the assumption in performing the mass balance is that the anions, cations, and water are all associated in some manner, but the exact chemistry of the association is not considered. Analytes contributing less than 0.2 wt% (generally trace ICP analytes, AA analytes, and radionuclides) are considered negligible in this assessment. The assays that contributed analytes to the mass balance were the ICP fusion, IC, total carbon (TC), total cyanide assays, and the gravimetric wt% water measurement. The ICP fusion value does not include nickel, which is a significant analyte in the sample but may be biased high. However, for the purpose of this exercise, the nickel value from the respective acid leach preparations will be inserted into the total mass of ICP fusion analytes to account for it.

Without considering the physical and chemical properties of the waste matrix and the context of the process history, the mass balances produced from these assays account for 77.2 to 97.9% of the mass. However, this range of recovery is expected because it is known that there are analytes present that were not measured in the analysis of the samples. The IC anion analysis only measures the water-soluble components; there is a substantial insoluble residue that must contain additional anions. There is no measurement of the sulfide content in any of the assays, even though it has been previously established that 28,100 g-mol of S^{-2}, was used in scavenging ^{60}Co. Thus, an additional contribution of 2,100 $\mu g/g$ has been estimated as necessary to aid in closing the balance. Bismuth was not reported in the assays, and $BiPO_4$ first cycle waste was recorded as being disposed here, introducing a potential shortfall.

Aluminum is likely to be present as $Al(OH)_3$, and other transition metals are also likely to be present as hydroxides or hydrous metal oxides. Neither hydroxide ion or oxide content has been measured in the waste, which introduces additional sources of shortfalls in the recovery. Therefore, multipliers for aluminum (2.9), iron (1.6), nickel (1.6), and uranium (1.3) will be used to account for the unmeasured hydroxide or oxygen, which are assumed to be present in combination with these analytes. Only metals making weight percent contributions to the waste matrix will be adjusted in this manner; the trace metals will be assumed to be lost in the error of the major constituents. Adjustments will be made individually for TOC, TIC, and total cyanide. It is assumed that the TOC and TIC assays did not consume or measure any cyanide present. In addition, a significant disparity can be corrected by comparing the soluble phosphorus from the water leach ICP (and assuming that it is PO_4^{3-}), PO_4^{3-} values from the IC, and the phosphorus from the ICP fusion assay; the phosphate was found to be only 27 to 44% soluble. The water leach ICP and IC values agree within 3%, strongly suggesting that the soluble phosphorus in the waste matrix is present as PO_4^{3-}. The process history of the tank also indicates that large amounts of

phosphate were used to encourage precipitate formation. Therefore, an assumption that the phosphorus in the fusion assay represents an insoluble PO_4^{3-} is not unwarranted. Convert the phosphorus in the ICP assay to PO_4^{3-} and add it along with the water soluble phosphate and other anions. A minor accounting shift is now necessary to avoid double counting (subtract the ICP fusion phosphorous value).

Accounting for the analytes in this manner aids in closing agreement and improving the percent recoveries to between 96.6% and 119.3% (near quantitative recoveries). However, there remain some aspects of the waste matrices that require examination. Tables 5-40, 5-41, and 5-42 present mass balances that have been adjusted to compensate for the contributions of unmeasured (but likely) analytes combined with the measured analytes. There may have been some error introduced from drying of the sample during the preparation of the core composite, especially in the case of core 34, where the subsegments have a substantially higher water content than does the core composite, biasing the results low. In the case of core 35, the disparity between the gravimetric water measurement and the TGA water content also suggests (1) drying of the sample before the gravimetric assay; or (2) incomplete drying during the gravimetric test, which biases the results low.

Table 5-40. Adjusted Mass Balance: Core 34 Composite.

Assay	Concentration (μg/g)
ICP--Fusion (+ Ni from acid leach; -P; Al, Fe, Ni, U, adjustments)	329,000
IC Anions (TOC, TIC, and CN⁻ adjustments; +P as PO_4^{3-}; +S^{2-})	306,000
Gravimetric Water	380,000
Total (1,000,000 μg/g)	1,015,000

Table 5-41. Adjusted Mass Balance: Core 35 Composite.

Assay	Concentration (μg/g)
ICP--Fusion (+ Ni from acid leach; -P; Al, Fe, Ni, U, adjustments)	414,000
IC Anions (TOC, TIC, and CN⁻ adjustments; +P as PO_4^{3-}; +S^{2-})	212,000
Gravimetric Water	340,000
Total (1,000,000 μg/g)	966,000

Table 5-42. Adjusted Mass Balance: Core 36 Composite.

Assay	Concentration (μg/g)
ICP--Fusion (+ Ni from acid leach; -P; Al, Fe, Ni, U, adjustments)	404,000
IC Anions (TOC, TIC, CN⁻ adjustments; +P as PO_4^{3-}; +S^{2-})	339,000
Gravimetric Water	450,000
Total (1,000,000 μg/g)	1,193,000

5.8.2 Suggested Components of Waste Matrix

The actual composition of the waste matrix is quite complex, and trace amounts of various compounds probably exist in the tank. However, with some simple assumptions regarding how the anions and cations will combine, a list of the most probable compounds that exist in the waste matrix and contribute significantly to its overall makeup can be developed.

Table 5-43 is a condensed version of a more general chart found on page D-147 in the *Handbook of Chemistry and Physics 64th Ed.* (Weast 1984). It provides solubility data on some of the most common anions and cations.

Table 5-43. Probable Solids in the Waste Matrix.

	NO_2^-	NO_3^-	PO_4^{3-}	SO_4^{2-}	OH⁻	CN⁻ (as Fe(CN)$_6^{4-}$)	CO_3^{2-}	S^{2-}
Al^{+3}			PPT		PPT	PPT		NL
Ca^{+2}			PPT	PPT			PPT	
$Fe^{+2,+3}$					PPT	PPT	PPT	PPT
Na^+								
Ni^{+2}			PPT		PPT	PPT	PPT	PPT
U^{+6}		NL	PPT	PPT	PPT			PPT

PPT = Precipitate forms.
NL = Precipitate formation not likely under tank conditions.

From the earlier tables and process information, chloride, fluoride, sulfide, and even ferrocyanide will not be significant mass contributors to the waste matrix. Sulfide and cyanide precipitates are significant because they provide a potential fuel source; however, it is generally believed that the sulfides were eventually converted to sulfates. Sodium, NO_2^-,

and NO_3^- are highly soluble, and thus probably do not contribute much to the insoluble solids. However, sodium, nitrite, and nitrate contribute significantly to the overall solids content of the waste (dissolved + insoluble solids). In addition, they represent three of the four most prevalent analytes, after water, in the waste. No analytical measurement of hydroxide was made, but it is known that in the process history of tank 241-C-112, basic solutions were added routinely to the tank. The following is a list of likely candidates for the insoluble solids:

- Aluminum hydroxide, $Al(OH)_3$
- Aluminum phosphate, $AlPO_4$
- Tetraaluminum ferrocyanide, $Al_4[Fe(CN)_6]_3$
- Calcium phosphate, $Ca_3(PO_4)_2$
- Calcium sulfate, $CaSO_4$
- Calcium carbonate, $CaCO_3$
- Iron carbonate, $FeCO_3$
- Iron(II) hydroxide, $Fe(OH)_2$
- Iron(III) hydroxide, $Fe(OH)_3$
- Iron(II) phosphate, $Fe_3(PO_4)_2$
- Iron(III) phosphate, $FePO_4$
- Iron sulfide, FeS
- Iron(III) ferrocyanide, $Fe_4[Fe(CN)_6]_3$
- Disodium nickel ferrocyanide, $Na_2NiFe(CN)_6$
- Nickel carbonate, $NiCO_3$
- Nickel sulfide, NiS
- Nickel hydroxide, $Ni(OH)_2$
- Dinickel ferrocyanide, $Ni_2Fe(CN)_6$
- Uranyl phosphate, $UO_2HPO_4 \bullet 4H_2O$
- Uranyl hydroxide, $UO_2(OH)_2$
- Uranyl sulfide, UO_2S
- Uranyl sulfate, $2(UO_2SO_4) \bullet 7H_2O$.

The ^{137}Cs present is still apparently bound with the ferrocyanide, and the ^{90}Sr is probably bound with phosphate, carbonate, or sulfate.

5.8.3 Comparison to Theoretical Estimates and Simulant Studies

Agreement between synthetic sludge properties and observed waste material characteristics is within the constraints of the synthetic recipes and assumptions regarding chemical behavior in tank 241-C-112. Comparisons with core 35 were not made because of its small sample size. On this basis it was assumed that valid comparisons could not be made, and that core 35 was not a representative sample of the tank waste. Table 5-44 compares some properties and analyte concentrations of the waste materials and comparable simulants.

Table 5-44. Tank 241-C-112 Comparison of Waste Material with Simulants for Selected Analytes.

Analyte	In-Farm 2 simulant values	Core 34 values subsegment range (composite value)	Core 36 values subsegment range (composite value)
Ni ug/g*	18,700	22,200 to 28,300 (30,000)	900 to 23,000 (12,800)
Wt% H_2O Gravimetric	51	45 to 58 (38)	41 to 64 (45)
Wt% Total Cyanide dry basis	9.1 to 11.3	0.43 to 0.83 (0.97)	0.40 to 0.92 (0.71)
ΔH kJ/dry g	-1.2	-0.01 to -0.03 (-0.02)	-0.01 to -0.03 (-0.04)
Density g/ml	1.39**	Bulk Value: 1.5 to 1.6	Bulk Value: 1.3 to 1.6

*Ni analysis is potentially biased high. Values are derived from ICP fusion performed in a Ni crucible. However, in each case the blank value was an order of magnitude (or greater) less than the measurement. The fusion values are comparable with acid digestion values, where both are available.

**Centrifuged for 30 g-yr.

5.9 RCRA-TYPE ANALYSIS: DATA VALIDATION/ VERIFICATION PROTOCOL

5.9.1 Chemical Data

Data validation procedures for chemical data were in place during the analysis of tank 241-C-112. The data validation and verification procedures followed to ensure reliable data for *Resource Conservation and Recovery Act of 1976* (RCRA)-type samples are described in detail in WHC-CM-5-3, *Sample Management and Administration*, Section 2.0. A brief list of the requirements for data packages are as follows:

- Requested versus Reported Analyses
- . Initial and Continuing Instrument Calibration
- Analysis Holding Times
- Gas chromatograph (GC)/MS Tune Criteria (GC/MS analysis)
- Matrix Spike/Matrix Spike Duplicate Analysis

- Surrogate Recoveries
- Duplicate Analysis
- Internal Standards (GC analysis)
- Analytical Blank Analysis
- Laboratory Control Samples (LCS)
- Additional Quality Assurance (QA)/QC Oversight, as designated in the Statement of Work (SOW)
- Interference Check Sample.

When determining the quality of the chemical data for tank 241-C-112, it is useful to consider the results of several of the QC assays. Blanks, matrix spikes, duplicates, and control samples can all provide further insight to the data and its reliability. Potential sample contamination problems are addressed using analytical blanks. Confounding effects of the sample matrix are resolved using matrix spike results. Duplicate analysis compares the difference between the replicate samples, providing an indication of laboratory precision (and in some cases, sample heterogeneity). The laboratory control sample offers a monitor of overall performance of an analytical method in all steps of the analysis. Overall, there were few problems with the data validation and compliance with established QC criteria. The 241-C-112 samples were generally free from calibration and contamination errors (core 36 was an exception to this). In addition, the interference control standards, matrix spike, laboratory control standard, and holding time requirements were largely met. Most of the observed out-of-control or out-of-limit problems were observed in the blank and duplicate assays of the trace analytes. In some cases, the sample behavior of the duplicate assays was attributed to significant sample heterogeneity. In almost all cases, the data obtained was qualifiable and usable. HASM (1993) contains the core sample data output and the associated QA documentation.

5.9.2 Radiological Data

Similarly, the data validation and verification procedures followed to ensure reliable data for radioactive, high-level, RCRA-type samples are also described in detail in WHC-CM-5-3, *Sample Management and Administration*, Section 2.4. They differ somewhat from the requirements for chemical data. A brief list of the requirements for data packages are as follows:

- Chain-of-Custody
- Laboratory Control Samples (LCS)
- Requested vs. Reported Analyses
- Initial Calibration
- Efficiency Checks
- Preparation Blank Analysis
- Matrix Spikes/Tracers/Carriers

- Duplicate Analysis
- Additional QA/QC Oversight, as designated in the SOW
- Background Checks.

When determining the quality of the radiological data for tank 241-C-112, it is also useful to consider the results of several of the QC assays. Chain-of-custody, calibrations, efficiency and background checks, blanks, matrix spikes/tracers/carriers, duplicate analyses, and LCSs can all provide further insight to the data and its reliability. Potential sample contamination problems or loss of sample control are addressed in using a chain-of-custody. Intial calibrations, efficiency and background checks, and analytical blanks ensure that the equipment is operating correctly and further address contamination problems. Confounding effects of the sample matrix are resolved using matrix spike/tracers/carriers. Duplicate analysis compares the difference between the replicate samples, providing an indication of laboratory precision (and in some cases, sample heterogeneity). The laboratory control sample offers a monitor of overall performance of an analytical method in all steps of the analysis. In the case of the radiological data, there were several problems with the data validation and compliance with established QC criteria. The 241-C-112 chain-of-custody documentation was in order, thus sample integrity was not lost. In addition, the sample blank results indicated that samples were generally free from contamination errors. However, there are several problems in compliance with the established QC criteria for initial calibrations, efficiency checks, matrix spike/tracers/carriers, and the use of laboratory control standards for these samples. These concerns were addressed and responded to in the PNNL Technical Task Plan (TTP), which outlined the alternative QC criteria that would be adhered to during the analysis of the samples. WHC agreed to the criteria outlined in the PNNL TTP and a formal audit response has been issued to clarify the matter further. In almost all cases, the radiological data obtained was qualifiable and usable. HASM (1993) contains the core sample data output and the associated QA documentation.

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6.0 INTERPRETATION OF ANALYTICAL RESULTS

6.1 TANK 241-C-112 WASTE PROFILE

Tank 241-C-112 received four major types of waste likely to deposit solids during its operating history. The waste types, in order, were as follows:

- Bismuth phosphate (1C)

- Unscavenged UR waste

- Ferrocyanide-scavenged waste (FeCN) from scavenging of UR waste initially stored in other tanks

- Cladding wastes (CW).

A small volume of strontium semiworks waste subsequently was received by the tank and a relatively large volume of B Plant ion-exchange waste after that. These last wastes would not be expected to contribute large amounts of solids to the tank. The purpose of this section is to attempt to identify the location of the tank waste solids, thereby allowing estimation of the tank inventory for various analytes of importance.

The approach taken to identify the waste profile was to examine the subsegment assays for analytes distinct to the waste types that were disposed in the tank, and combine that information with what is known regarding the tank's process history. The first waste placed in the tank, via the cascade inlet from tank 241-C-111, was $BiPO_4$ 1C waste. This waste would be comparatively high in bismuth, phosphate, and aluminum because aluminum decladding waste was combined with it. The 1C solids volume was measured as 57,000 L (15,000 gal) in 1952 (Anderson 1990). This volume would amount to 32.8 cm (12.9 in.) in the tank bottom. The tank next received UR waste with solids comparatively high in uranium and iron and low in bismuth and aluminum. The available records do not show whether this waste was added directly to the tank or via the cascade overflow line from tank 241-C-111. Assuming direct addition to tank 241-C-112, the estimated volume of these solids would be approximately 55,000 L (14,500 gal) at 1.8 vol% of the waste input. This corresponds to 13.5 cm (5.3 in.) distributed evenly across the tank.

The tank then received ferrocyanide-scavenged waste. The solids from this waste would be high in nickel, calcium, cyanide (as ferrocyanide), ^{137}Cs, and uranium, although the uranium may have settled out in the tank originally receiving the UR waste. Because some of the ferrocyanide waste feed was concentrated 1C waste (evaporator bottoms), the waste could also be high in aluminum. The estimated solids volume in tank 241-C-112 at the end of the scavenging program was between 174,000 and 318,000 L (46,000 and 84,000 gal). This would amount to 61.5 to 96.5 cm (24.2 to 38.0 in.) distributed evenly across the tank. The last major waste type was aluminum CW. These materials would be high in aluminum

and silica; however, the solids volume is unknown because the majority of the solids would be deposited in the first tank to receive the wastes, which was not tank 241-C-112. The grey/white solids seen in the video recordings of the core extrusions are believed to be CWs; the tan/dark brown solids are thought to be ferrocyanide sludge. Their observed position during extrusion agrees with the historical record. The volume of strontium semiworks waste was small and probably would not have been visually detectable. However, it would have a relatively high ^{90}Sr content because it includes strontium recovery and purification waste losses. This characteristic would be readily observable in the radiochemistry analyses.

6.2 REVIEW OF THE SUBSEGMENT ANALYTE PROFILES

The following conclusions are drawn from review of the subsegment analyses presented in Section 5.0.

Core 34 -- Chemical analyses indicate this material is ferrocyanide waste. The relatively high nickel, calcium, TOC, and ^{137}Cs support this conclusion. In addition, the total cyanide analysis indicates residual cyanide in the waste. The DSC traces show a small exotherm in the same temperature range where the In-Farm simulants begin to show reactions; however, the overall energy release is somewhat lower than expected for the measured cyanide concentration and much lower than that expected from simulant information. The high uranium value in subsegment 2D is believed to be due to some UR waste also being present.

Core 35 -- The only solids recovered from this core were 7.6 cm (3.0 in.) of solids from the lower segment (assumed to be subsegment 2D); thus, the waste origin is indeterminate. Some analytes (e.g., nickel, aluminum, calcium, and phosphate), match core 34 well. Other analytes such as nitrate, nitrite, uranium, and TOC/TC values agree better with core 36. No total cyanide measurements are available.

Core 36 -- Subsegments 1C and 1D show high nickel, calcium, TOC, and ^{137}Cs analyte concentrations characteristic of ferrocyanide wastes. Subsegments 2C and 2D are relatively low in these analytes but high in uranium and phosphate, which is typical of UR wastes. The 2A and 2B subsegments appear to be a mix of these two types of waste. The total cyanide values trend generally lower as a function of waste depth. In subsegments 1C, 1D, 2A, and 2B, the DSC traces show a small exotherm in the same temperature range where the simulants show reactions; again the exotherms were much lower than expected. Although some measurable cyanide is in the lower subsegments, no corresponding exotherm is recorded in the DSC measurements. High TOC and ^{90}Sr values in 36-1C may indicate hot semiworks/strontium semiworks, which would have had organic complexants.

6.2.1 Tank Entrance/Exit Effects on Analyte Distribution

Figure 6-1 shows an elevation and plan of where the core samples were taken. Important items to note are that core 36 (and 35) were taken from risers near the ferrocyanide waste inlet, while core 34 was obtained near the waste pumpout riser. The decant "float and flex" pump contained a 6.1-m (20-ft) section of flexible hose that could traverse a relatively wide area on that side of the tank. The cascade fill line (where $BiPO_4$ 1C waste entered the tank) is closer to the core 34 sample point than to the core 35 and 36 risers. The elevation view shows this is a shallow-dish bottom tank and the bottom of the core samples were 33 cm (13.0 in.) above the centerline inside bottom of the tank. Also shown is the waste surface, measured from the centerline as 115.1 ± 1.3 cm (45.3 ± 0.5 in.). The total sample and solids recovery for each of the three cores is shown as well as the top and bottom of the sample segments. If the solids recovery in the bottom of segment core 34, is adjusted to equal core 36, there is a 18.3-cm (7.2-in.) decrease in the top solids surface going from riser 8 across the tank to riser 2 (18.3-cm [7.2-in.] decrease in approximately 20.1 m [66 ft]). Figures 6-2A and 6-2B shows a representation of the overall waste profile of tank 241-C-112 and the assumed volumes, boundaries, and positions of the various individual layers as they are believed to exist.

As new wastes entered the tank and distributed themselves across the tank, the material under and around the tank pumpout (core 34) would be routinely disturbed (and occasionally solids transferred) in behavior akin to the last in-first out principle. However, the material beneath the waste inlet (cores 35 and 36) would be disturbed initially but, over time, large stratified layers resistant to mixing would eventually build up. Thus the influence of the waste inlet and outlet locations provides insight to the contrast in the analyte and waste profiles between cores 34 and 36.

The ^{137}Cs concentrations are nearly uniform in the core composites. However, the ^{137}Cs concentration as a function of depth shows profiles consistent with the wastes believed to be associated with the subsegments--low ^{137}Cs values for unscavenged UR wastes, higher ^{137}Cs values for ferrocyanide wastes. In addition, the ^{137}Cs profile was relatively uniform in core 34; it shows a decreasing trend as a function of depth in core 36. The ^{90}Sr concentration, however, shows a highly skewed distribution between cores 34 and 36 and unusual concentration profiles as a function of depth for both cores. Core 34 has extremely high ^{90}Sr concentration throughout; core 36 is high in the top subsegment and then the concentration falls dramatically. The consistently high ^{90}Sr values in core 34 are believed to be a localized phenomenon and not representative of the ^{90}Sr concentration in the tank. This observation is supported by heat-load estimates developed from thermal models of the temperature profiles in the tank. In addition, no mechanism has been demonstrated that has the capability to concentrate radionuclides to levels of concern (Dickinson et al. 1993). The temperature profile model results give a heat load of 2.37 ± 0.06 kW (Simpson et al. 1993, Appendix C). A preliminary heat-load estimate developed from the core 34 cesium and strontium inventories representing the bulk analyte concentrations is 5.63 kW. This value is not realistic, given the temperature measurements of the tank waste. Another preliminary estimate, calculated from core 36 values (2.33 kW), is much more in agreement with the

heat-load model value for tank 241-C-112. Core 35 does not have enough waste to provide a basis for a meaningful comparison. Using the model heat load as a basis, a realistic ^{90}Sr concentration/inventory may be developed. This will be shown in Section 6.3. The high localized ^{90}Sr values may be attributable to the location of the core samples relative to the waste inlet and outlet.

One of the upper subsegments of core 36 (36-1C) was relatively high in aluminum, which is typical for CWs that were deposited on top of the ferrocyanide wastes. Aluminum shows similar distribution behavior to ^{90}Sr; a high and relatively uniform concentration in core 34, and in core 36, a much lower concentration that decreases as a function of depth. This may be reasonable because the same tank conditions and disturbances would have affected later waste transfers. As noted previously (Borsheim and Simpson 1991), some large transfers from tank 241-C-112 in the 1970's probably transferred some of the uppermost solids from the tank. None of the core sample subsegments demonstrated convincing evidence of being 1C waste solids. Unfortunately, the bismuth values that distinguish this waste were not determined via the ICP fusion analysis. It is expected that the bulk of the BiPO$_4$ 1C waste lies below the depth that can be core sampled through the available risers.

6.2.2 Comparisons with the Borsheim/Simpson Model Estimates

Calculations of the ^{137}Cs, nickel, and Fe(CN)$_6^{4-}$ inventories are analytes appropriate for comparison with the model. Assumptions regarding the tank used in the calculations for the analytical estimates, and the calculations themselves, are presented in Simpson et al. (1993). Table 6-1 presents comparisons of the calculated values with the original and revised Borsheim/Simpson values after scavenging was finished.

Table 6-1. Comparisons of Initial and Revised Borsheim/Simpson Model Estimates with Values Calculated from Analytical Results.

Analyte	Borsheim/Simpson		Revised Borsheim/Simpson		Analytical Estimates
	Retained	Input	1.0 (retained) vol%	1.5 (retained) vol%	(from Section 6.3)
Ni, moles	31,000	78,500	98,800[1]	93,300[1]	95,200
^{137}Cs, kCi (decayed to 1993)	124.8	334.8	224.7	205.9	217.2
Fe(CN)$_6^{-4}$, moles	31,000	78,500	65,900	60,400	8,700

[1]Includes the ^{60}Co scavenging contribution.

Figure 6-1.   Elevation and Plan of Tank 241-C-112.

29304019.5
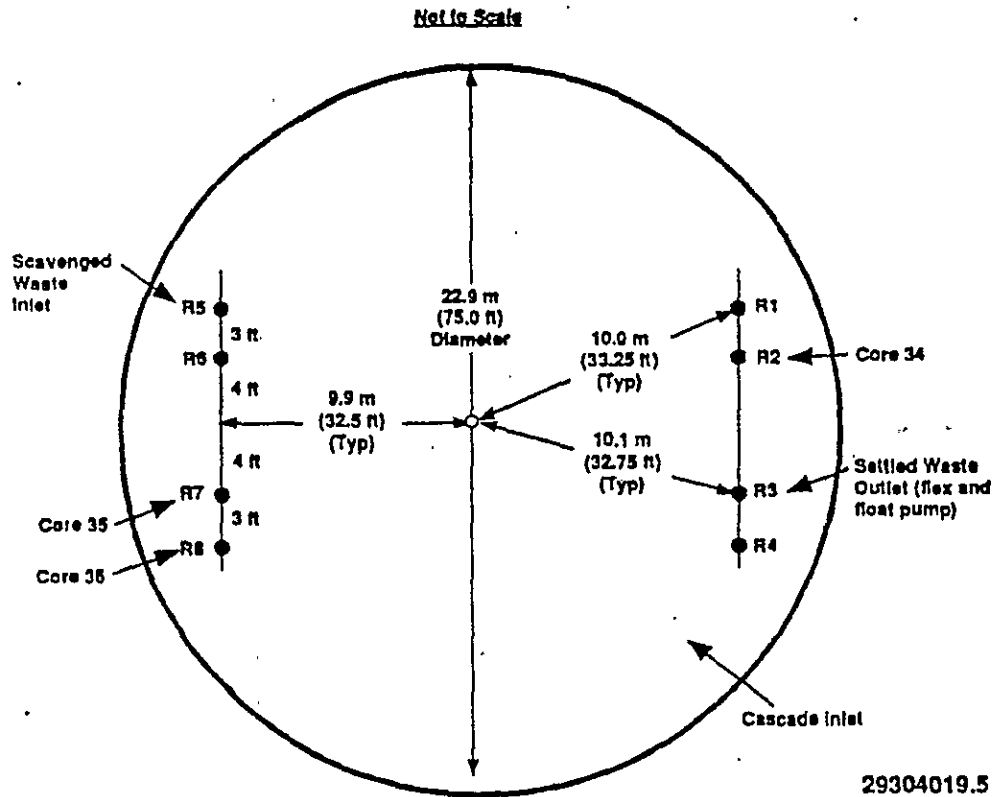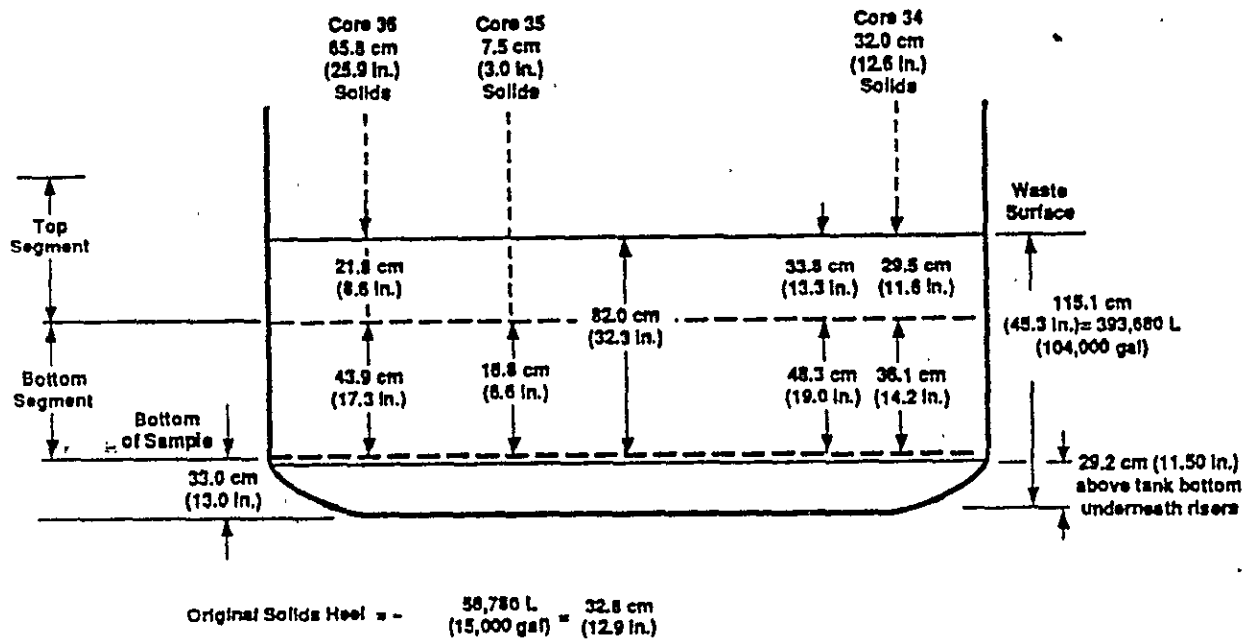
Figure 6-2A. Waste Profile of Tank 241-C-112.
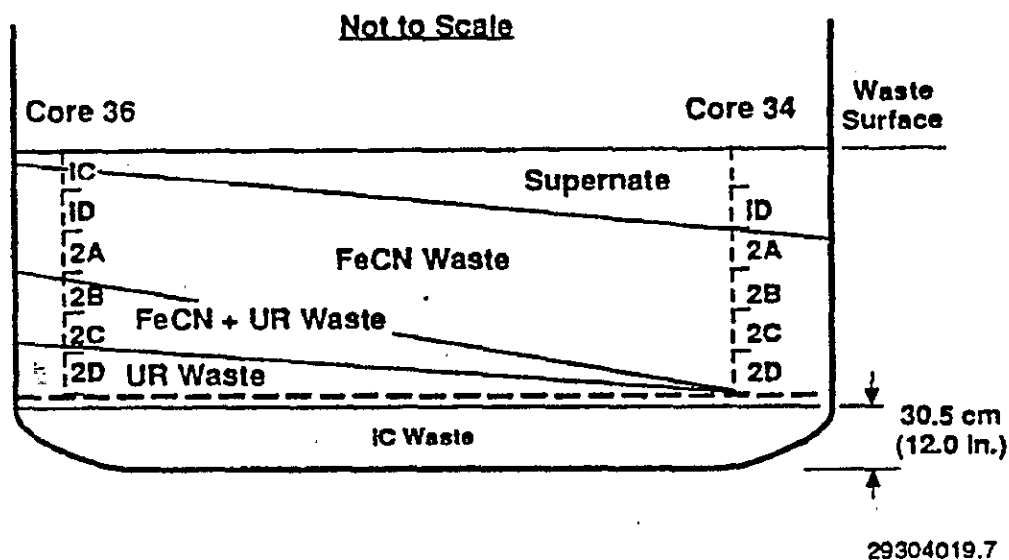
29304019.7

- Dished Bottom: First cycle BiPO$_4$ waste or unscavenged uranium recovery (UR) waste 57,700 L (15,000 gal)

- Tank Layer 1: Unscavenged UR waste 37,100 L (9,800 gal)

- Tank Layer 2: Unscavenged UR waste and ferrocyanide scavenging waste 37,100 L (9,800 gal)

- Tank Layer 3: Ferrocyanide scavenging waste 158,000 L (41,800 gal)

- Supernatant: 137,000 L (36,300 gal).

Figure 6-2B.   Alternative Waste Profile of Tank 241-C-112.

Not to Scale
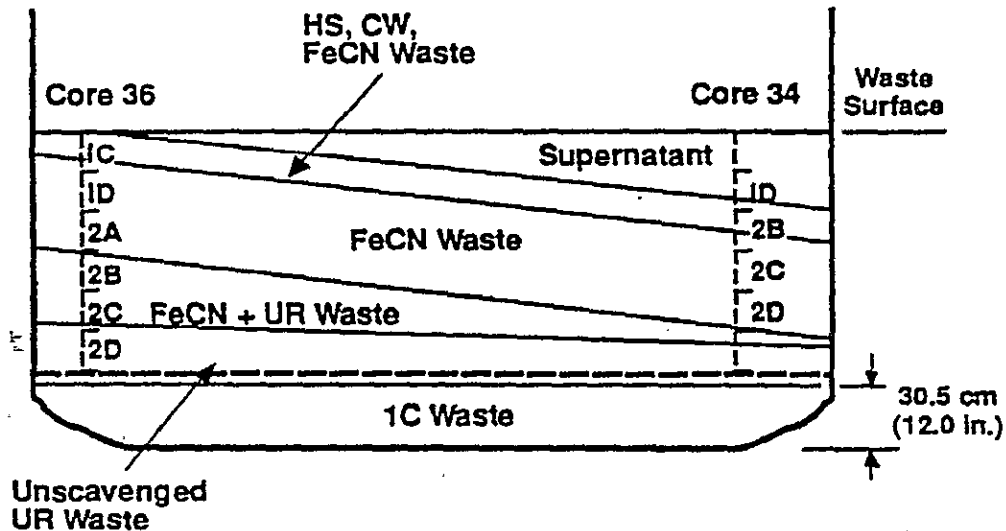
- Dished Bottom: First cycle BiPO$_4$ waste or unscavenged uranium recovery (UR) waste 57,700 L (15,000 gal)

- Tank Layer 1: Unscavenged UR waste 74,200 L (19,600 gal)

- Tank Layer 2: Unscavenged UR waste and ferrocyanide scavenging waste 37,100 L (9,800 gal)

- Tank Layer 3: Ferrocyanide scavenging waste 158,000 L (41,800 gal)

- Tank Layer 4: Miscellaneous waste residues (CW, HS) 49,200 L (13,000 gal)

- Supernatant: 37,500 L (9,900 gal).

Several assumptions must be made to calculate the tank contents before making comparisons to the Borsheim/Simpson model predictions for selected analytes. In addition, several assumptions of that model must be examined because they affect the original predictions regarding the waste in the tanks, such as (1) the assumption of 4.25 vol% solids formation (representative of the U Plant materials, not the In-Farm waste); (2) no additional settling or compaction; (3) negligible waste transfer (input/output) effects; and (4) transfers after the scavenging program did not meaningfully affect the condition of the waste. However, at the time they were obtained, these data and assumptions were the best available. As the ferrocyanide program evolved, more and better data became available.

The development of the model provided some preliminary understanding to the condition and distribution of the waste in the tank. Generally, the model gave values that were within $\pm$ 50% of the values calculated from the analytical results. Where agreement was not good, further investigation found reasonable sources for the difference. The range of values developed from the model was adequate for defining initial conditions (and bounding values); however, for analytes like ^{90}Sr and ferrocyanide itself, further process history contributed meaningfully to the present inventory in the tank as determined from laboratory analysis. Further clarification was provided by physical and chemical characterization of flowsheet materials, as well as aging and energetics studies. The model functioned well within the constraints placed on its operation, and it remains flexible enough to run further trials with new parameters, which have been done and are presented in Simpson et al. (1993).

As noted previously, the analytical nickel values are biased high, perhaps as much as 20 to 25%, by the use of a nickel crucible in the ICP fusion assay. Because the nickel tracer is biased high and it is unknown how much of the ferrocyanide solids were actually transported to the cribs, initial agreement is not good between the model estimates and the analytical results. In addition, ^{60}Co scavenging was done in several of the batches that were settled in tank 241-C-112, adding to the nickel inventory but not contributing to the ferrocyanide content. Approximately 32,900 g-mol of additional nickel was added to the tank in these process runs. Therefore, the nickel inventory determined from Borsheim/ Simpson should be adjusted upwards by that amount to account for the additional nickel; the model only accounted for nickel deposited with ferrocyanide. Agreement between the model values and the analysis-based estimates closes when the analytical bias is considered and with inventory adjustments from the cobalt-scavenging contribution. However, with all of the caveats associated with it, the nickel assay provides no more than a bounding condition for the ferrocyanide inventory as well as indicating that ferrocyanide was (or is) present.

Values for ^{137}Cs from Borsheim/Simpson adequately bound the inventories calculated from the analytical results. The calculated inventory can vary somewhat depending on the which core's density and concentration values are used in the computation. No overt biases were found in the analysis, and no other waste type that was disposed to tank 241-C-112 is believed to have a ^{137}Cs concentration high enough to confound inventory estimates (unlike ^{90}Sr). In addition, ^{137}Cs values provide a potential check on the waste retention of the tank. Cesium was widely dispersed in the slurry while it was settling, and its concentration profile

as a function of depth is relatively uniform in the subsegments believed to be ferrocyanide waste. The Borsheim/Simpson model has a significant amount of solids disposed to the cribs as a result of some of the model assumptions. If the ^{137}Cs values are used to estimate the waste retention in the tank, the amount of ferrocyanide waste retained in the tank is much higher than the estimated 40% retention from present model values.

A large degree of uncertainty is associated with the amount of ferrocyanide waste that may have been disposed to the cribs. The original model run has a large amount of solids being discharged, even though the available records indicate that the discharged effluent had only traces of suspended solids in it. The model basis of 4.25 vol% has been determined to be flawed for this waste type. Simulant studies indicate that an appropriate solids formation value for the In-Farm process is 1.0 to 1.5 vol% (Jeppson and Wong 1993). This additional information can be used to develop better model parameters and waste inventory estimates. A rerun of the model using these new volume parameters gives significantly better agreement.

The ferrocyanide inventory calculated from the total cyanide analysis ranges between 8,700 to 11,500 g-mol. The revised model value for the estimated remaining ferrocyanide of 60,000 to 66,000 g-mol is significantly higher than that determined from experimental results. The model estimated the total ferrocyanide used in processing waste through tank 241-C-112 is 78,500 g-mol. Given the improvement in agreement of the other analytes, this result, along with the energetic results, suggests a degradation or aging mechanism of some type.

6.3 CALCULATED BULK INVENTORIES OF SELECTED ANALYTES

Two different interpretations of the waste distribution in the tank were evaluated. The major difference in the two interpretations is the distribution of mass between supernate and wet solids. The second interpretation increases the mass of wet solids by approximately 30%, and the inventory relationship is linear for nearly all of the analytes (i.e., most of them increase by about 30%). The proportional contributions of each of the major analytes remain almost unchanged. The rationale for two interpretations of the tank conditions come from core recovery data and reviewing the in-tank photographs. The first interpretation, based on core recovery data, interpreted a significant amount of standing liquid remaining in the tank, and that the waste in the tank had a significant spatial bias (tilt) from the waste heel and input/output effects. The second interpretation was based on the in-tank photographs. The waste surface appears littered with debris and seems relatively solid; there is little standing liquid. Several liquid pools are observed and seem to be quite deep, some of them are located near the risers. The debris observed are potential obstructions to core sampling and may be the cause of poor sample recovery encountered in some of the samples. The waste surface appears wet, but solid, pocked with pools of liquid. However, this appearance may be deceiving--it could be a surface scum or a quicksand-like material.

Several safety issues are defined by certain bulk amounts or weight percent of a given analyte. Tables 6-2 through 6-5 present the calculated bulk amounts of some selected analytes and their weight percent contribution to the waste matrix in the first interpretation. Tables 6-6 through 6-9 present the same information for the 2nd interpretation. The gross waste inventory in the tank is estimated to range between 608,000 and 610,000 kg (434,000 kg wet solid, and 174,000 kg of drainable liquid in the first case; 563,000 kg wet solid and 47,800 kg of drainable liquid in the second case). Simpson et al. (1993) presents the data, assumptions, and calculations used to determine the following values. Values presented as concentration estimates and used in bulk inventory calculations are generally within the 95% confidence interval (CI) and RPD between sample and duplicate are usually However, given the heterogeneous nature of the tank waste, the degree of stratification in the tank, and the potential bias introduced from incomplete sample recovery, these values could be skewed by as much as 50% in some cases. In some cases, independent data can be used to verify or bound several of the presented values.

First Interpretation: refer to Figure 6-2A

Table 6-2. Energetics Related Analyte Values.

	TOC	Total Cyanide	NO_2^-	NO_3^-	H_2O
Bulk Inventory (Mg)	1.86	1.55	27.78	37.32	355.37[1]
Wt% (total)	0.31	0.25	4.57	6.14	58.42[1]
Bulk Inventory, wet solids (Mg)	1.56	1.35	18.19	24.76	233.22
Wt% (wet solids)	0.36	0.31	4.19	5.71	53.77

[1]Water content combines interstitial and free water (i.e., supernate).
TOC = Total organic carbon.

The bulk inventory of disodium nickel ferrocyanide in the tank is 8,700 g-mol, assuming the calculated inventory of total cyanide is present as that analyte. Molar ratios for ferrocyanide, nitrate, and nitrite in the wet solids (assuming this value for ferrocyanide) are 1: 45.4 : 45.9.

Table 6-3. Fission Product Inventory.

	^{137}Cs	^{90}Sr
Bulk Inventory (Ci) (wet solids)	217,200	183,500
Heat Generation (w)	1025	1230

The total heat load of the tank is 2,255. The volumetric heat generation for the waste in the tank based on the solids volume is 7.8 E-3 w/L solids at 1.50 g/mL.

Table 6-4. Plutonium/Americium Inventory.

	^{238}Pu	$^{239/240}$Pu	^{241}Am
Bulk Inventory (Ci) (wet solids)	5.9	67.3	330
Bulk Inventory (g)	0.34	1,100	100

Table 6-5. ICP Major Cation Inventory (Fusion Preparation).

	Al	Ca	Fe	Na	Ni	P	U
Bulk Inventory (Mg)	4.34	6.08	8.38	40.12	5.52	10.08	22.09
Wt% (Wet solids)	1.00	1.40	1.93	9.25	1.27	2.32	5.09

Second Interpretation: refer to Figure 6-2B

Table 6-6. Energetics Related Analyte Values.

	TOC	Total Cyanide	NO_2^-	NO_3^-	H_2O
Bulk Inventory (Mg)	2.08	1.85	26.23	35.60	327.01[1]
Wt% (total)	0.34	0.30	4.29	5.83	53.54
Bulk Inventory, wet solids (Mg)	2.00	1.80	23.61	32.17	293.70
Wt% (wet solids)	0.35	0.32	4.19	5.71	52.15

[1]Water content combines interstitial and free water (i.e., supernate).
TOC = Total organic carbon.

The bulk inventory of disodium nickel ferrocyanide in the tank is 11,500 g-mol, assuming the calculated inventory of total cyanide is present as that analyte. Molar ratios for ferrocyanide, nitrate, and nitrite in the wet solids (assuming this value for ferrocyanide) are 1: 44.6 : 45.1.

Table 6-7. Fission Product Inventory.

	^{137}Cs	^{90}Sr
Bulk Inventory (Ci) (wet solids)	286,000	162,000
Heat Generation (w)	1,350	1,030

The total heat load of the tank is 2,440. The volumetric heat generation for the waste in the tank based on the solids volume is 6.5 E-3 w/L solids at 1.50 g/mL.

Table 6-8. Plutonium/Americium Inventory.

	^{238}Pu	$^{239/240}Pu$	^{241}Am
Bulk Inventory (Ci) (wet solids)	7.7	87.1	430
Bulk Inventory (g)	0.44	1,400	125

Table 6-9. ICP Major Cation Inventory (Fusion Preparation).

	Al	Ca	Fe	Na	Ni	P	U
Bulk Inventory (Mg)	5.50	8.98	10.50	52.31	7.15	13.76	28.10
Wt% (Wet solids)	0.98	1.59	1.87	9.29	1.27	2.44	4.99

6.4 GAMMA SCANNING OF CORE SAMPLE SEGMENTS

Qualitative measurements of the gamma emitters in the waste were obtained. The information, although interesting and somewhat useful, was not as comprehensive as was hoped and could have been obtained at less cost through a lower technology method or simply by waiting for extrusion and analysis of the sample. The scans showed that radiocesium is the only significant gamma emitter in the waste. After evaluation of the data results from the gamma scans from tank 241-C-112, a decision was made to suspend any further gamma scanning on the remaining ferrocyanide tanks until hard salt cake samples (e.g., tank 241-BY-104) can be obtained. This decision was made on several bases; however the most significant factor was cost.

The information from the scan was also found to be an accurate indicator of core condition (i.e., full, partial, or empty) and it did confirm that there was some gross layering of the wastes, in addition to differences in activity between batches. This information is of some value because it allows modifications to the analysis plan prior to extrusion, if necessary, avoiding hot cell delays. Thus, field radiography is under consideration as a procedure to indicate percent recovery. The information regarding layering and variations batch content is also somewhat valuable; however, now that the hypothesis has been tentatively confirmed, the analytical horizons in the characterization plan for core sample analysis are now believed narrow enough to find ferrocyanide and ^{137}Cs concentrations.

Based on the information obtained from these initial scans, the cost associated with this analytical technique was not thought to be good value and drained resources from the program. Investigation of tank 241-BY-104 core samples (salt cake and sludge) is the next gamma scanning effort under consideration. Those core samples differ enough in composition and process history for the technique to be of potential value.

In addition to the cost-benefit argument, PNNL raised objections to continuing the gamma scanning effort because of the schedule delays caused by logistical problems in sample transport between the various facilities and hot cell scheduling. Given the limited sample transport, limited hot cell and personnel resources, and the uncertainty regarding the availability of these resources at any given time, the schedule constraints regarding data package delivery also became major considerations in discontinuing the effort.

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7.0 QUANTITATIVE/STATISTICAL INTERPRETATION OF THE DATA

This section contains the statistical analysis of core samples taken from tank 241-C-112. The analysis is divided into four sections. The first section contains mean concentration estimates in the form of CIs for each analyte of interest. Second, an examination of samples taken at two locations from homogenized subsegments was conducted to determine the ability of the 325 Analytical Chemistry Laboratory to homogenize solid core segments. The results from this examination indicate that the 325 Analytical Chemistry Laboratory is able to homogenize core subsegments satisfactorily. The third section is a comparison of the core composite concentration estimates with a simulated core composite computed from individual subsegment data. In 89% of the cases tested, the core composite could not be statistically distinguished from the simulated core composite. This indicates that the ability of the 325 Analytical Chemistry Laboratory to make core composites is satisfactory. The last topic in this section addresses the issue of spatial (between core) variability and analytical error. Variance component estimates and CIs are presented and discussed. In general, the spatial variability is of the same magnitude or larger than the analytical error.

Samples were obtained for subsegments 1D, 2B, 2C, and 2D from core 34; a single subsegment, 2D, from core 35; and subsegments 1C, 1D, 2A, 2B, 2C, and 2D from core 36. Drainable liquids were recovered from cores 34 and 35. Composite samples were made from the homogenized subsegment samples for each core. It should be noted that the core 35 composite and subsegment 2D are one and the same. Drainable liquid composites (DLC) were also made for cores 34 and 35. A sample and duplicate were taken from each core composite and prepared for analysis in the laboratory. Laboratory analyses were conducted on additional homogenized subsegment samples for the homogenization tests.

The laboratory results from tank 241-C-112 samples are tabulated in Jensen and Whitcher (1992). Sample preparations, assays, and analytes chosen for statistical evaluation are as follows: ICP acid digestion, ICP potassium hydroxide fusion dissolution, and ICP water leach analyses were conducted on all core composite samples and duplicates for aluminum, calcium, iron, sodium, nickel, lead, uranium, and phosphorous. Core 34 composite duplicate analysis results were not reported for ICP potassium hydroxide fusion dissolution and ICP water leach methods. The potassium hydroxide\nickel fusion dissolution was the only ICP method used to analyze subsegment results. Radiochemistry composite sample results were reported for uranium, ^{238}Pu, $^{239/240}$Pu, ^{137}Cs, and ^{90}Sr, as well as subsegment results for ^{137}Cs and ^{90}Sr. IC analysis results for chloride, nitrite, nitrate, phosphate, and sulfate were reported for core composite and subsegment samples. The ICP acid digestion analysis results were used to evaluate the homogenization test samples. Figures for the homogenization test data are also found in Jensen and Whitcher (1992). Drainable liquids were recovered from cores 34 and 35. They are included in part of the analysis and results for concentration estimates.

7.1 CONCENTRATION ESTIMATES

A task directed by WHC-EP-0210 to tank 241-C-112 was to estimate the constituent concentrations (Hill et al. 1991). This task was accomplished using means and appropriate CIs on these means. It should be emphasized that concentration estimates will be obtained rather than an inventory of the tank.

Each segment was subdivided into subsegments. The subsegments were each homogenized (laboratory core homogenization ability is covered in Section 4.0). Each core composite was formed by combining samples from the homogenized subsegments. The core composite sample was then also homogenized. Two aliquots were drawn from each core composite and prepared for chemical analysis. Concentration estimates were computed based on the results of these chemical analyses. Jensen and Whitcher (1992) contains the core composite data used to obtain concentration estimates and intervals. The "NA" symbol indicates that the data is not available.

The concentration estimates are given in the form of 95% CIs on the mean concentration of each analyte in the tank. It is assumed that each sample and duplicate are analyzed independent of one another to yield adequate estimates of analytical error. Due to the hierarchical structure of the data, the analytical error alone is not the appropriate error term to use in the CI calculations. A linear combination of the analytical error and spatial variability is the appropriate measure of spread for the CIs. The derivation of the formulas used to calculate these CIs is given in a Westinghouse Hanford Company internal memo (Jensen and Whitcher 1993).

Tables 7-1, 7-2, 7-3, 7-4, and 7-5 contain the following summary statistics for all three ICP analyses (acid digestion, water leach, potassium hydroxide/nickel fusion dissolution), radiochemistry, and IC anion analyses, respectively.

For some analytes the lower confidence limit L was negative. Because concentrations are strictly greater than or equal to zero, any negative L values were set to zero.

The DLCs were added to the solid composite data to increase the information and, therefore, give more accurate CIs. Some of the DLC results from the ICP acid digestion analysis are quite different from the solid core composite results. For this reason two sets of summary statistics on all core composite results are presented, one including and another set excluding the DLC results. Plutonium, uranium, nickel, calcium, and iron DLC results are quite different than the solid core composite data. These differences will inflate the analytical error, which is a component in $\hat{\sigma}_{\bar{y}}^2$ for the CI. The DLCs do not create a problem for ICP acid digestion sodium analysis or any of the ICP water leach analyses.

A close examination of the ICP water leach sodium results revealed a large residual associated with one of the core 36 composite result of 130,962 μg/g. Its duplicate result was 85,753 μg/g, which is within the range of the other core composite results. Additional summary statistics are given excluding this outlying result.

Table 7-1. Acid Digestion Inventory Statistics (μg/g).

Excluding Drainable Liquid Composites						
Analyte	$\bar{y}$	BMS	$\hat{\sigma}_{\bar{y}}^2$	df	L	U
Al	24,464	7.65 E+08	1.28 E+08	2	0	73,056
Ca	16,743	5.12 E+07	8.53 E+06	2	4,179	29,307
Fe	21,853	2.63 E+07	4.40 E+06	2	12,836	30,871
Na	89,567	5.57 E+08	9.30 E+07	2	48,115	131,018
Ni	13,113	3.20 E+07	5.33 E+06	2	3,179	23,047
Pb	2,553	4.59 E+06	7.64 E+05	2	0	6,315
U	58,389	3.22 E+09	5.37 E+08	2	0	158,129
P	22,915	9.4 E+07	1.57 E+07	2	5,864	39,967
Including Drainable Liquid Composites						
Analyte	$\bar{y}$	BMS	$\hat{\sigma}_{\bar{y}}^2$	df	L	U
Al	24,464	7.65 E+08	1.28 E+08	2	0	73,056
Ca	10,189	8.35 E+07	8.34 E+06	2	0	22,621
Fe	13,473	5.28 E+07	5.27 E+06	2	3591	23,354
Na	84,235	7.71 E+08	7.71 E+07	2	46447	122,024
Ni	8,084	1.95 E+06	1.95 E+06	2	2078	14,089
Pb	2,553	4.59 E+06	7.64 E+05	2	0	6,315
U	35,446	4.05 E+09	4.05 E+08	2	0	122,048
P	15,283	3.02 E+08	3.02 E+07	2	0	38,949

$\bar{y}$: Arithmetic mean of the concentration data
BMS: "Between mean squares" from the one-way analysis of variance (ANOVA)
$\hat{\sigma}_{\bar{y}}^2$: Estimated variance of $\bar{y}$
df: Degrees of freedom associated with BMS
L: Lower limit to the 95% CI on the mean
U: Upper limit to the 95% CI on the mean.

Table 7-2. Potassium Hydroxide Fusion Dissolution Inventory Statistics (μg/g).

Analyte	$\bar{y}$	BMS	$\hat{\sigma}^2_{\bar{y}}$	df	L	U
Al	26,540	7.53 E+08	1.51 E+08	2	0	79,337
Ca	19,950	6.56 E+07	1.31 E+07	2	4,367	35,533
Fe	27,915	3.62 E+07	7.23 E+06	2	16,344	39,486
Na	103,893	8.43 E+08	1.69 E+08	2	48,026	159,760
Ni	NA	NA	NA	NA	NA	NA
Pb	2,937	6.51 E+0	1.30 E+06	2	0	7,846
U	80,730	2.87 E+09	5.74 E+08	2	0	183,779
P	27,971	1.29 E+08	2.58 E+07	2	6,135	49,808

$\bar{y}$: Arithmetic mean of the concentration data
BMS: "Between mean squares" from the one-way analysis of variance (ANOVA)
$\hat{\sigma}^2_{\bar{y}}$: Estimated variance of $\bar{y}$
df: Degrees of freedom associated with BMS
L: Lower limit to the 95% CI on the mean
U: Upper limit to the 95% CI on the mean.

Table 7-3. Water Leach Inventory Statistics (μg/g).

Excluding Drainable Liquid Composites					
Analyte	$\bar{y}$	BMS	$\hat{\sigma}_{\bar{y}}^2$	df	L	U
Al	521	1.43 E+05	2.87 E+04	2	0	1,250
Ca	338	7.04 E+03	1.41 E+03	2	176	499
Fe	1,168	3.66 E+05	7.31 E+04	2	5	2,332
Na	88,541	1.31 E+09	2.62 E+08	2	18,838	158,244
Na #						
Ni	679	1.38 E+05	2.76 E+04	2	0	1,395
Pb	NA	NA	NA	NA	NA	NA
U	2,166	9.87 E+06	1.97 E+06	2	0	8,211
P	10,108	6.72 E+07	1.34 E+07	2	0	25,886
Including Drainable Liquid Composites					
Analyte	$\bar{y}$	BMS	$\hat{\sigma}_{\bar{y}}^2$	df	L	U
Al	521	1.43 E+05	2.87 E+04	2	0	1,250
Ca	253	2.57 E+04	3.67 E+03	2	0	513
Fe	1,092	4.20 E+05	4.68 E+04	2	161	2,022
Na	84,078	1.34 E+09	1.49 E+08	2	31477	136,679
Na #						
Ni	662	1.65 E+05	1.84 E+04	2	79	1,245
Pb	NA	NA	NA	NA	NA	NA
U	1,673	1.11 E+07	1.23 E+06	2	0	6,442
P	7,350	1.07 E+08	1.18 E+07	2	0	22,153

\# Excluding outlier data result
$\bar{y}$: Arithmetic mean of the concentration data
BMS: "Between mean squares" from the one-way analysis of variance (ANOVA)
$\hat{\sigma}_{\bar{y}}^2$: Estimated variance of $\bar{y}$
df: Degrees of freedom associated with BMS
L: Lower limit to the 95% CI on the mean
U: Upper limit to the 95% CI on the mean.

Table 7-4. Radiochemistry Statistics (μCi/g).

Element	$\bar{y}$	BMS	$\hat{\sigma}_{\bar{y}}^2$	df	L	U
^{137}Cs	747	8.38 E+03	1.68 E+03	2	571	924
^{90}Sr	2,189	4.74 E+06	9.48 E+05	2	0	6,380
U_{FL}(μg/g)	68,600	3.89 E+09	1.30 E+09	1	0.0	525,911
$^{239/240}$Pu	0.06	0.013323	0.002665	2	0.0	0.28
^{238}Pu	0.06	0.025023	0.005005	2	0.00	0.37

FL = Laser fluorimetry.

Table 7-5. Ion Chromatography Anion Statistics (μg/g).

Compound	$\bar{y}$	BMS	$\hat{\sigma}_{\bar{y}}^2$	df	L	U
Chloride	980	109,000	21,800	2	345	1,615
Nitrite	47,200	2.99 E+08	5.98 E+07	2	13,930	80,470
Nitrate	62,000	5.95 E+08	1.19 E+08	2	15,080	108,921
Phosphate	30,760	5.86 E+08	1.17 E+08	2	0	77,337
Sulfate	12,040	2.09 E+07	4.19 E+06	2	3,237	20,843

$\bar{y}$: Arithmetic mean of the concentration data
BMS: "Between mean squares" from the one-way analysis of variance (ANOVA)
$\hat{\sigma}_{\bar{y}}^2$: Estimated variance of $\bar{y}$
df: Degrees of freedom associated with BMS
L: Lower limit to the 95% CI on the mean
U: upper limit to the 95% CI on the mean.

The radiochemistry U confidence limits should be viewed with caution. This interval is based on three sample results from cores 34 and 36. The CI for $\bar{y}$ uses a t-statistic with one degree of freedom, which inflates the CI width due to the uncertainty of the U results.

7.2 HOMOGENIZATION TESTS

A second task, directed by WHC-EP-0210 Rev 1 (Winters et al. 1990a) to core samples from tank 241-C-112, was to evaluate the ability of the 325 Analytical Chemistry Laboratory to homogenize core subsegments. Each subsegment (from cores 34, 35, and 36) was homogenized and arbitrarily divided into two parts (top and bottom). One subsample was obtained from each part. Two aliquots were taken from each subsample and prepared for chemical analysis. The homogenization test data described above are tabulated in Jensen and Whitcher (1992). ICP acid digestion analyses were conducted on the aliquots for the following analytes: aluminum, iron, sodium, nickel, lead, uranium, and phosphorus.

Due to the structure of the data, a hierarchical model can be fit to the data. This model is used to separate different components of variability in the data. The total variability in the data is composed of three separate components: one due to variability between the segments; one due to the variability between samples from the homogenized material (σ_h^2) within the segments (this component of variability is referred to as being due to homogenization); and one due to analytical error. The analytical error is the variance component that measures the variability between duplicate measurements from aliquots prepared from the same sample. Jensen and Whitcher (1992) contains a description of the hierarchical model. From the analysis of variance (ANOVA) results of the hierarchical model, a test can be conducted to determine if σ_h^2 is significantly greater than zero. If σ_h^2 is significantly greater than zero, then it can be concluded that the laboratory does not have the ability to homogenize core segments. If it cannot be concluded that σ_h^2 is significantly greater than zero, then the laboratory does have the ability to homogenize core segments. The F-statistic p-value from the ANOVA was used to test if σ_h^2 is greater than zero. If the p-value is smaller than 0.05, it is concluded that the $\sigma_h^2 > 0$.

The results of the homogenization test are given in Table 7-6. For all analytes tested (aluminum, calcium, iron, sodium, nickel, phosphorus, lead, uranium), the p-value from the F-test was greater than 0.05. It cannot be concluded that the variability between the sampled locations of the mixed subsegments is significantly different from zero. The overall conclusion is that the ability of the 325 Analytical Chemistry Laboratory to homogenize core samples is satisfactory for this material. It is interesting to note that segment-to-segment variability always accounts for more than 77% of the total variability, and analytical error claims most of the rest. The percent of the variability due to the homogenization is usually less than 1%.

Table 7-6. Homogenization Test Results.

Analyte	p-value	% Variability due to:		
		segment	Homogenization	Analytical
Al	0.65	94.82	0.00	5.18
Ca	0.10	99.45	0.34	0.21
Fe	0.58	77.10	0.00	22.90
Na	0.72	92.95	0.00	7.05
Ni	0.42	99.17	0.04	0.79
P	0.26	96.55	0.74	2.71
Pb	0.22	99.73	0.87	2.40
U	0.08	99.55	0.21	0.24

An observation with a large residual was found for core 34, subsegment 2D iron data. The removal of this observation did not change the results of the homogenization tests.

7.3 COMPARISON BETWEEN THE SIMULATED CORE COMPOSITE AND THE CORE COMPOSITE

Another task directed by WHC-EP-0210 (Winters et al. 1990a,b) was to evaluate the ability of the 325 Analytical Chemistry Laboratory to prepare core composite samples from the individual subsegment samples. A core composite sample was formed by combining equally weighted individual samples from each homogenized subsegment in the core. The core composite was then homogenized and a sample and duplicate were taken. A simulated core composite was constructed to compare with the core composite samples. The simulated core composite is the mean of the results from the chemical analysis of the individual aliquots from each segment.

For the chemical analysis methods reporting subsegment results (ICP potassium hydroxide fusion dissolution, radiochemistry, IC), core 34 did not have core composite duplicate results reported. For this reason, the method used to compare the core composite with the simulated core composite for core 36 is slightly different than the method used for core 34.

7.3.1 Core 36 Statistical Methods

The comparison for core 36 is made by computing a CI on the difference between the two means. If zero is in the CI, the simulated core composite mean cannot be statistically distinguished from the core composite sample mean. If zero is not in the CI, the two means are significantly different.

The CI on the difference between the simulated core composite and core composite mean is (L, U), where the lower (L) and (U) values are

$$L = (\bar{y}_w - \bar{y}_c) - t\sqrt{\hat{\sigma}^2(\bar{y}_w - \bar{y}_c)} \ , \quad U = (\bar{y}_w - \bar{y}_c) + t\sqrt{\hat{\sigma}^2(\bar{y}_w - \bar{y}_c)}$$

and where

$\bar{y}_w$:	=	Subsegment data mean (simulated core composite)
$\bar{y}_c$:	=	Core composite sample mean
t:	=	97.5 percentile point from Student's t distribution
$\hat{\sigma}^2(\bar{y}_w - \bar{y}_c)$:	=	Estimated variance of the difference.

Jensen and Whitcher (1992) contains a discussion on how $\hat{\sigma}^2(\bar{y}_w - \bar{y}_c)$ is calculated. The calculation of $\hat{\sigma}^2(\bar{y}_w - \bar{y}_c)$ uses the data from all three cores because of the limited data available. The "t" has degrees of freedom (df) associated with it, which was calculated using Satterwaite's approximation (Snedecor and Cochran 1980).

7.3.2 Core 34 Statistical Methods

A CI for $\bar{y}_w$ is calculated to make the comparison between the core composite and the simulated core composite. If the composite sample result is contained within the CI, then $\bar{y}_w$ cannot be statistically distinguished from the core composite sample result. If the core composite sample results are not contained within the CI, then $\bar{y}_w$ and the core composite sample are statistically different.

The CI on $\bar{y}_w$ is (L, U) where the lower (L) and (U) values are

$$L = \bar{y}_w - t\sqrt{\hat{\sigma}^2(\bar{y}_w)}, \quad U = \bar{y}_w + t\sqrt{\hat{\sigma}^2(\bar{y}_w)}$$

and where $\hat{\sigma}^2(\bar{y}_w)$ is the estimated variance of $\bar{y}_w$. Jensen and Whitcher (1992) contains a discussion of how $\hat{\sigma}^2(\bar{y}_w)$ is calculated.

7.3.3 Results from Cores 34 and 36

Tables 7-7 and 7-8 present the results of the simulated core composites mean versus the core composite mean for cores 34 and 36 respectively. Jensen and Whitcher (1992) contains figures of the segment and core composite data for all three cores, which may aid in the understanding the results from the tables.

Only three analytes (sodium, chloride, sulfate) had composite samples that were outside of the $\bar{y}_w$ CI for core 34. In all three cases, the composite sample exceeds the upper limit of the CI. For core 36, all the CIs of the difference $\bar{y}_w - \bar{y}_c$ contained zero, indicating that there was no significant difference between $\bar{y}_w$ and $\bar{y}_c$. Overall, the simulated core composite could be statistically distinguished from the mean in 89% of the cases tested. On this basis, it can be concluded that the 325 Analytical Chemistry Laboratory is able to make core composites satisfactorily.

7.4 SPATIAL VARIABILITY AND ANALYTICAL ERROR ESTIMATION

Using the hierarchical structure of the data (cores selected randomly and two aliquots selected per core composite), the spatial variability and the analytical error can be separated from each other. The spatial variability is the variability from core composite to core composite. The analytical error, as it is called here, is not only true analytical error. Confounded with it is segment homogenization and sampling variability. Variability from all three of these sources will be referred to as analytical error.

Table 7-7. Segment vs. Composite Statistics for Core 34.

Analyte	$\bar{y}_w$	$\hat{\sigma}^2(\bar{y}_w)$	Composite	df*	95% CI for $\bar{y}_w$	
					L	U
Al	23,549	5.24 E+07	29,798	2	0	54,708
Ca	25,894	1.89 E+07	28,984	8	15,864	35,923
Fe	13,295	1.95 E+07	22,848	8	3,107	23,483
Na #	87,128	1.35 E+07	114,891	9	78,806	95,451
Pb	3,220	1.10 E+06	3,326	3	0	6,559
U	20,098	6.98 E+08	14,369	5	0	88,039
P	17,318	6.50 E+06	21,956	6	11,076	23,561
Chloride #	963	4,464	1,300	6	799	1,126
Nitrite	51,875	1.88 E+07	62,000	5	40,740	63,010
Nitrate	68,125	2.80 E+07	80,000	5	54,528	81,722
Phosphate	13,200	3.37 E+07	19,100	7	0	26,922
Sulfate #	12,587	8.57 E+05	15,600	6	10,322	14,852
Cs-137	539	17,666	750	9	238	840
Sr-90	2,456	5.37 E+05	3,510	4	422	4,490

#: Significant difference between segment data and composite.

units: μg/g: Al, Ca, Fe, Na, Pb, U, P, chloride, nitrite, nitrate, phosphate, sulfate
μCi/g: Cs-137, Sr-90.

Table 7-8. Segment vs. Composite Statistics for Core 36.

Analyte	$\bar{y}_w$	$\bar{y}_c$	$\hat{\sigma}^2(\bar{y}_w-\bar{y}_c)$	df*	95% CI for $\bar{y}_w-\bar{y}_c$	
					L	U
Al	4,948	6,410	4.29 E+08	3	-67,364	64,439
Ca	15,484	20,391	5.17 E+07	5	-23,393	13,580
Fe	18,355	26,012	3.76 E+07	7	-22,158	6,844
Na	92,724	120,730	4.35 E+08	2	-117,756	61,744
Pb	1,788	1,049	4.36 E+06	3	-5,902	7,379
U	96,380	104,910	2.13 E+09	4	-136,748	119,688
P	24,733	36,761	7.10 E+07	2	-48,288	24,232
Chloride	783	1,050	58,964	2	-1,312	778
Nitrite	40,750	52,500	1.68 E+08	3	-53,025	29,525
Nitrate	55,667	71,500	3.25 E+08	2	-93,463	61,796
Phosphate	28,858	49,500	3.27 E+08	2	-98,414	57,130
Sulfate	10,600	13,700	1.13 E+07	2	-17,578	11,377
Cs-137	559	793	21,857	11	-559	92
Sr-90	384	508	2.90 E+06	3	-7,462	7,214

#: Significant difference between segment data and composite.
units: μg/g: Al, Ca, Fe, Na, Pb, U, P, chloride, nitrite, nitrate, phosphate, sulfate
μCi/g: Cs-137, Sr-90.

The size of the analytical error compared to the spatial variability has an impact on the analyte concentration CIs from Section 3.0. The variance used in the CI calculations is a linear function of estimates of spatial and analytical variability. Because the size of the analytical error can potentially be controlled, this comparison is an interesting factor.

Estimates of the spatial variability ($\hat{\sigma}_s^2$) and analytical error ($\hat{\sigma}_a^2$) were obtained for each analyte by solving a system of equations using ANOVA results and expected mean squares (Snedecor and Cochran 1980). Snedecor and Cochran (1980) also explain how CIs for σ_s^2 and σ_a^2 were obtained. An approximate CI for σ_s^2 was obtained using the ANOVA results and four F-statistic values. Exact CIs were obtained for σ_a^2 using $\hat{\sigma}_a^2$ and a chi-square statistic. Tables 7-9, 7-10, 7-11, 7-12, and 7-13 contain estimates and 95% CIs for σ_s^2 and σ_a^2 for all 14 analytes and chemical analysis methods. Using the variance component estimation methods described above, it is possible to obtain negative estimates. When negative variance components estimates were obtained they were set equal to zero. Lower limits (L) on the 95% CIs were also set equal to zero if the were negative.

Table 7-9. Acid Digestion Spatial and Analytical Error Estimates.

Excluding Drainable Liquid Composites						
Analyte	$\hat{\sigma}_s^2$	$\hat{\sigma}_a^2$	95% CI for $\hat{\sigma}_s^2$ *		95% CI for $\hat{\sigma}_a^2$	
			L	U	L	U
Al	3.67 E+08	3.10 E+07	9.69 E+07	7.44 E+09	9.93 E+06	4.30 E+08
Ca	2.49 E+07	1.44 E+06	7.11 E+06	4.98 E+08	4.62 E+05	1.20 E+07
Fe	0	2.76 E+07	0	2.43 E+08	8.85 E+06	3.83 E+08
Na	2.66 E+08	2.39 E+07	6.91 E+07	5.42 E+09	7.66 E+06	3.32 E+08
Ni	1.55 E+07	1.06 E+06	4.28 E+06	3.11 E+08	3.41 E+05	1.48 E+07
Pb	2.28 E+06	24,113	7.43 E+05	4.47 E+07	7,737	3.35 E+05
U	1.56 E+09	1.10 E+08	4.28 E+08	3.14 E+10	3.54 E+07	1.53 E+09
P	4.52 E+07	3.83 E+06	1.19 E+07	9.17 E+08	1.23 E+06	5.32 E+07
Including Drainable Liquid Composites						
Analyte	$\hat{\sigma}_s^2$	$\hat{\sigma}_a^2$	95% CI for $\hat{\sigma}_s^2$ *		95% CI for $\hat{\sigma}_a^2$	
			L	U	L	U
Al	3.67 E+08	3.10 E+07	9.69 E+07	7.44 E+09	9.93 E+06	4.30 E+08
Ca	12,440	8.34 E+07	0	4.82 E+08	3.65 E+07	3.46 E+08
Fe	0	1.55 E+08	0	2.73 E+08	6.77 E+07	6.41 E+08
Na	2.31 E+08	3.28 E+07	4.63 E+07	4.69 E+09	1.43 E+07	1.36 E+08
Ni	0	5.82 E+07	0	1.00 E+08	2.55 E+07	2.41 E+08
Pb	2.28 E+06	24,112	7.43 E+05	4.47 E+07	7,736.657	3.35 E+05
U	9.72 E+08	9.39 E+08	0	2.44 E+10	4.11 E+08	3.89 E+09
P	7.36 E+07	6.70 E+07	0	1.82 E+09	2.93 E+07	2.77 E+08

*Approximate 95% CI.

Table 7-10. Water Leach Spatial and Analytical Error Estimates.

Excluding Drainable Liquid Composites						
Analyte	$\hat{\sigma}_s^2$	$\hat{\sigma}_a^2$	95% CI for $\hat{\sigma}_s^2$ *		95% CI for $\hat{\sigma}_a^2$	
			L	U	L	U
Al	70,595	30,458	0	1.73 E+06	8,254	120,388
Ca	0	37,798	0	61,606	10,243	149,400
Fe	2.26 E+05	3,433	7.28 E+04	4.45 E+06	930	13,569
Na	5.01 E+08	5.11 E+08	0	1.57 E+10	1.38 E+08	2.02 E+09
Ni	85,597	1,287	2.76 E+04	1.68 E+06	349	5,087
Pb	NA	NA	NA	NA	NA	NA
U	4.55 E+06	2.60 E+06	0	1.19 E+08	7.03 E+05	1.03 E+07
P	3.50 E+07	1.13 E+07	2.27 E+06	8.12 E+08	3.06 E+06	4.46 E+07
Including Drainable Liquid Composites						
Analyte	$\hat{\sigma}_s^2$	$\hat{\sigma}_a^2$	95% CI for $\hat{\sigma}_s^2$ *		95% CI for $\hat{\sigma}_a^2$	
			L	U	L	U
Al	70,595	30,458	0	1.73 E+06	8,254	1.20 E+05
Ca	0	41,181	0	2.29 E+05	14,840	3.40 E+05
Fe	1.38 E+05	22,384	25,347	2.83 E+06	9,327	1.08 E+05
Na	3.86 E+08	2.31 E+08	0	9.00 E+09	9.63 E+07	1.12e+09
Ni	54,992	6,588	12,290	1.11 E+06	2,745	3.19 E+04
Pb	NA	NA	NA	NA	NA	NA
U	3.50 E+06	9.30 E+05	3.11 E+05	7.43 E+07	3.87 E+05	4.50 E+06
P	3.52 E+07	4.94 E+06	7.18 E+06	7.17 E+08	2.06 E+06	2.39 E+07

*Approximate 95% CI.

Table 7-11. Potassium Hydroxide/Nickel Fusion Dissolution Spatial and Analytical Error Estimates.

Analyte	$\hat{\sigma}_s^2$	$\hat{\sigma}_a^2$	95% CI for $\hat{\sigma}_s^2$ *		95% CI for $\hat{\sigma}_a^2$	
			L	U	L	U
Al	4.70 E+08	5.67 E+05	1.57 E+08	9.17 E+09	1.54 E+05	2.24 E+07
Ca	4.09 E+07	1.56 E+05	1.37 E+07	7.99 E+08	42,269	6.16 E+06
Fe	1.65 E+07	9.81 E+06	7.56 E+06	4.34 E+08	2.66 E+06	3.88 E+08
Na	5.24 E+08	3.96 E+06	1.76 E+08	1.03 E+10	1.07 E+06	1.57 E+08
Ni	NA	NA	NA	NA	NA	NA
Pb	3.98 E+06	1.47 E+05	1.36 E+06	7.92 E+07	3.99 E+04	5.82 E+06
U	1.78 E+09	1.64 E+07	5.99 E+08	3.49 E+10	4.44 E+06	6.48 E+08
P	8.04 E+07	1.37 E+05	2.69 E+07	1.57 E+09	37,210	5.43 E+06

*Approximate 95% (CI).

Table 7-12. Radiochemistry Spatial and Analytical Error Estimates.

Analyte	$\hat{\sigma}_s^2$	$\hat{\sigma}_a^2$	95% CI for $\hat{\sigma}_s^2$ *		95% C. I. for $\hat{\sigma}_a^2$	
			L	U	L	U
^{137}Cs	4,949	463	1,752	1.02 E+05	125	18,281
^{90}Sr	3.56 E+06	1,360	9.91 E+05	5.78 E+07	369	53,767
U	0	2.76 E+07			7.47 E+06	1.09 E+09
$^{239/240}$Pu	8.29 E-03	5.4 E-05	2.78 E-03	1.62 E-01	1.5 E-05	2.13 E-03
^{238}Pu	1.56 E-02	3.2 E-05	5.23 E-03	3.05 E-01	0.9 E-05	1.27 E-03

*Approximate 95% (CI).

Table 7-13. Ion Chromatography Anion Spatial and Analytical Error Estimates.

Analyte	$\hat{\sigma}_s^2$	$\hat{\sigma}_a^2$	95% CI for $\hat{\sigma}_s^2$ *		95% CI for $\hat{\sigma}_a^2$	
			L	U	L	U
Chloride	52,500	25,000	22,784	1.31 E+06	6,775	9.88 E+05
Nitrite	1.04 E+08	1.33 E+08	6.25 E+07	3.56 E+09	3.59 E+07	5.24 E+09
Nitrate	2.21 E+08	2.41 E+08	1.24 E+08	7.09 E+09	6.52 E+07	9.51 E+09
Phosphate	2.97 E+08	1.10 E+08	1.22 E+08	7.07 E+09	2.99 E+07	4.36 E+09
Sulfate	9.17 E+06	6.25 E+06	4.37 E+06	2.51 E+08	1.69 E+06	2.47 E+08

*Approximate 95% (CI).

Including or excluding the DLC results greatly affected $\hat{\sigma}_s^2$ and $\hat{\sigma}_a^2$ for the acid digestion and water leach ICP analyses (refer to Tables 7-9 and 7-10). When the DLCs were excluded, $\hat{\sigma}_s^2$ was generally larger than $\hat{\sigma}_a^2$. When the DLCs were included, $\hat{\sigma}_a^2$ was generally of the same magnitude or larger than $\hat{\sigma}_s^2$. For the radiochemistry analysis, an estimate and approximate CI for σ_s^2 were not obtained due to the small number of data reported.

In general, the upper limits (U) on the approximate 95% CIs for σ_s^2 were larger than the 95% CI upper limits for σ_a^2 (88% of the cases). From this result it can be generally concluded that the spatial variability (σ_s^2) is of the same magnitude or larger than the analytical error σ_a^2.

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8.0 CONCLUSIONS AND RECOMMENDATIONS

8.1 ANALYTICAL RESULTS

Analyses of the waste show a very small number of analytes comprising a large portion of the waste. Water is the single largest analyte, making up over 50% of the total mass. Seven elements (aluminum, calcium, iron, sodium, nickel, phosphorus, and uranium) constitute approximately 22.3% of the solids mass. They also represent over 95% of the total cations. Two anions (NO_2^- and NO_3^-) constitute approximately 10% of the mass. The fraction of the total anions that nitrate and nitrite represent cannot be adequately determined because the analytical method measured only soluble anions. The total cyanide content was measured and found to be less than 1% (dry basis) in each core and for the tank as a whole.

The only significant gamma emitter found in the waste was ^{137}Cs. Although ^{60}Co was also precipitated during the scavenging process, it has decayed below any level of concern and does not contribute to the heat load of the tank. No meaningful regional concentrations (hot spots) of radioisotopes or fuel were detected along the vertical axis in either core. The ^{137}Cs concentration was relatively constant between core 34 and core 36 and the regions with high nickel concentrations correlated with the ^{137}Cs peaks in the gamma scans. These observations are consistent with the historical information regarding the ferrocyanide-scavenging process, the gamma activity pattern obtained from the scans, and the ICP element distribution through the subsegments. The other major source of radiological activity was ^{90}Sr, which had a very skewed distribution between core 34 and core 36, probably arising from the later waste transfers into tank 241-C-112 after the ferrocyanide-scavenging runs were completed. This is not a "hot spot" type phenomena, but an artifact of sampling near the waste inlet and outlet. There is no evidence of a self-concentration mechanism (Dickinson et al. 1993). Heat-load calculations are further evidence supporting the contention that the relatively high ^{90}Sr concentrations are believed to be a local phenomenon. The bulk waste temperature in the tank, obtained from two thermocouple trees, is 29 °C (85 °F). Comparisons of heat-load calculations, using the temperature profiles from the thermocouple trees and the higher ^{90}Sr concentrations, do not agree with the tank waste temperature measurements and other observations of the tank waste. The radiological activity of tank 241-C-112 waste material was relatively low (ranging from 0.15 to 2 R/hr, measured through the drill string). No significant radiological activity was found in the drainable liquid in the tank or in the water or acid digestion of the samples. This suggests that ^{90}Sr and ^{137}Cs are quite insoluble.

Indications from core 34 data show that material has physical and chemical properties corresponding to those expected for ferrocyanide waste. Water content, nickel concentration, and density values are consistent with the simulant values. Indications from core 36 data show that material appears to have ferrocyanide waste overlying a bismuth phosphate 1C and/or UR waste heel. However, no firm evidence of $BiPO_4$ 1C waste was ever found from the assays. The concentration of nickel, calcium, ^{137}Cs, and uranium as a function of depth appears to confirm this observation. In addition, this behavior is consistent with historical

information. However, in both cases DSC results from the suspected ferrocyanide waste in tank 241-C-112 indicate that the material is considerably less energetic than the corresponding waste simulant. Core 35 is indeterminate. It may be a combination of ferrocyanide waste and unscavenged UR waste. Results of aging studies now underway on flowsheet simulants may demonstrate that radiolytic, hydrolytic, and thermal processes in the tanks, over the last 35 years, have combined to dissolve, dilute, and destroy the reactive ferrocyanide compounds. The data from tank 241-C-112 strongly indicate that the waste lacks the fuel concentration to sustain any propagating exothermic behavior, and a heat source intense enough to trigger a reaction is absent.

8.2 ENERGETICS ANALYSIS

Thermodynamic and kinetics (propagation) studies of simulants are bounding the reactivity of the ferrocyanide/oxidizer reactions. Results indicate that U Plant ferrocyanide wastes cannot create a propagating hazard; T Plant waste is expected to behave similarly. Ferrocyanide simulants made by the In-Farm flowsheet are more reactive. The waste in tank 241-C-112 and the other C Farm tanks, representing 20 to 25% of the ferrocyanide inventory added to the tanks, was made by a similar process and was a potential cause for concern. However, if the In-Farm simulants contained at least 15 wt% water, that moisture content precluded an uncontrolled, propagating reaction (Fauske 1992).

Three core samples were obtained from tank 241-C-112, a tank considered to contain one of the highest concentrations of ferrocyanide. All three cores were broken down into smaller subsegments and examined for reactivity using DSC; none of the samples exhibited any propagating behavior. The samples had a moisture content ranging from 34 to 64 wt% water and a ΔH ranging from -0.02 kJ/g of dry material to -0.04 kJ/g dry material (4.4 to 8.6 cal). The onset temperatures for the exotherms ranged between 275 °C and 290 °C-- close to that predicted by the simulants. The simulant may have represented the waste as it was initially precipitated in the tank; several of the physical and chemical properties of the simulants are quite close to those of the tank waste. However, further chemical analysis indicates that the waste material has a total cyanide content much lower than expected from the simulant formulations and, correspondingly, tank 241-C-112 waste material is not as energetic as the analogous waste simulants. Tank 241-C-112 sample material is nearly 30 times less chemically reactive than the comparable In-Farm simulant material (tank 241-C-112 waste: -0.04 kJ/g; In-Farm simulant: -1.20 kJ/g). The causes of this behavior are hypothesized to be long-term exposure to radiation fields and high pH CW. Both of these conditions are believed to degrade the ferrocyanide complexes. Further testing of simulant materials, as well as results from other ferrocyanide tanks, will aid in confirming this hypothesis.

Calculations of the bulk waste inventory and inventories for several analytes of interest to the various safety issues (ferrocyanide, NO_2^-, NO_3^-, ^{137}Cs, ^{90}Sr, plutonium, and water) were made. The calculated bulk inventory of ferrocyanide (between 8,700 and 11,500 g-mol) was far in excess of the 1,000 g-mol Ferrocyanide Watch List criteria, but the energetics results

indicate that particular criteria do not account for the dispersion of the ferrocyanide in the waste (i.e., the concentration may be too low to support a self-sustaining reaction). None of the other calculated bulk inventory values exceeded any level of concern (refer to Table 8-1).

Experimental and analytical evidence from tank 241-C-112 suggests the risk from ferrocyanide compounds in Hanford Site high-level waste tanks is acceptable and that a propagating exothermic ferrocyanide reaction is incredible.

Table 8-1. Comparison of Tank 241-C-112 Analyte Values to
Safety Issue Criteria.

Analyte	Safety Issue Criteria[1]	Calculated/Measured Value
$Na_2NiFe(CN)_6$	1,000 g-mol	8,700 to 11,500 g-mol
ΔH (dry basis)	-75 cal/g	-9 cal/g
$^{239/240}Pu$	50 kg	1.1 to 1.4 kg
Temperature	300 °F (149 °C)	29 °C (85 °F)
Heat Load	11.72 kw	2.26 to 2.44 kw
Organic Content (TOC, Dry basis)	3.0 wt% TOC (10% sodium acetate equivalent)	0.75 wt% TOC

[1](Lindsey 1986; RHO 1988; Boyles 1992; Reep 1992).

8.3 RECOMMENDATIONS

The following recommendations are made based on the data and analyses presented in this report and the goals of the characterization effort:

- Gamma scanning assay of segments should be discontinued for SSTs until an adequate core sample from tank 241-BY-104 is obtained.

- The Watch List criteria for the Ferrocyanide waste tank USQ should be expanded so that it includes concentration or energetics based measurements.

- Tank 241-C-112 should be removed from the Ferrocyanide Watch List.

- A statistical comparison between the results from tank 241-C-109 (when available) and tank 241-C-112 should be performed to determine the degree of similarity between the tanks.

- An analysis for the formate anion should be done on the water soluble material to aid in confirming that the ferrocyanide has degraded.

- The DSC analysis should be performed under an inert atmosphere to better represent the waste conditions in the tank.

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Tank Characterization Panel
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